

RECYT

Year 24 / N° 38 / 2022 / 55–62

DOI: <https://doi.org/10.36995/j.recyt.2022.38.007>

## Extraction of Stevioside and Rebaudioside A from *Stevia rebaudiana* by percolation with ethanol-water solvents

### Extracción de Steviosido y Rebaudiósido A de *Stevia rebaudiana* por percolación con solventes hidroalcohólicos

Liliana Celaya<sup>1,2,\*</sup>, Darío Ferreyra<sup>1</sup>, Eugenio Kolb<sup>1</sup>, Nicolás Kolb Koslobsky<sup>1 (†)</sup>

1- Laboratorio Central, Facultad de Ciencias Exactas Químicas y Naturales, Universidad Nacional de Misiones 3300, Posadas, Argentina; 2- Consejo Nacional de Investigaciones Científicas y Técnicas (CONICET - UNAM), Universidad Nacional de Misiones 3300, Posadas, Argentina.

\* Email: [lilianacelaya@fceqyn.unam.edu.ar](mailto:lilianacelaya@fceqyn.unam.edu.ar)

Received: 21/01/2022; Accepted: 27/07/2022

#### Abstract

The aim of this study was to analyze the optimum temperature and percentage of ethanol to extract steviol glycosides (Stevioside and Rebaudioside A) from stevia leaves in a percolation extractor. The experiments were performed in broad ranges of binary solvent (0-100 % ethanol) and temperature (30–70 °C) in order to investigate their effects on steviol glycosides recovery and purity. The optimal extraction conditions, which gave a best steviol glycosides extraction kinetic were 70 °C and 35% of ethanol. In addition, the optimal extraction condition giving maximum purity was 70% of ethanol between 30 and 70°C.

Keywords: Steviol Glycosides; Crude Extract; Percolation; Ethanol-Water; Purity; Yield.

#### Resumen

El objetivo de esta investigación fue optimizar la temperatura y el porcentaje de etanol en la extracción de dos esteviol glicósidos (Esteviósido y Rebaudiósido A) desde hojas de stevia utilizando lecho percolador. Los experimentos se llevaron a cabo en todo el rango de porcentajes de etanol en agua (entre 0 y 100%) y a temperatura entre 30 y 70°C, para investigar su efecto en la recuperación y pureza de los esteviol glicósidos. La condición óptima que proporciona la mejor cinética de extracción de Esteviósido y Rebaudiósido A es 35% de etanol realizando la extracción a 70°C mientras que la condición que proporciona extractos de mayor pureza es 70% de etanol a temperatura entre 30 y 70°C.

Palabras clave: Steviol Glicósidos; Extracto Crudo; Percolación; Etanol-Agua; Pureza; Rendimiento.

#### Introduction

*Stevia rebaudiana* Bertoni, commonly called stevia, is a rich source of the natural sweeteners named steviol glycosides (SGs). Nowadays, stevia sweeteners are used in many countries in food formulations and for different pharmaceutical purposes. The percentage of SGs in stevia leaves varies from 4% to 20%, depending on the genotype, the phenotype and the agronomic practices (Kolb et al., 2001; Brandle and Telmer, 2007; Geuns, 2003). The main SGs in leaves are Stevioside, Stv (1 - 10%) and Rebaudioside A, RbA (2 - 13%); other SGs are normally found in smaller amounts (Morita et al., 2011; Wang et al., 2020; Pfister and Sehgal, 2012). Stv and RbA are largely

responsible for the sweetness in stevia leaves; these SGs determine the quality of stevia for production process (Jonnala et al., 2006; Kinghorn et al., 2001).

The method to obtain SGs from stevia, in general involves pretreatments, extraction, purifications and crystallizations (Chhaya et al. 2012; Chabot and Beaulieu, 2012). The first step to obtain SGs is traditionally a discontinuous solid-liquid extraction (maceration or decoction) of pre-treated leaves (Abelyan et al., 2006; Chabot and Beaulieu, 2012; Jonnala et al., 2006; Kienle, 1992). SGs are extracted from leaves with water or alcohols (Barba et al., 2015; Ameer et al., 2017). New extraction techniques are being used such as: microwave-assisted extraction, pressurized liquid extraction, supercritical-

fluid extraction, ultrasound-assisted extraction, high pressure assisted extraction and enzyme-assisted extraction (Bursać-Kovačević *et al.*, 2018; Xu *et al.*, 2019). Selection operation technique and operation condition affect directly the quality of stevia extract (Chabot and Beaulieu, 2012). The product in the extraction step (crude extract) is a colored and dense mixture containing SGs, among other components of leaves (Chabot and Beaulieu, 2012). Crude extract is normally subject to a multi-step purification to obtain individual components (Kienle, 1992; Payzant *et al.*, 1999; Wang *et al.*, 2020): Stv, RbA, Rebaudioside D, Rebaudioside C and SGs mixtures.

There are some desirable requirements for the industrial production of SGs: high processing capacity, high yields, high concentrations and high purity. In a previous work, Persinons (1973) used a percolator for the recovery of Stv from pre-treated stevia leaves. More recently Fromigoni *et al.* (2018) used a column to treat stevia leaves with ethanoic solvent. In another study, Ameer *et al.* (2017) optimized the ethanol-water solvent for the SGs recovery. Raspe *et al.* (2021) applied sonication for increase the recovery of SGs with ethanol-water solvents and Ciotta *et al.* (2022) used a semicontinuous extraction with boiling water followed by microfiltration to obtain purified stevia extracts used as sweetener.

In South American countries, there are diverse commercially-available stevia products since the approval of SGs as food additives (CAA, 2011; Celaya *et al.*, 2021; JECFA/69, 2008). Despite this, in Argentina there is currently no large-scale SGs industrial production. For this reason, there is a strong need to explore the possibilities inherent in stevia cultivation and SGs industrial production in the country.

In the present study, we aim to optimize the recovery of Stv and RbA from stevia leaves, in a semi-continuous process in column percolation by using ethanol-water as solvent, while working at moderate temperatures (from 30 to 70°C) in the extraction process. The kinetic for the recovery of Stv and RbA was established in the best conditions for the extraction.

## Materials and methods

### Chemicals and plant material

HPLC grade solvents were supplied by Merck (Darmstadt, Germany). Cicarelli (Reagents S.A., San Lorenzo, Argentina) supplied ethanol. The distilled water used for the extraction was deionized in a Milli-Q system. Stv (purity of 95.0%) and RbA (purity of 99.5%) were obtained in-house by recrystallizations in methanol and ethanol, as determined by a chromatographic analysis (Celaya *et al.*, 2021). The binary solvents employed in the percolation study

were prepared by mixing ethanol and water in adequate proportion to obtain dissolutions with density according to reported values (Celaya *et al.* 2016; Poling *et al.*, 2008). All chemicals were analytical grade.

Stevia was harvested in the experimental farm when the flowers were opened (0-5% open flower), in Posadas (Departamento Capital, Argentina). Samples were collected after sun-drying for 4 days. Stevia leaves were ground and sieved to obtain 70% of particle size of 5-40 mesh and 30% of particle size > 40 mesh (Norma IRAM 20514/2007). The SGs contents were quantified as described below. The moisture content was determined by drying samples (1-3 g) at  $102 \pm 2$  °C until it reached constant weight. Stevia leaves pre-dried and grinded contained the following SGs: RbA  $10.1 \pm 0.6\%$  and Stv  $2.4 \pm 0.1\%$ . The lot contained the following amount of moisture:  $7.9 \pm 3.1\%$ .

### Sample extraction with the percolator

The solid/liquid extraction of SGs was performed by means of a lab-scale percolator. The percolator consists in a cylindrical vessel in which the solids rest on a cotton bed on a perforated support (Fig. 1). The temperature of percolation was controlled by a thermostatic system with recirculation of heating water (Shott Gerate, model CTI 150, USA). The extraction temperature was measured by three thermometers. The solvent extraction was adjusted for the assay by immersing a glass beaker (containing the binary solvent) into the thermostatic system (Fig. 1).

The stevia weight and percolation flowrate were determined in preliminary experiments. For each test, 200 g of dried and grinded leaves were packed and impregnated with 1700 mL of solvent. After impregnation (6 min), the solvent was eluted by using gravity and 10 fractions of 100 mL were obtained. The total percolation time used was 60 min; the flowrate used was 16.7 mL/min. After extraction, the fractions were rapidly cooled to room temperature and filtered. Afterward, the fractions were mixed and a total extract of 1000 mL was used for the optimization.

### Soxhlet extraction

For comparison purposes, multiple-batch extractions were carried out employing a soxhlet system (Rouhani, 2019). The solvents used were: water, ethanol 96° and ethanol. Ground stevia leaves were pre-dried at 60°C for 2.0 h; then 20 g of stevia were placed in an extraction thimble. The thimble was placed in the soxhlet extractor. The plant material was extracted using 500 mL of solvent for 10 h. The residual solvent was removed by evaporation at 60°C up to constant weight. Extracted materials and extracts were used for analytical purposes.

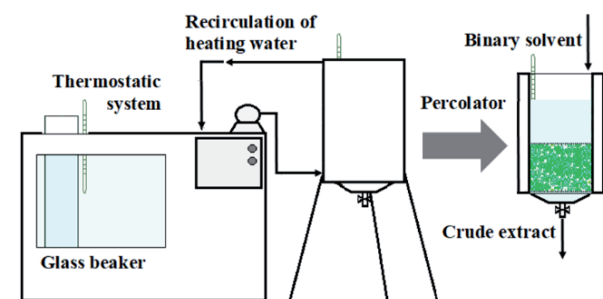


Figure 1: Schematic representation of the lab-scale percolator.

#### HPLC quantification of SGs in leaves and extracts

The quantification of SGs in leaves and extracts was performed with a high-performance liquid chromatography diode array detector (HPLC-DAD) by using the external standard method proposed by Kolb *et al.* (2001). The analyses were performed on a Prominence LC-20AT chromatograph (Shimadzu Corporation, Kyoto, Japan). The operating conditions were as follows: the column was Agilent Zorbax NH<sub>2</sub> (25.0 × 0.46 cm; 5 μm, particle size; Waters, Milford, MA), the UV detector was set at 210 nm, samples were run at room temperature, the mobile phase was acetonitrile:water 80:20 (v/v), and pH was 5.

For the calibration curve, stock solutions of 0.2–1.0 g/L of standard compounds (Stv and RbA) in ethanol-water 70:30 (w/w) were prepared. The parameters for external calibration curves were obtained by fitting experimental data through linear regression from replicate injections of standard solutions. The compounds in each sample were identified by comparing their retention times with those of the standards.

#### Factorial design

The solid/liquid extractions of SGs were carried out under the experimental conditions defined in a factorial design 3x5 (Table S1). The factorial design 3x5 was used to study the effect of temperature, (30°C, 50°C and 70°C) and ethanol percentage, (0%, 30%, 50%, 70%, 100%), on the responses: purity of extracts, retention of solvents in solid waste and RbA/Stv ratio. The statistical software Statgraphics Plus 5.1 software (Manugistics, Rockville, MD, USA) was used to carry out this study.

#### Optimization of concentration and recovery

Temperature (°C) and ethanol percentage (%) were optimized using response surface method. The dependent variables were: concentration (amount of SGs and solids extracted) and recovery (Stv and RbA). The total amount of solids extracted (TSE) was measured after drying 10 mL of crude extract at 102 ± 2 °C until it reached constant weight on a tared steel plate (Celaya *et al.*, 2021).

The experimental ranges selected for the independent variables were: temperature (from 30 to 70°C) and

percentage of ethanol in the binary solvent (from 0 to 100%). The experimental conditions and responses for conducting the experiments were analyzed following a pre-defined composite design with 19 experiments summarized in Table 1. Two replicates of the central point were conducted.

#### Equations

The variables determined in the extracts were: concentration (g/L), recovery of sweeteners (%; dried basis), purity (% of crude extracts) and yield (on a dried basis).

The concentration in the extract was determined as:

$$SGs = \frac{g(Stv + RbA)}{L} \quad (\text{Eq. 1})$$

$$TSE = \frac{g TSE}{L} \quad (\text{Eq. 2})$$

The recovery of sweeteners (on a dried basis) was calculated as:

$$\text{Recovery (RbA)} = \frac{\text{RbA recovered}}{\text{RbA in leaves}} \cdot 100\% \quad (\text{Eq. 3})$$

$$\text{Recovery (Stv)} = \frac{\text{Stv recovered}}{\text{Stv in leaves}} \cdot 100\% \quad (\text{Eq. 4})$$

The purity was determined as:

$$\text{Purity (SGs)} = \frac{\text{SGs in extract}}{\text{TSE in extract}} \cdot 100\% \quad (\text{Eq. 5})$$

The yield (on a dried basis) was calculated as:

$$\text{Yield (SGs)} = \frac{g \text{ SGs recovered}}{100 \text{ g leaves}} \cdot 100\% \quad (\text{Eq. 6})$$

$$\text{Yield (TSE)} = \frac{g \text{ TSE recovered}}{100 \text{ g leaves}} \cdot 100\% \quad (\text{Eq. 7})$$

A quadratic prediction (Eq.1) was used to optimize the SGs extraction:

$$Y = c + a \cdot X_1 + b \cdot X_2 + aa \cdot X_1^2 + bb \cdot X_2^2 + ab \cdot X_1 X_2 \quad (\text{Eq. 8})$$

In Eq. (8), *Y* is the response, concentration and recovery (SGs and TSE); *c*, *a*, *aa*, *b*, *bb*, *ab* are the coefficients of the equation. Significant effects (*P* < 0.05) were defined by using a standardized Pareto chart dropping coefficients without statistical significance. Models established were used to find the optimal mixture compositions with desirable responses (i.e., maximum concentration, maximum recovery) using the “Response Optimizer” (evaluated in Statgraphics Plus 5.1).

## Extraction kinetics

SGs extraction kinetics were undertaken in the stevia percolator under selected conditions for recovery and purity.

When performing the stevia percolation process in these selected conditions 10 fractions of 100 mL were obtained. The total percolation time used was 60 min. The fractions were rapidly cooled to room temperature and filtered. The 10 fractions (100 mL in 6 min each one) were used for the kinetic study. Extractions were carried out in triplicate. The experimental data was adapted to the model of Pilosof *et al.* (1985):

$$\text{Yield} = \frac{\text{Yield}_{\infty} t}{t_{1/2} + 0.5 t} \quad (\text{Eq. 9})$$

In Eq. (9), shows the yield of compound (Stv, RbA and TSE), that could be extracted; , the necessary time to extract half of . The quality of the fit was evaluated by using four statistical parameters (Scipioni *et al.*, 2010): R-squared ( $R^2$ ), Chi-square ( $\chi^2$ ), root square mean error (RSME) and the mean percentage error (MPE) (Eq. 10–13).

$$R^2 = 1 - \frac{\sum (\text{Yield} - \text{Yield}_{(\text{Eq.9})})^2}{\sum (\text{Yield} - \overline{\text{Yield}})^2} \quad (\text{Eq. 10})$$

$$\chi^2 = \sum \frac{(\text{Yield} - \text{Yield}_{(\text{Eq.9})})^2}{N - np} \quad (\text{Eq. 11})$$

$$\text{RMSE} = \left[ \frac{\sum (\text{Yield} - \text{Yield}_{(\text{Eq.9})})^2}{N} \right]^{0.5} \quad (\text{Eq. 12})$$

$$\text{MPE} = \frac{100}{N} \sum \left| \frac{\text{Yield} - \text{Yield}_{(\text{Eq.10})}}{\text{Yield}} \right| \quad (\text{Eq. 13})$$

## Results and discussion

### Optimization of concentration and recovery

Out of the dependent variables analyzed here, only the behavior of concentration (and recovery) can be described by the polynomial Eq. (8), proposed by the software Statgraphics Plus 5.1 by using the *standardized Pareto chart dropping coefficients without statistical significance. The final equation that predicts concentrations (and recovery) is:*

$$Y = c + a.X_1 + b.X_2 + bb.X_2^2 \quad (\text{Eq. 14})$$

The response surface plots that predict the concentration of extracts (SGs and TSE), for extraction volume of 1000 mL (Table 1) are shown in Fig A1 (Appendix). Coefficients c, a, b, bb are shown in Table 2.

The extraction efficiency could be influenced by multiple parameters, such as time, temperature and solvent polarity and their effects may be either independent or interactive (Celaya *et al.*, 2015). The responses show linear temperature dependence and a quadratic dependence of ethanol percentage (Table 2, Fig.A1). Different results were obtained in another study (Raspe *et al.*, 2021), investigating the influence of temperature, and binary mixtures of water and ethanol as solvent.

**Table 1:** Experimental design and responses of the composite design.

Independent variables		Reduced variables		Run	Responses <sup>a</sup>					
X <sub>1</sub> (T, °C)	X <sub>2</sub> (Ethanol, %)	X <sub>1'</sub>	X <sub>2'</sub>		SGs (g/L)	TSE (g/L)	RbA/Stv	Stv recovered (%)	RbA recovered (%)	SGs purity (%)
30	0	-1.0	-1.0	4	16.4	56.2	4.2	61.2	60.3	29.2
30	30	-1.0	-0.4	12	19.3	64.0	4.3	69.8	71.5	30.1
30	50	-1.0	0.0	3	16.7	53.7	4.2	62.3	61.6	31.2
30	70	-1.0	+0.4	13	15.5	47.6	3.8	61.5	56.2	32.6
30	100	-1.0	+1.0	18	1.7	9.8	3.7	7.1	6.2	17.7
36	15	-0.7	-0.7	20	17.1	56.0	3.5	73.5	60.8	30.6
36	85	-0.7	+0.7	17	9.9	31.8	2.9	48.5	33.6	31.1
50	0	0.0	-1.0	16	17.8	61.8	4.6	61.7	66.8	28.9
50	30	0.0	-0.4	14	22.7	79.5	4.3	82.7	83.8	28.5
50	50	0.0	0.0	10	22.1	74.1	4.3	80.2	81.9	29.8
50	50	0.0	0.0	6	19.8	70.6	4.3	72.5	73.3	28.1
50	50	0.0	0.0	8	22.2	75.1	4.0	84.6	81.2	29.6
50	70	0.0	+0.4	5	18.6	56.9	4.1	70.4	68.0	32.6
50	100	0.0	+1.0	11	2.6	12.6	3.7	10.8	9.5	20.9
64	15	+0.7	-0.7	15	21.1	68.3	3.5	90.6	74.7	30.9
64	85	+0.7	+0.7	19	17.1	49.3	3.3	77.1	59.8	34.7
70	0	+1.0	-1.0	2	17.4	68.8	3.7	71.7	62.5	25.3
70	30	+1.0	-0.4	9	20.8	73.6	4.1	77.9	76.5	28.3
70	50	+1.0	0.0	1	20.7	76.4	3.9	80.6	75.3	27.1
70	70	+1.0	+0.4	7	23.2	74.4	3.9	92.1	84.1	31.2
70	100	+1.0	+1.0	21	5.3	22.3	3.1	24.8	18.3	23.8
100	0 <sup>b</sup>							91.8 (Stv+RbA)		
100	0 <sup>c</sup>							91.6-93.8 SGs		24.1-25.7
78	0 <sup>d</sup>							97.2		
68	51 <sup>e</sup>									
50	34 <sup>f</sup>									21.8
Soxhlet extraction <sup>g</sup>		Water			5.5	21.7	4.2	> 95.0	> 95.0	25.1
		Ethanol 96%			4.7	15.7	4.0	84.1	83.1	30.0
		Ethanol			4.2	15.0	3.4	68.6	59.5	28.0

<sup>a</sup> In each extraction, 200 g of stevia were extracted and 1000 mL of crude extract were eluted using gravity

<sup>b</sup> Jentzer *et al.* (2015)

<sup>c</sup> Ciotta *et al.* (2022)

<sup>d</sup> Chhaya *et al.* (2012)

<sup>e</sup> Xu *et al.* (2019)

<sup>f</sup> Raspe *et al.* (2021)

<sup>g</sup> In the soxhlet extraction, 20 g of stevia were extracted with 500 mL of solvent

Temperatures from 50°C to 78°C were selected in other works, to obtain higher SGs values with different extraction process (Afandi *et al.*, 2013; Chhaya *et al.*, 2012; Raspe *et al.*, 2021; Rouhani, 2019; Xu *et al.*, 2019). Higher temperatures lead to better SGs solubility, faster diffusion rates, lower solvent viscosities and weakening of the solute–matrix interactions (Ciotta *et al.*, 2022; Jentzer



et al., 2015). However, when working with alcohol-water mixtures, higher temperatures may cause preferential vaporization and ethanol loss (Singh, 2008). The high temperature used here is below the boiling point of ethanol, avoiding high loss of preferential evaporation of the organic solvent during the extraction process.

The selection of the best solvent to extract the phytochemicals from plant materials is essential (Rouhani, 2019). The effects of ethanol percentage on the SGs recovery indicates that binary solvents resulted in a higher concentration of SGs in crude extracts (Fig. 2 and Fig. A1). These results are in accordance with previous findings (Celaya et al., 2016; Raspe et al., 2021; Xu et al., 2019). In general, binary solvents effectively promote the solubilization of SGs, and this behavior is favored by the presence of other SGs and biomolecules. Previously, Afandi et al. (2013) obtained a different result in the recovery of SGs with the best recovery in pure alcohol.

**Table 2:** Coefficients of the Eq. (14) for the prediction of concentration (SGs and TSE).

Response, Y	Coefficients				Determination coefficient	
	c	a	b	bb	R <sup>2</sup>	R <sup>2</sup> (D.F.)
SGs (g/L)	10.9127	0.1072	0.3098	-0.0042	90.6 %	88.9 %
	(P<0.001)*	(P=0.0028)	(P<0.001)	(P<0.001)		
TSE (g/L)	36.9029	0.4392	0.8653	-0.0128	94.4 %	93.5 %
	(P<0.001)	(P<0.001)	(P<0.001)	(P<0.001)		

\* P, statistical significance in bracket; D.F.: 20.

The effect of temperature and ethanol percentage on the Stv and RbA recovery is shown in Fig. 2. According to the results obtained (Fig. 2.A and Fig. 2.B), binary solvents are superior to those found in water and ethanol up to 70 °C, since it allows for a more rapid recovery of Stv and RbA at moderate temperatures.

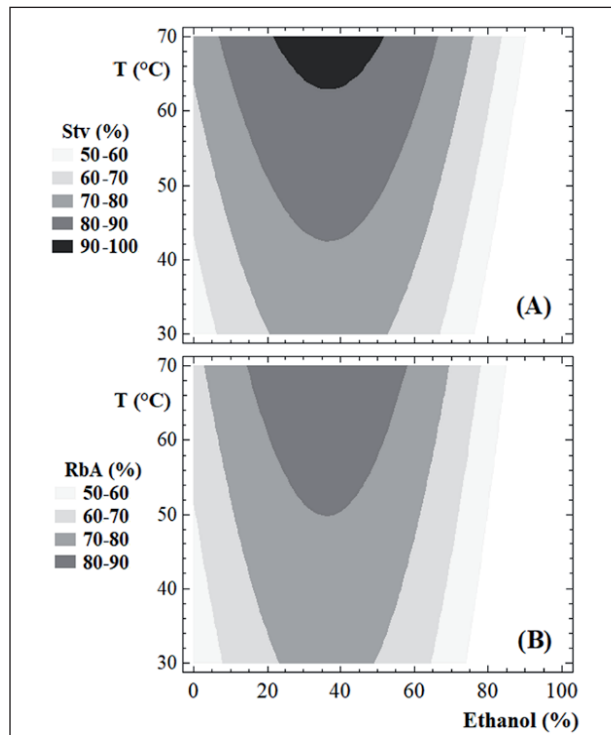
Recovery of Stv and RbA were maximized and the optimized process conditions were: temperature: 70°C and ethanol: 35%. When the solvent to leaf ratio was equivalent to 5:1 (1000 mL/ 200 g), the predicted recoveries were 93.0% and 87.5% for Stv and RbA, respectively.

#### Effects of percolation conditions on purity

In order to conduct research on the effects of percolation conditions on the purity, a 3x5 factorial design was used (Table A1). The effect of ethanol percentage on the SGs purity is shown in Fig. 3 (and Table 1). It indicates that binary solvents with high percentage of ethanol result in higher purity of crude extracts. On the other hand, over the temperature range (between 30-70 °C), temperature had no effect on the purity ( $P > 0.05$ ; CL 95%).

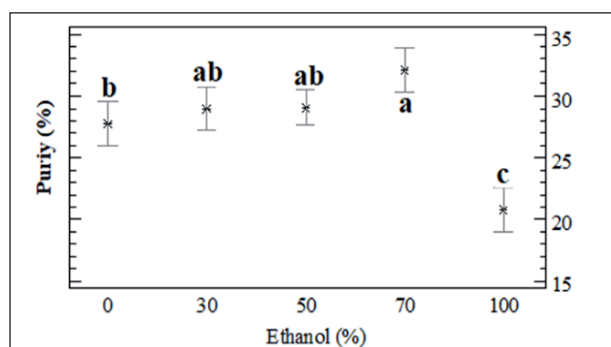
Selection of the solvent for a percolation process depends on the solubility of the desired components in the solvent (Singh, 2008). The impurities present in stevia extracts are components co-extracted with SGs from the leaves (Žlabur et al., 2015; Xu et al., 2019); the impurities

include saccharides and crude fiber, chlorophylls, phenolic compounds, resins, phytosterols and triterpenes, proteins, organic acids, fatty acids, gums and colloids (Celaya et al., 2021; Chabot and Beaulieu, 2013).



**Figure 2:** Contour plot of predicted recovery of Stv (A) and RbA (B).

For Ameer et al. (2017) and Xu et al. (2019), 75% and 59% of ethanol v/v (76% and 59% of ethanol w/w) was superior to other binary solvents for a maximum SGs yield in the extraction of Stv and RbA from stevia. Different was for Raspe et al. (2021), with the best result at 40% of ethanol v/v (34% w/w).



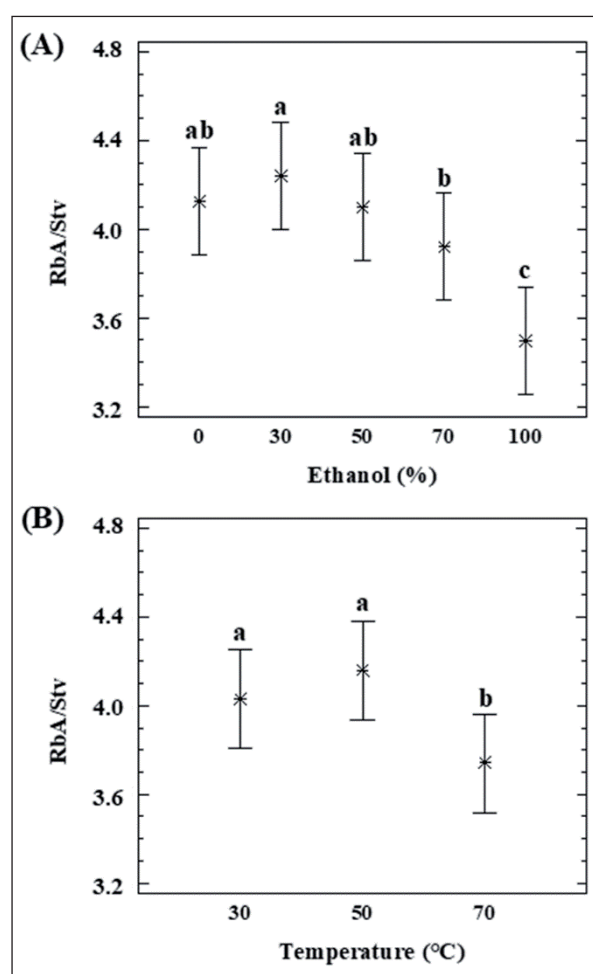
**Figure 3:** Dependence of purity (Eq. 5) with ethanol (% w/w) in a binary solvent for SGs extraction (statistical significance,  $P < 0.004$ ; D.F. 14).

In order to obtain a better SGs recovery in the present study, the 35% (w/w) was optimal (Fig. 2.A and Fig. 2.B). On the other hand, a percentage of ethanol of 70% (w/w), similar to proposed previously (Ameer et al., 2017; Chabot and Beaulieu, 2012; Kolb et al., 2001), was appropriate for the recovery of an extract with the best purity. In the present study, the high purity was for 85% w/w of ethanol

at 70°C (Table 1, Fig. 3), but the recovery with this binary solvent is very low (77.1% for Stv and 59.8% for RbA).

#### RbA/Stv ratio

The effect of the percolation conditions on the RbA/Stv ratio over the extracts was also been statistically analyzed by means of the 3x5 factorial design (Table A1). The RbA/Stv relations that appear in the extracts are summarized in Table 1. Stv and RbA show a similar behavior in the percolation, but there are some differences regarding the recoveries (Figure 2). Xu *et al.* (2019) obtained similar results previously, in the recovery of SGs for analytical purposes. RbA contains one more glycosyl group than Stv. RbA can form more hydrogen bonds with water or alcohols (Celaya *et al.*, 2016). A higher RbA in relation to Stv was favored by the presence of a high percentage of water in the binary solvent (Figure 4.A). Temperature had an impact on this variation. Such differences could be reflected on the RbA/Stv ratio (Figure 4.B). Based on the results, in general lower temperatures and lower percentage of ethanol improve RbA/Stv ratio.



**Figure 4:** Dependence of RbA/Stv ratio with (A) a percentage of ethanol (P 0.0028, D.F. 14) and (B) temperature (P 0.0088; D.F. 14).

#### Soxhlet extraction

Soxhlet apparatus is a laboratory equipment that works at the boiling temperature (Takeuchi *et al.*, 2009). The procedure was carried out in this study with ethanol, ethanol 96 and water. The obtained yields in these extractions are shown in Table 1. Previously, Rouhani (2019) obtained different results for the soxhlet extraction of Stv from stevia leaves. According to the results obtained here, polar solvents result in a higher RbA/Stv ratio and high recovery. The solvent ethanol 96% result in a higher SGs purity.

In other study, Ciotta *et al.* (2022) used semicontinuous percolation with boiling water as solvent to recover SGs from stevia leaves. In the study, Ciotta and co-workers used five extractive cycles to recover 91.64 of the SGs present in the leaves. Here the best recovery was for soxhlet extraction using boiling water as solvent (recovery > 95.0 %).

#### Extraction kinetics

The chosen conditions for the kinetic study were 35% of ethanol at 70°C (condition 1, fast recovery) and 70% of ethanol at 70°C (condition 2, best purity). SGs extraction kinetics were undertaken in the percolator under these selected conditions. The data of ten fractions of 100 mL of crude extract obtained in each condition (percolation extract, 1000 mL), were adapted to the model of Pilosof *et al.* (1985). This model was previously applied to study the release rate of RbA from matrices formed with agglutinants in *yerba-mate* powder (Scipioni *et al.*, 2010). The results are in agreement with the recovery and purity previously estimated.

Figures 5.A and 5.B show the experimental and predicted yields obtained with the Pilosof model. The regression coefficients and statistical parameters (Eq. 9) are listed in Table 3. A good fitness was obtained in all cases;  $R^2$  values were high (>95%), and RSME and MPE were low.

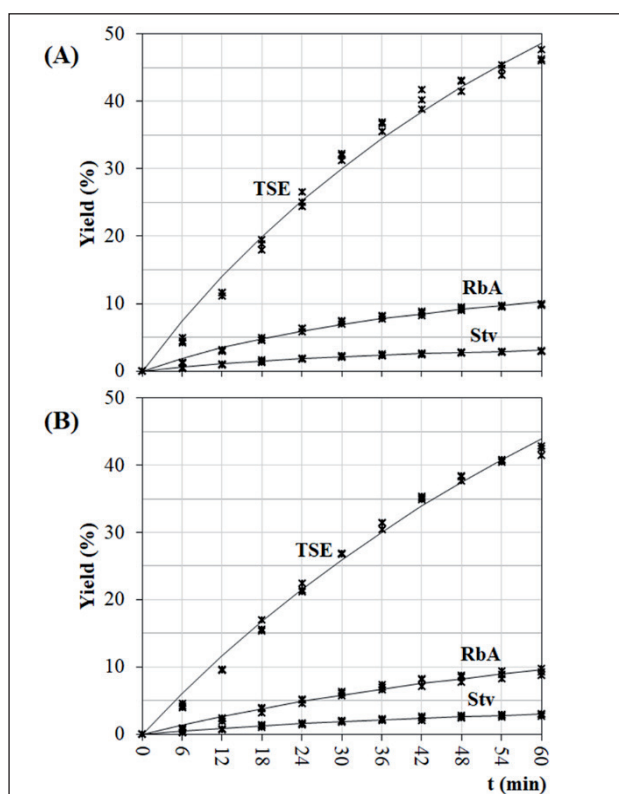
According to the model, high values of  $t_{1/2}$  indicate a low recovery rate of compounds. The highest value of  $t_{1/2}$  was obtained for TSE with 70% of ethanol at 70°C (68.6 min). SGs showed the lowest values, in 35% of ethanol at 70°C, 27.9 and 28.9 min for Stv and RbA, respectively (Table 3).

For Raspe *et al.* (2021), the increase in the solvent/leaf ratio provided the highest yield of total solids and SGs, as the increase of ethanol in the solvent, while the temperature promoted an increase only in the of total solids. The result of kinetic behavior obtained here are in accordance with the results of Raspe and coworkers from 54 min of percolation (Figures 5.A and 5.B).

**Table 3:** Constants and statistical parameters of the kinetic model of Pilosof (Eq. 9) for optimal conditions.

Optimal condition	Response	Coefficients <sup>a</sup>		Statistical parameters			
		Yield <sub>∞</sub>	t <sub>1/2</sub>	R <sup>2</sup>	χ <sup>2</sup>	RSME	MPE
Condition 1:	Stv yield	2.9	27.9	98.4	0.01	0.11	0.04
35% of ethanol at 70 °C	RbA yield	10.1	28.9	98.5	0.15	0.37	0.06
	TSE yield	64.0	48.4	98.4	3.68	1.85	0.08
Condition 2:	Stv yield	3.8	47.7	96.6	0.03	0.17	0.08
70% of ethanol at 70 °C	RbA yield	13.7	55.3	97.6	0.21	0.48	0.08
	TSE yield	72.3	68.6	99.1	1.73	1.27	0.06

<sup>a</sup> The coefficients of the kinetic model were adjusted to 100 g of dried leaves as calculation base; D.F.: 29

**Figure 5:** Extraction kinetics in the best conditions: (A) 35% of ethanol at 70 °C; (B) 70% of ethanol at 70 °C. Coefficients for the kinetic models (Eq. 9), are listed in Table 3.

## Conclusions

In this study we aim to optimize the recovery of Stv and RbA in a stevia percolator by using mixtures of ethanol and water, while working at moderate temperatures in the extraction process. When using ethanol-water as a solvent for the percolation, the optimum extraction temperature is 70°C. Regarding the percentage of ethanol, the statistical analysis showed that the optimal extraction conditions, which gave the best SGs extraction kinetic is 35% of ethanol, but the highest purity (at least 31%) is obtained with 70% of ethanol. Based on the results, it is not possible to obtain the fast SGs recovery and high purity at the same time in the stevia percolation extraction with binary solvent.

The obtained results could be used industrially for the process of stevia leaves obtaining crude extracts with high purity or high yield.

## Acknowledgements

Financial support by Project UNaM-16Q511 and Project UNaM-16Q1204-IDP. The authors would like to thank Evangelina Nuñez and Cristina Mayol for their support and collaboration.

## References

1. Afandi, A.; Sarijan, S.; Shaha, R.K. *Optimization of Rebaudioside A extraction from Stevia rebaudiana (Bertoni) and quantification by high performance liquid chromatography analysis*. Journal of Tropical Resources and Sustainable Science 1: 62-70. 2013.
2. Ameer, K.; Bae, S.-W.; Jo, Y.; Lee, H.-G.; Ameer, A.; Kwon, J.-H. *Optimization of microwave-assisted extraction of total extract, stevioside and rebaudioside-A from Stevia rebaudiana (Bertoni) leaves, using response surface methodology (RSM) and artificial neural network (ANN) modelling*. Food Chemistry 229: 198-207. 2017.
3. Brandle, J.E.; Telmer P.G. *Steviol glycoside biosynthesis*. Phytochemistry. 68, 1855–1863.
4. Bursać-Kovačević, D.; Marasa, M.; Barba, F.J.; Granato, D.; Roohinejad, S.; Mallikarjunan, K.; Montesano, D.; Lorenzo, J.M.; Putnik, P. *Innovative technologies for the recovery of phytochemicals from Stevia rebaudiana Bertoni leaves: A review*. Food Chemistry 268: 513-521. 2018.
5. CAA. Código Alimentario Argentino Cap. XVIII, *Aditivos Alimentarios*. Artículo 1398 [update 6/2011]. [www.anmat.gov.ar](http://www.anmat.gov.ar)
6. Celaya, L.; Martina, P.; Kolb-Koslobsky N. *Infusions prepared with Stevia rebaudiana: application of a simplex centroid mixture design for the study of natural sweeteners and phenolic compounds*. Journal of Food Science and Technology 59: 55-64. 2022.
7. Celaya, L.; Kolb, E.; Kolb, N. *Solubility of Stevioside and Rebaudioside A in water, ethanol and their binary mixtures*. International Journal of Food Studies 5: 158-166. 2016.
8. Chabot, S.; Beaulieu, M. *Patent Application Publication* US/2013/0108718. 2013.
9. Chhaya, R.; Majumdar, G.C.; Sirshendu, D. *Optimization of process parameters for water extraction of stevioside using response surface methodology*. Separation Science and Technology. 47: 1-9. 2012.
10. Ciotta, S. R.; Zorzenon, M.R. T.; Dacome, A. S.; Hodas, F.; de Almeida do Couto, J. M. F. Fernandes, P. G. M.; Mareze da Costa, C. E.; da Costa, S. C. *Extraction of sweeteners from Stevia*

- rebaudiana* by semicontinuous percolation of untreated leaves and leaves pretreated with ethanol. *Journal of Food Processing and Preservation* 46(3): e16303. 2022.
11. Formigoni, M.; Gimenez-Milani, P.; da Silva-Avíncola, A.; dos Santos, V.J.; Benossi, L.; Dacome, A.S.; Pilau, E.J.; da Costa S.C. Pretreatment with ethanol as an alternative to improve steviol glycosides extraction and purification from a new variety of stevia. *Food Chemistry* 241: 452-459. 2018.
  12. Geuns, J.M. *Molecules of Interest, Stevioside*. *Phytochemistry* 64: 913-921. 2003.
  13. JECFA/69. *Steviol glycosides, chemical and technical assessment (CTA, 2007) revised by harriet wallin for the 69th JECFA*. Original Prepared by Harriet Wallin for the 63rd JECFA, Revised by Paul M. Kuznesof for the 68th JECFA. 2008. [www.fao.org/ag/agn/agns/jecfa/cta/69/Steviol\\_glycosides\\_CTA\\_69.pdf](http://www.fao.org/ag/agn/agns/jecfa/cta/69/Steviol_glycosides_CTA_69.pdf)
  14. Jentzer, J.; Alignan, M.; Vaca-garcia, C.; Rigal, L.; Vilarem, G. Response surface methodology to optimise Accelerated Solvent Extraction of steviol glycosides from *Stevia rebaudiana* Bertoni leaves. *Food Chemistry* 166: 561-567. 2015.
  15. Kienle, U. *Patent Application Publication* US/1992/5112610. 1992.
  16. Kinghorn, A.D.; Wu, C.D.; Soejarto, D.D. In: *Alternative Sweeteners Third Edition*, Revised and Expanded (p. 167-184). Marcel Dekker Inc. 2001.
  17. Kolb, N.; Herrera, J.L.; Ferreyra, D.J.; Uliana R.F. *Analysis of Sweet Diterpene Glycosides from Stevia rebaudiana: Improved HPLC Method*. *Journal of Agricultural and Food Chemistry* 9: 4538-4541. 2001.
  18. Morita, T.; Morita, K.; Komai, K. *Patent Application Publication* US/2011/7884265B2. 2011.
  19. Payzant, J.D.; Laidler, J.K.; Ippolito, R.M. *Patent Application Publication* US/1999/5962678. 1999.
  20. Persinos, G.J. *Patent Application Publication* US/1973/3723410. 1973.
  21. Pfister, W.R.; Sehgal, S. *Patent Application Publication* US/2011/0021111A1. 2012.
  22. Pilosof, A.M.R.; Boquet, R.; Bartholomai, G.B. *Kinetics of water uptake by food powders*. *Journal of Food Science* 50: 278-283. 1985.
  23. Poling, B.E.; Thomson, G.H.; Friend, D.G.; Rowley, R.L.; Wilding, W.V. *Physical and Chemical Data*. McGraw-Hill Ed., *Perry's Chemical Engineers' Handbook*. McGraw-Hill Companies. 2008.
  24. Raspe, D.T.; Ciotta, S.R.; Zorzenon, M.R.T.; Dacome, A.S.; da Silva, C.; Milani, P.G.; da Costa, S.C. *Ultrasound-assisted extraction of compounds from Stevia leaf pretreated with ethanol*. *Industrial Crops and Products* 172: 114035. 2021.
  25. Rouhani, M. *Modeling and optimization of ultrasound-assisted green extraction and rapid HPTLC analysis of stevioside from Stevia Rebaudiana*. *Industrial Crops and Products* 132: 226-235. 2019.
  26. Scipioni, G.P.; Ferreyra, D.J.; Acuña M.G.; Schmalko, M.E. *Rebaudioside A release from matrices used in a yerba maté infusion*. *Journal of Food Engineering* 100: 627-633. 2010.
  27. Singh, J. In: *Extraction Technologies for Medicinal and Aromatic Plants* (pp. 72-73). International Centre for Science and High Technology Eds. 2008.
  28. Takeuchi, T.M.; Pereira, C.G.; Braga, M.E.M.; Maróstica, M.R.Jr.; Leal, P.F.; Meireles M.A.A. In: *Extracting Bioactive Compounds for Food Products: Theory and Applications*, chapter 4. Taylor & Francis Group. 2009.
  29. Wang, J.; Zhao, H.; Wang, Y.; Lau, H.; Zhou, W.; Chen, C.; Tan, S. *A review of stevia as a potential healthcare product: Up-to-date functional characteristics, administrative standards and engineering techniques*. *Trends in Food Science and Technology* 103: 264-281. 2020.
  30. Xu, S.; Wang, G.; Guo, R.; Wei, Z.; Zhang, J. *Extraction of steviol glycosides from Stevia rebaudiana (Bertoni) leaves by high-speed shear homogenization extraction*. *Journal of Food Processing and Preservation* 43: e14250. 2009.
  31. Zaidan, U.H.; Zen N.I.M.; Amran N.A. *Biochemical evaluation of phenolic compounds and steviol glycoside from Stevia rebaudiana extracts associated with in vitro antidiabetic potential*. *Biocatalysis and Agricultural Biotechnology* 18: e101049. 2019.
  32. Žlabur, J.S.; Voća, S.; Dobričević, N.; Brnčić, M.; Dujmić, F.; Brnčić, S.R. *Optimization of ultrasound assisted extraction of functional ingredients from Stevia rebaudiana Bertoni leaves*. *International Agrophysics* 29: 231-237. 2015.





**Available in:**

<https://www.redalyc.org/articulo.oa?id=382679632007>

How to cite

Complete issue

More information about this article

Journal's webpage in redalyc.org

Scientific Information System Redalyc  
Diamond Open Access scientific journal network  
Non-commercial open infrastructure owned by academia

Liliana, Celaya, Darío, Ferreyra, Eugenio, Kolb,  
Nicolás, Kolb Koslobsky

**Extraction of Stevioside and Rebaudioside A from Stevia rebaudiana by percolation with ethanol-water solvents**  
**Extracción de Steviosido y Rebaudiósido A de Stevia rebaudiana por percolación con solventes hidroalcohólicos**

*Revista de Ciencia y Tecnología*

no. 38, p. 1 - 70, 2022

Centro de Investigación y Desarrollo Tecnológico, Facultad de Ciencias Exactas, Química y Naturales, Universidad Nacional de Misiones,

**ISSN:** 0329-8922

**ISSN-E:** 1851-7587

**DOI:** <https://doi.org/10.36995/j.recyt.2022.38.007>