

# Influence of Agitator Geometry in the Dissolution Kinetics of a Spherical Solid

## Influencia de la geometría del agitador en la cinética de disolución de un sólido esférico

Gustavo A. Orozco <sup>1</sup>, Clara T. González-Hidalgo<sup>2</sup>, and Fabio Fajardo<sup>3</sup>

### ABSTRACT

An experimental study was conducted with the purpose of establishing the effect of both agitator geometry and agitation rate on the dissolution kinetics of spherical candies. To this effect, the dissolution constants were determined in aqueous solutions by way of the Hixson-Crowell model for three types of agitators, namely a paddle agitator with two blade, and two turbine agitators with four and six blades, respectively. Subsequently, for a fixed agitation rate, the temperature was varied in order to experimentally determine the activation energy of each agitator. It was found that, while the geometry of the agitator has an effect on the pre-exponential factors of the dissolution constant, the activation energy remains constant. Finally, differences in the dissolution constants as a function of the angular frequency of the agitator were also found.

**Keywords:** solid dissolution constant, activation energy, agitator geometry, agitation rate

### RESUMEN

Se realizó un estudio experimental con el fin de establecer el efecto de la geometría de un agitador y su tasa de agitación en la cinética de disolución de dulces esféricos. Para ello se determinaron las constantes de disolución en soluciones acuosas usando el modelo de Hixson-Crowell en tres tipos de agitadores, a saber: un agitador de paleta de dos aspas, y dos agitadores tipo turbina de cuatro y seis aspas respectivamente. Posteriormente, para una tasa fija de agitación, se varió la temperatura con el propósito de determinar experimentalmente la energía de activación de cada agitador. Se encontró que, mientras la geometría del agitador tiene un efecto en el factor pre-exponencial de la constante de disolución, la energía de activación permanece constante. Finalmente, también se encontraron diferencias en las constantes de disolución en función de la frecuencia angular del agitador.

**Palabras clave:** constante de disolución de sólidos, energía de activación, geometría del agitador, tasa de agitación

**Received:** March 4<sup>th</sup>, 2021

**Accepted:** October 06<sup>th</sup>, 2021

### Introduction

The dissolution phenomenon plays an important role in several industrial and biological processes such as drug delivery via mucosal surfaces (Jug *et al.*, 2018), water treatment (Pham, Sedlak and Doyle, 2012), or controlled release of inhibitors from coatings (Pham *et al.*, 2012), among others. Once a solute is introduced in a solvent, it is expected to engage in a dynamical process where the solute particles migrate from the highest to the lowest concentration regions. This phenomenon is strongly dependent on thermodynamic variables such as pressure, temperature, and the nature of the solvent.

From a mathematical point of view, the dissolution rate can be accounted for by using the Noyes-Whitney equation: (Noyes and Whitney, 1897)

$$-\frac{dm}{dt} = \frac{DS}{h} (C_s - C_t) \quad (1)$$

where  $m$  is the solid mass at a given time  $t$ ,  $S$  is the solid surface area,  $h$  is the diffusion layer thickness,  $C_s$  is the solute solubility,  $C_t$  is the solute concentration at a time  $t$ , and  $D$  is the diffusion coefficient. It is known that the diffusion coefficient characterizes the ease with which each particle moves into a given solute and depends on many factors such as the size and shape of the solute, solvent viscosity, and temperature.

<sup>1</sup> Chemical Engineer, Universidad de América, Colombia. Physicist, Universidad Nacional de Colombia, Colombia. MSc Chemical Engineering, Universidad Nacional de Colombia, Colombia. PhD Chemical Engineering, Universitat Rovira i Virgili, Spain. Affiliation: Assistant Professor, Universidad Nacional de Colombia, Colombia. E-mail: gaorozcoa@unal.edu.co

<sup>2</sup> Physicist, Universidad Nacional de Colombia, Colombia. MSc Physics, Universidad Nacional de Colombia, Colombia. PhD Chemical Engineering, Universitat Rovira i Virgili, Spain. Affiliation: Assistant Professor, Pontificia Universidad Javeriana, Colombia. E-mail: gonzalezclara@javeriana.edu.co

<sup>3</sup> Physicist, Universidad Nacional de Colombia, Colombia. PhD Physics, Universidade Estadual de Campinas, Brasil. Affiliation: Associate Professor, Universidad Nacional de Colombia, Colombia. E-mail: fefajardot@unal.edu.co

**How to cite:** Orozco, G. A., González-Hidalgo, C. T., Fajardo, F. (2022). Influence of Agitator Geometry in the Dissolution Kinetics of a Spherical Solid. *Ingéniería e Investigación*, 42(3), e94110. <https://doi.org/10.15446/ing.investig.94110>



Attribution 4.0 International (CC BY 4.0) Share - Adapt

If a solute with a spherical geometry is considered, it is possible to see that the integration of [Equation \(1\)](#) results in the Hixson-Crowell model ([Hixson and Crowell, 1931](#)). Details about the procedure can be found in the literature ([Beauchamp, 2001](#); [Hixson and Crowell, 1931](#); [Hixson and Baum, 1942](#)).

$$m^{1/3} = m_0^{1/3} - kt \quad (2)$$

where  $k$  represents the constant kinetic of dissolution given by

$$k = \frac{D}{h} \left( \frac{4\pi}{3} \right)^{1/3} \left( \frac{1}{p} \right)^{2/3} C_s. \quad (3)$$

It can be observed that  $k$  depends on density ( $\rho$ ), diffusivity ( $D$ ), and solubility parameters ( $C_s$ ).  $k$  is also explicitly dependent on temperature through an Arrhenius-like law given by the following Equation:

$$k = Ae^{-E_a/RT} \quad (4)$$

where  $A$  is the pre-exponential factor,  $E_a$  is the activation energy,  $R$  is the ideal gas constant, and  $T$  is the temperature.

Although the dissolution constants of spherical solids have been previously reported, to the best of our knowledge, the effect of the agitator geometry on it has not yet been considered. This effect is of great importance to understand the dissolution dynamics in stirred tanks where additional variables need to be considered, such as the type of impeller device and the proportions of the tank, deflectors, and agitators ([Mc Cabe et al., 2007](#)). Previous works have been carried out on the effect of different impeller features on different processes. For instance, ([Mangwandi et al., 2010](#)) showed that, for a high-shear granulator, impeller speed has an effect on granule size distribution, as well as on structure and shape, which in turn affects the dissolution rate measured by means of the time a certain percentage of granules takes to dissolve. ([Legendre et al., 2016](#)) simulated bubble dissolution in a fluid tower to optimize the process of  $CO_2$  mineralization. Using a specific impeller geometry, they calculated the Hatta number, demonstrating that the effect of the chemical reaction is minimal in contrast to the mass transfer process, as well as presenting the dissolution of mass over time. ([Kravtchenko et al. \(1999\)](#) studied the dissolution of food hydro-colloids, for which they empirically found a certain configuration between stirrer and container in order to overcome the issue with lump formation. They were able to quantify dissolution kinetics by assessing the mass dissolved over time. ([Jirout et al., 2011](#)) showed that mixing system aspect ratios have an effect on the mixing of suspensions calculated through the Froude number and dimensionless power consumption

necessary for off-bottom particle suspension. ([Thakur et al., 2004](#)) demonstrated that, in complex fluids (Newtonian and non-Newtonian), the column-to-impeller distance has an effect on mixing, which was evidenced by means of computing the effective shear rate and the power consumption. ([Akrap et al., 2012](#)) studied the effect of variation in the ratio of the impeller diameter to tank diameter, on the nucleation rate of borax dehydrate and crystal growth, but they did not examine dissolution. ([Abreu-Lopez et al., 2017](#)) numerically computed the effect of different impeller designs in a batch reactor for aluminum degassing, estimating pressure fields and dimensionless oxygen concentration with respect to time. ([Alok et al., 2014](#)) analyzed the effect of different kinds of impeller on mixing in an aerobic stirred tank fermenter, determining the volumetric mass transfer coefficient. They reported a particular kind of impeller to be the most efficient in regards to mass transfer in comparison with the others by measuring the turbulent dissipation rate, which, according to the authors, represents the volumetric mass transfer rate. ([Ćosić et al., 2016](#)) studied the influence of the impeller blade angle on the crystallization kinetics of borax. Some of the findings were that the flow pattern of the liquid charge changes significantly with the blade angle, the rate of nucleation increases as the blade angle decreases, and increasing the impeller blade angle significantly impacts the crystal growth rate constant. By means of electrical resistance tomography and computational fluid dynamics, ([Kazemzadeh et al., 2020](#)) analyzed the effect of impeller type on the mixing of highly concentrated large-particle slurries. Using three types of axial impellers, they found that the impeller type has impact on the homogeneity of the mixture, the average particle velocity along the tank, and the energy dissipation, which therefore affects power consumption. ([Gu et al., 2020](#)) studied floating particle drawdown and dispersion in a stirred tank with four pitched-blade impellers using computational fluid dynamics simulation. The type of impeller can reduce the power consumption at the same impeller speed, and it can also enhance the solid integrated velocity and the level of homogeneity for the floating particle dispersion process under constant power consumption. ([Liu et al., 2021](#)) investigated gas-liquid mass transfer and power consumption in jet-flow high-shear mixers. They found that gas-liquid mass transfer performance and power consumption are significantly affected by structural parameters, among them rotor blade angle and rotor blade arc. They also found the gas-liquid mass transfer performance is improved with increasing blade angle and that power consumption is reduced by decreasing the blade angle.

This work aims to experimentally study the effect of impeller geometry on the dissolution rate of spherical candy. In order to do that, the dissolution constant rate will be evaluated considering variations in the geometry of the stirrers, their agitation angular frequency, and the solvent temperature. The activation energy of the

dissolution is determined, as well as the pre-exponential factors for every agitator.

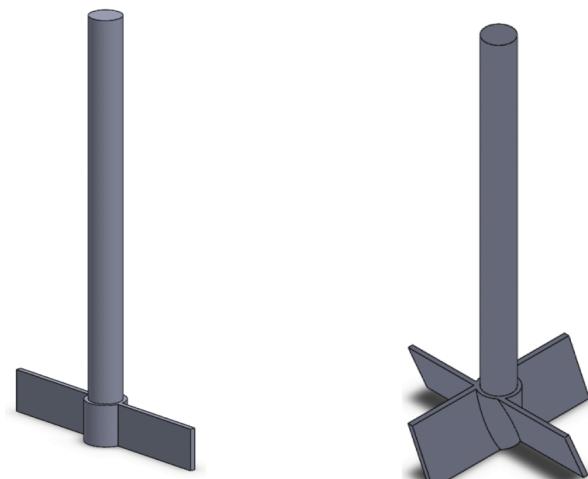
## Experimental development

Inside a mixing cylindrical tank, it is possible to distinguish three velocity components: the radial component, which acts perpendicularly to the stirrer axis; the longitudinal component, which acts parallel to the axis; and the rotational component, which acts tangentially to the circular trajectory around the axis.

Impellers can be divided into two types: those which generate currents parallel to the axis, called axial flow impellers; and those which generate currents in the radial or tangential direction, called radial flow impellers. The main impeller types for liquids are propellers, paddles, and turbines (Mc Cabe *et al.*, 2007).

In this work, three different impellers were used to generate diverse flow patterns inside the system and observe the effect of different geometries on the dissolution process. The types of conventional agitators used were:

- *Agitator 1:* This agitator is a paddle which consists of two blades (Figure 1); it mainly produces radial and tangential flow currents, as well as intense shear stresses.
- *Agitator 2:* This agitator is a turbine type which consists of four blades (Figure 1) pitched at 45°. It produces both axial and radial flow, as well as high shear stresses.
- *Agitator 3:* This agitator is also a turbine type which consists of six straight blades (Figure 1). It produces radial and tangential flow, as well as a higher shear stress than agitator 2.



**Figure 1.** Agitator 1: paddle agitator, two blades; agitator 2: turbine agitator, four inclined blades; agitator 3: turbine agitator, six disk straightblades

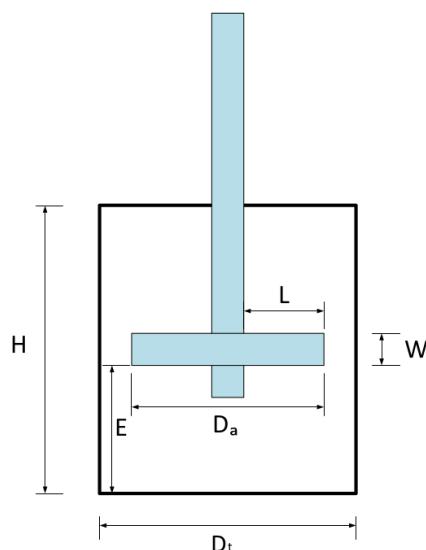
**Source:** Authors

Empirical geometric proportions of the tank and agitator, which are illustrated in Figure 2, were taken into account for the paddle-type and turbine-type agitators shown in Table 1, where  $D_a$  is the propeller diameter,  $D_t$  is the water tank diameter,  $W$  is the width of the propeller blade,  $H$  is the water tank height, and  $E$  is the height of the propeller with respect to the water tank bottom. Dimensions were measured using a caliper with an error of  $\pm$ .

**Table 1.** Ratios used between tank and agitator size

	Paddle-type	Turbine-type
$D_a/D_t$	0, 9	0, 7
$H/D_t$	0, 8	0, 8
$E/D_t$	0, 3	0, 3
$W/D_a$	0, 2	0, 3
$L/D_a$	0, 4	0, 4
$W/L$	0, 4	0, 7

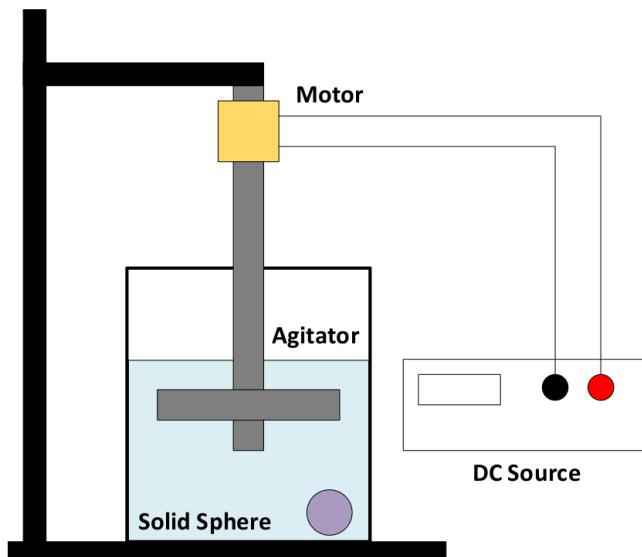
**Source:** Authors



**Figure 2.** Geometric factors considered in the blade design of the agitators

**Source:** Authors

A sketch of the experimental set-up is shown in [Figure 3](#). It consisted of a 1 000 ml beaker filled with 700 ml of water and a candy sphere (solute) immersed in it. The agitation was provided by means of an agitator coupled to a DC motor, whose angular frequency ( $\omega$ ) was controlled using a variable DC voltage source. The rotational speed was practically achieved as soon as the motor is started. Therefore, by changing  $\omega$ , modifications of the flow pattern inside the tank were obtained.



**Figure 3.** Sketch of the experimental setup

Source: Authors

The spherical mass was initially measured using a precision digital balance (OHAUS) with an accuracy of 0,01 g. Averaging ten candies, the initial mass of the candy was  $10,01 \text{ g} \pm 0,49 \text{ g}$ . Although the radius was not used, the candy spheres had an average diameter of  $17,07 \text{ mm} \pm 0,25 \text{ mm}$ . The solid was introduced into the beaker, and the solution was stirred for 2 min under a fixed voltage applied to the motor, which corresponds to a given angular velocity. Afterwards, the candy sphere was taken out and dried. Then, its mass was measured again. This process was repeated 10 times in 2 min intervals at room temperature ( $20^\circ \text{C}$ ). Time was measured using a digital chronometer with an accuracy of 0,01 s. Five variations in the stirring frequency were chosen for each agitator. The described procedure was performed for the three types of agitators illustrated in [Figure 1](#). Each measurement was repeated five times with different candies. It is important to clarify that, while the agitation occurred, there were no collisions between the candy spheres and the agitator paddles, and no changes in their spherical shape were observed. The candy was introduced and taken out with a strainer in a very short time with the agitator at rest, and it was put on a napkin so water was absorbed the mass was measured, which is why measurement errors in this regard were not significant.

As a second stage of the study, the experimental setup was placed inside a thermal bath under a controlled temperature

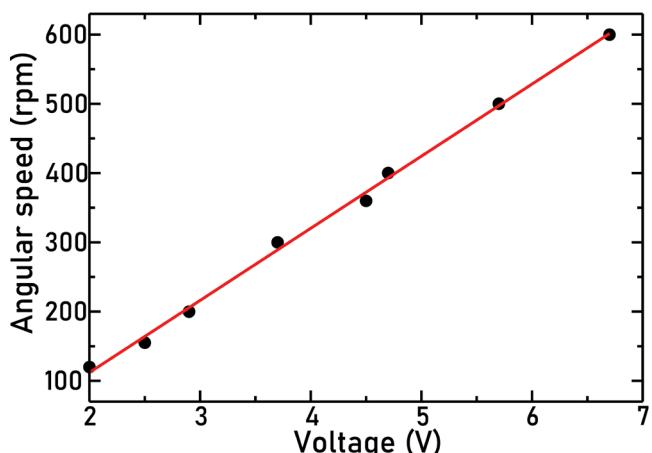
(Thermo HAAKE Phoenix II circulator). The voltage applied to the motor was kept at 7 V (632 rpm) 0,1 V, and temperature variations between 20 and  $70\ 1^\circ \text{C}$  were implemented and measured with a mercury thermometer. Because the solid dissolution rate and mass loss increased at higher temperatures, it was necessary to take mass measurements more often during this stage (they were taken every minute).

The angular velocity of the agitator was controlled by the voltage applied to the motor and measured using a tachometer with an accuracy of 0,1 rpm. The motor calibration curve is shown in [Figure 4](#), which evidences a direct relation between the voltage applied to the motor and the rotational speed. The results were very similar for the different agitators coupled to the motor because all of them have a small moment of inertia. Numerical values of the calibration curve can be consulted in the Supplementary Information section.

## Results and discussion

### Dissolution constants at room temperature

For Figures 4 to 11, dots represent the average of five measurements, and error bars are not included since they are smaller than the symbol size.

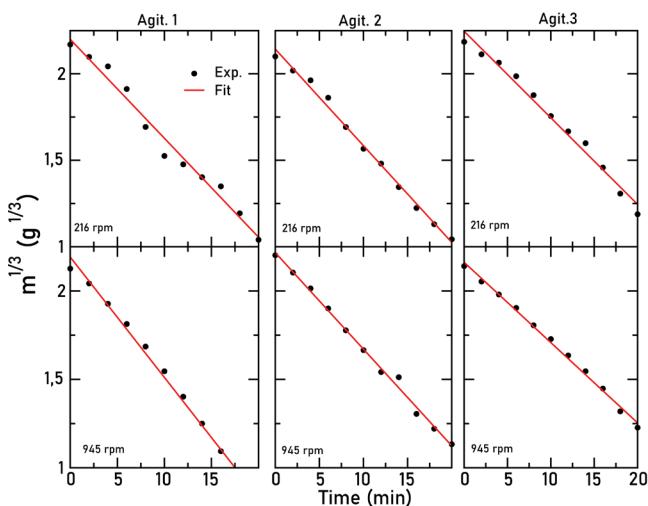


**Figure 4.** Motor calibration curve

Source: Authors

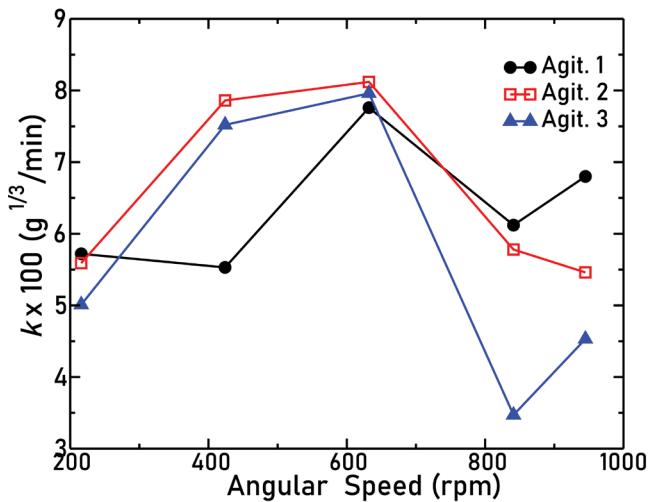
[Figure 5](#) represents the cube root of mass as a function of time and at room temperature. Dots correspond to experimental points and lines to a linear fit. Although several angular frequencies were considered, for this Figure only, the lowest and the highest angular frequencies are shown for every agitator.

It can be observed that, at all frequencies, the three types of agitators follow a linear relation between the cube root of mass and time, which agrees with the Hixson-Crowell model. From the slope of these curves, and considering [Equation \(2\)](#), it was possible to determine the value of the dissolution constant  $k$ .



**Figure 5.** Cube root solid mass as a function of time for several angular frequencies of the three agitators

Source: Authors



**Figure 6.** Solid dissolution constant as a function of the angular frequency for the three agitators

Source: Authors

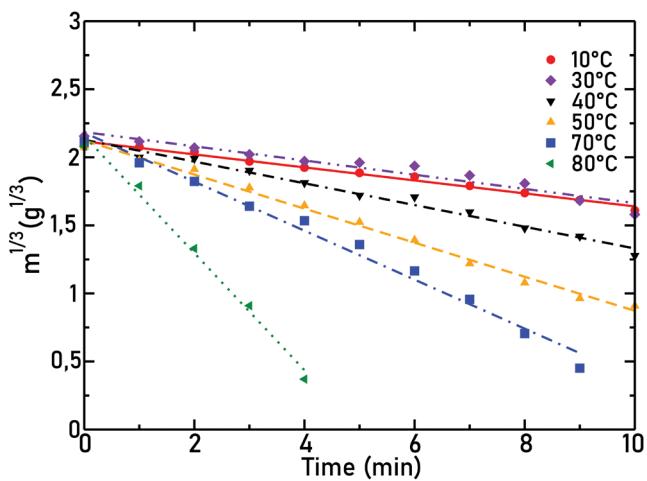
The relation between the angular frequency  $\omega$  of the agitators and the solid dissolution constant  $k$  is represented by Figure 6.

Note how, for  $\omega \leq 632$  rpm, the value of  $k$  increases. For agitator 1, the difference between the two first dissolution constant values is so small (around 3%) that we can assume that the  $k$  remained constant. However, for higher values of  $\omega$ , a decrease with respect to the maximum achieved dissolution constant is found, which is a characteristic of the three types of agitators. This behavior can be attributed to the fact that, for  $\omega > 650$  rpm, vortex and high turbulence inside the beaker were experimentally observed, especially for agitator 3, which showed the highest drop in  $k$ . The combination of these effects could make the mass transfer between solute and solvent difficult and therefore produce a decrease in mass transfer during the diffusion process. From Figure 6, it is also possible to infer how, for  $\omega \leq 632$

rpm, the higher values of dissolution constants are obtained with agitator 2, which is the one that actually has its blades pitched, thus allowing an axial flow. For  $\omega > 632$  rpm, agitator 1 attains the highest rates of dissolution, presumably since it is the one which reported the least of turbulence. This led us to infer that the dissolution constant depends indeed on the geometry used in the agitation system.

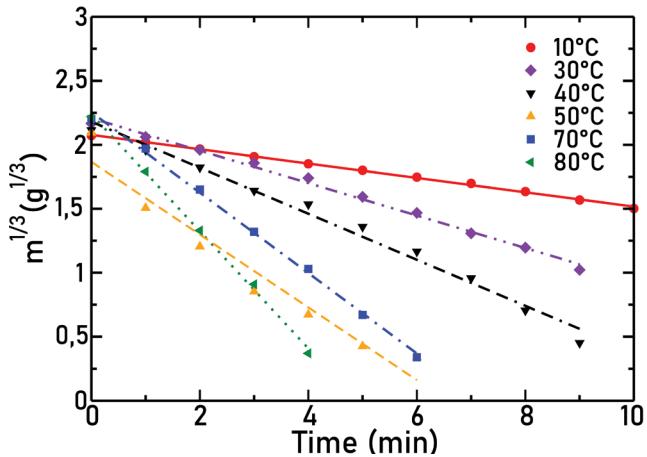
#### Dissolution constants under different temperatures

Figures 7, 8, and 9 show the cube root of mass as a function of time for agitators 1, 2, and 3, respectively, for different temperatures as indicated in each Figure. Lines correspond to the fit to the experimental points. In all cases, the angular frequency was held at 632 rpm.



**Figure 7.** Cube root of mass as a function of time for agitator 1 at different temperatures,  $\omega = 632$  rpm

Source: Authors



**Figure 8.** Cube root of mass as a function of time for agitator 2 at different temperatures,  $\omega = 632$  rpm

Source: Authors

In these Figures, it is observed how, as the temperature increases, there is an increase in the slope absolute value and therefore in the solid dissolution constant. This was expected, since, when temperature increases, there is also an increase in the solubility of particles.

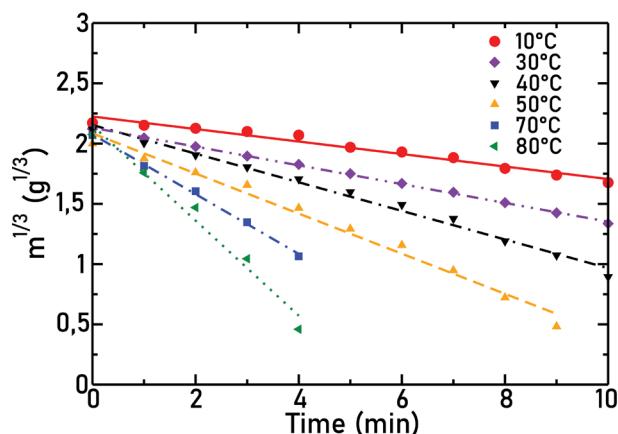


Figure 9. Cube root of mass as a function of time for agitator 3 at different temperatures,  $\omega = 632$  rpm

Source: Authors

Figure 10 shows the dependence of the solid dissolution constant as a function of temperature for the different agitators used in this experiment. As predicted by Equation (4), there is an increase in the solid dissolution constant as the temperature increases. It is also possible to observe that the dissolution constant is remarkably dependent on the type of agitator, and it is again the agitator 2 which attains the highest dissolution constants for all temperatures. For the highest temperature studied, the difference between the constants decreases, possibly due to the effect of temperature, which exerts a greater influence on the dissolution process than the agitator geometry. Still, agitator 2 keeps attaining the highest constant. Therefore, taking into consideration the results shown in Figure 6, we can say that agitator 2, which produces both axial and radial flow, also induces the highest solid dissolution constants.

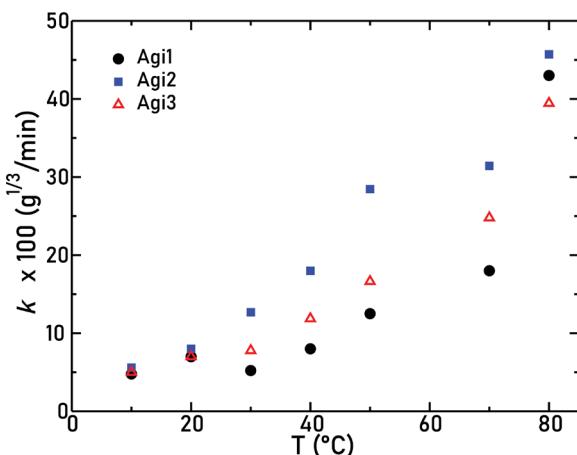


Figure 10. Variation of the dissolution constant as a function of temperature for the three types of agitators

Source: Authors

## Activation energies

Activation energies  $E_a$  for the dissolution process are determined by taking the natural logarithm of Equation (4), so that

$$\ln k = \ln A - \frac{E_a}{RT} \quad (5)$$

where  $A$  represents the pre-exponential constant.

Figure 11 shows the natural logarithm of the solid dissolution constant as a function of the inverse of the temperature for the three types of agitators used in this work. In this Figure, it is possible to observe how every straight line has approximately the same slope (Table 2). Differences between them are about 6% at most. Therefore, it is possible to conclude that the activation energy is approximately the same for all cases and is independent of the agitator geometry. This activation energy is similar to the 23 kJ/mol value reported by (Beauchamp, 2001). However, in Table 3, which summarizes the pre-exponential factors ( $A$ ), some remarkable variations appear, since line intercepts in Figure 11 are different in each situation.

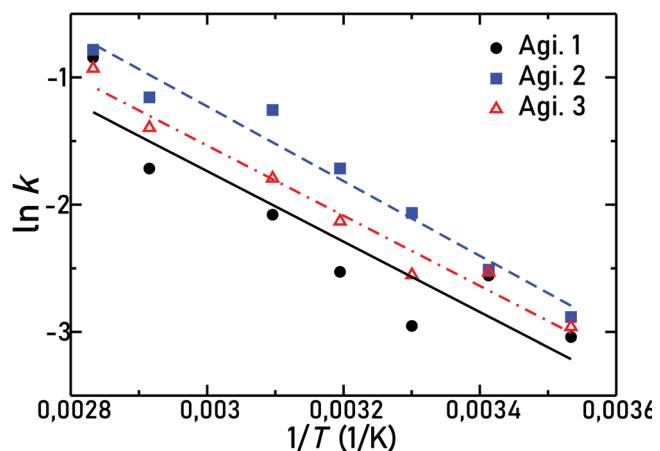


Figure 11. Relation between natural logarithm of dissolution constant, for all the agitators, and the inverse of temperature

Source: Authors

This leads us to conclude that the agitation phenomenon is modified specifically in the pre-exponential factor of the dissolution constant, while the activation energy of the process remains practically constant.

Table 2. Activation energy for the different agitators

	$E_a$ (kJ/mol)
Agi 1	$23,0 \pm 1,0$
Agi 2	$24,4 \pm 1,0$
Agi 3	$22,9 \pm 1,0$

Source: Authors

Table 3. Pre-exponential factors for the different agitators

	$A$ ( $\text{g}^{1/3}/\text{min}$ )
Agi 1	$712 \pm 14$
Agi 2	$1947 \pm 16$
Agi 3	$843 \pm 17$

Source: Authors

## Conclusions

Hixson-Crowell model appropriately predicts the kinetics of the dissolution of solids with spherical geometry. By using this model, it was possible to determine both the kinetic constants and activation energy necessary in the dissolution process. Additionally, the results showed that the agitation rate and agitator geometries affect the kinetic dissolution constant. Therefore, it is important to consider these variables when conducting a dissolution experiment.

Values of the solid dissolution constant are affected at constant temperature by the angular frequency of the agitator in all the studied geometries. The highest values in the constants were obtained for the turbine-type agitator 2, which produces axial and radial flow. However, for higher angular frequency values, the constant decreases, which can be explained in terms of the high turbulence manifested in the formation of vortices.

The activation energy was approximately the same for the three types of agitators used, regardless of their geometry. For the pre-exponential factors, a remarkable difference appeared among the agitators, which leads us to conclude that the geometry of the agitator can affect the kinetic constant specifically in the pre-exponential factor.

## References

Abreu-López, D., Amaro-Villeda, A., Acosta-González, F., González-Rivera, C., and Ramírez-Argaez, M. (2017). Effect of the impeller design on degasification kinetics using the impeller injector technique assisted by mathematical modeling. *Metals*, 7(4), 132. <https://doi.org/10.3390/met7040132>

Akrap, M., Kuzmanić, N., and Kardum, J. P. (2012). Impeller geometry effect on crystallization kinetics of borax-decahydrate in a batch cooling crystallizer. *Chemical Engineering Research and Design*, 90(6), 793-802. <https://doi.org/10.1016/j.cherd.2011.09.015>

Alok, S. (2014). Effect of different impellers and baffles on aerobic stirred tank fermenter using computational fluid dynamics. *Journal of Bioprocessing & Biotechniques*, 4(7), 1000184. <https://doi.org/10.4172/2155-9821.1000184>

Beauchamp, G. (2001). Dissolution kinetics of solids: Application with spherical candy. *Journal of Chemical Education*, 78(4), 523. <https://doi.org/10.1021/ed078p523>

Ćosić, M., Kaćunić, A., and Kuzmanić, N. (2016). The investigation of the influence of impeller blade inclination on borax nucleation and crystal growth kinetics. *Chemical Engineering Communications*, 203(11), 1497-1506. <https://doi.org/10.1080/00986445.2016.1212023>

Gu, D., Shi, X. and Liu, Z. (2020). Intensification on drawdown process of floating particles by circle package fractal impellers. *Journal of the Taiwan Institute of Chemical Engineers*, 106, 62-73. <https://doi.org/10.1016/j.jtice.2019.09.028>

Hixson, A., and Baum, S. (1942). Power requirement of turbine agitators. *Industrial and Engineering Chemistry*, 34(2), 194-208. <https://doi.org/10.1021/ie50386a013>

Hixson, A., and Crowell, J. (1931). Dependence of reaction velocity upon surface and agitation I – Theoretical considerations. *Industrial and Engineering Chemistry*, 23(8), 923-931. <https://doi.org/10.1021/ie50260a018>

Jirout, T., and Rieger, F. (2011). Impeller design for mixing of suspensions. *Chemical Engineering Research and Design*, 89(7), 1144-1151. <https://doi.org/10.1016/j.cherd.2010.12.005>

Jug, M., Hafner, A., Lovrić, J., Kregar, M. L., Pepić, I., Vanić, Ž., Cetina-Čižmek, B., and Filipović-Grčić, J. (2018). An overview of *in vitro* dissolution/release methods for novel mucosal drug delivery systems. *Journal of Pharmaceutical and Biomedical Analysis*, 147, 350-366. <https://doi.org/10.1016/j.jpba.2017.06.072>

Kazemzadeh, A., Ein-Mozaffari, F., and Lohi, A. (2020). Effect of impeller type on mixing of highly concentrated slurries of large particles. *Particuology*, 50, 88-99. <https://doi.org/10.1016/j.partic.2019.07.004>

Kravtchenko, T., Renoir, J., Parker, A., and Brigand, G. (1999). A novel method for determining the dissolution kinetics of hydrocolloid powders. *Food Hydrocolloids*, 13(3), 219-225. [https://doi.org/10.1016/S0268-005X\(99\)00002-8](https://doi.org/10.1016/S0268-005X(99)00002-8)

Legendre, D., and Zevenhoven, R. (2016). A numerical Euler-Lagrange method for bubble tower  $\text{CO}_2$  dissolution modeling. *Chemical Engineering Research and Design*, 111, 49-62. <https://doi.org/10.1016/j.cherd.2016.04.010>

Liu, Y., Guo, J., Li, W., Li, W., and Zhang, J. (2021). Investigation of gas-liquid mass transfer and power consumption characteristics in jet-flow high shear mixers. *Chemical Engineering Journal*, 411, 128580. <https://doi.org/10.1016/j.cej.2021.128580>

Mangwandi, C., Adams, M. J., Hounslow, M. J., and Salman, A. D. (2010). Effect of impeller speed on mechanical and dissolution properties of high-shear granules. *Chemical Engineering Journal*, 164(2-3), 305-315. <https://doi.org/10.1016/j.cej.2010.05.039>

McCabe, W., Smith, J., and Harriott, P. (2007). *Unit operations of chemical engineering*. McGraw-Hill.

Noyes, A., and Whitney, W. (1897). The rate of solution of solid substances in their own solutions. *Journal of the American Chemical Society*, 19(12), 930-934. <https://doi.org/10.1021/ja02086a003>

Pham, A. L. T., Sedlak, D. L., and Doyle, F. M. (2012). Dissolution of mesoporous silica supports in aqueous solutions: Implications for mesoporous silica-based water treatment processes. *Applied Catalysis B: Environmental*, 126, 258-264. <https://doi.org/10.1016/j.apcatb.2012.07.018>

Tabor, Z., and Warszyński, P. (2014). Modeling dissolution controlled release of inhibitor from a damaged coating. *Corrosion Science*, 82, 290-296. <https://doi.org/10.1016/j.corsci.2014.01.026>

Thakur, R. K., Vial, C., Djelveh, G., and Labbafi, M. (2004). Mixing of complex fluids with flat-bladed impellers: Effect of impeller geometry and highly shear-thinning behavior. *Chemical Engineering and Processing: Process Intensification*, 43(10), 1211-1222. <https://doi.org/10.1016/j.cep.2003.11.005>

**Available in:**

<https://www.redalyc.org/articulo.oa?id=64379889004>

How to cite

Complete issue

More information about this article

Journal's webpage in redalyc.org

Scientific Information System Redalyc  
Diamond Open Access scientific journal network  
Non-commercial open infrastructure owned by academia

Gustavo A. Orozco, Clara T. González-Hidalgo, Fabio Fajardo

**Influence of Agitator Geometry in the Dissolution Kinetics  
of a Spherical Solid**

**Influencia de la geometría del agitador en la cinética de  
disolución de un sólido esférico**

*Ingeniería e Investigación*

vol. 42, no. 3, e203, 2022

Facultad de Ingeniería, Universidad Nacional de Colombia.,

**ISSN:** 0120-5609

**ISSN-E:** 2248-8723

**DOI:** <https://doi.org/10.15446/ing.investig.94110>