



Revista Ambiente & Água

ISSN: 1980-993X

Instituto de Pesquisas Ambientais em Bacias Hidrográficas

Pereira, Chrystopher Allan Miranda; Brito, Núbia Natália
Integration of treatment technologies with Fenton reagent for laboratory effluent remediation
Revista Ambiente & Água, vol. 13, no. 5, 2018
Instituto de Pesquisas Ambientais em Bacias Hidrográficas

DOI: 10.4136/ambi-agua.2243

Available in: <http://www.redalyc.org/articulo.oa?id=92860438010>

- How to cite
- Complete issue
- More information about this article
- Journal's homepage in redalyc.org

UAEM redalyc.org

Scientific Information System Redalyc

Network of Scientific Journals from Latin America and the Caribbean, Spain and Portugal

Project academic non-profit, developed under the open access initiative



Integration of treatment technologies with Fenton reagent for laboratory effluent remediation

ARTICLES doi:10.4136/ambi-agua.2243

Received: 19 Feb. 2018; Accepted: 08 Jul. 2018

Chrystopher Allan Miranda Pereira; Núbia Natália Brito*

Universidade Federal de Goiás (UFG), Goiânia, GO, Brasil

Instituto de Química (IQ). E-mail: chrystophermiranda@gmail.com, nubiabrito@ufg.br

*Corresponding author

ABSTRACT

This study investigated of the potential value of the integration of the coagulation/flocculation, Advanced Oxidation Processes (AOP) (Fenton reagent) and slow sand filtration technologies, with the aim of treating laboratory wastewater. The treatment system was designed in laboratory scale through coagulation/flocculation. It involved the use of Jar Test equipment with a sequence of two rotational phases: fast mixes to 300 rpm for 20 seconds and slow mixes to 30 rpm for 6 minutes and 10 seconds, with the addition of anionic polymer and sedimentation for 60 minutes at ambient temperature. In the treatment via Fenton reagent, two rotational phases were used: rapid mixing at 300 rpm for 20 seconds with the addition of iron (Fe^{2+}) and slow mixing at 30 rpm for 6 minutes and 10 seconds with the addition of hydrogen peroxide, followed by 60 minutes of sedimentation at ambient temperature. A cylindrical tank of polyvinyl chloride, sands and non-woven synthetic fabrics were used in the slow filtration. The filtration rate adopted was $3 \text{ m}^3 \text{ m}^{-2} \text{ d}^{-1}$ with a hydraulic retention time of 264 minutes. The best concentrations of chemical reagents used in the treatments were: 0.80 mg L^{-1} of polymeric anionic, 200.00 mg L^{-1} of H_2O_2 and 13.00 mg L^{-1} of total soluble iron. The integration of the treatment technologies made it possible to achieve a removal rate of 75.27% of COD and 94.12% of total phenols. Furthermore, the conjugation of the processes allowed the removal of 87.58% of TOC.

Keywords: coagulation/flocculation, cosmetics, textile dyes.

Integração de tecnologias de tratamento com reagente de Fenton para remediação de efluente laboratorial

RESUMO

Neste trabalho a potencialidade da integração de tecnologias coagulação/floculação, processo oxidativo avançado (reação de Fenton) e filtração lenta foram estudados com o objetivo de tratar resíduo líquido laboratorial apresentando compostos da indústria cosmética e têxtil, principalmente. O sistema de tratamento foi realizado em escala laboratorial através da coagulação/floculação, o equipamento de jarros foi utilizado com uma sequência de duas fases de rotação: mistura rápida a 300 rpm por 20 segundos, mistura lenta a 30 rpm durante 6 minutos e 10 segundos com a adição de polímero aniônico e com um tempo de sedimentação de 60 minutos em temperatura ambiente. No tratamento via reagente de Fenton duas fases de rotação



foram estudadas: mistura rápida a 300 rpm por 20 segundos com a adição de ferro (Fe^{2+}), mistura lenta a 30 rpm durante 6 minutos e 10 segundos com a adição de peróxido de hidrogênio e 60 minutos de sedimentação em temperatura ambiente. Na filtração lenta foi utilizado um tanque cilíndrico de policloreto de vinila, areia e mantas sintéticas não tecidas. A taxa de filtração adotada durante os experimentos foi de $3 \text{ m}^3 \text{ m}^{-2} \text{ d}^{-1}$ com um tempo de detenção hidráulica de 264 minutos. As melhores concentrações de reagentes químicos utilizados nos tratamentos foram: $0,80 \text{ mg L}^{-1}$ de polímero aniônico, $200,00 \text{ mg L}^{-1}$ de H_2O_2 e $13,00 \text{ mg L}^{-1}$ de ferro solúvel total. A integração das tecnologias de tratamento permitiu a remoção de 75,27% de DQO e 94,12% de fenóis totais. Além disso, a conjugação dos processos permitiu remoção de COT de 87,58%.

Palavras-chave: coagulação/floculação, cosméticos, corantes têxteis.

1. INTRODUCTION

The cosmetic industry in Brazil showed a rapid growth up to 2014, compared with other industrial sectors. However, in 2015 the sector showed a decline of about 9% for the first time in the last 20 years (ABIHPEC, 2017). The cosmetics are produced and used on a massive scale for aesthetic purposes and the huge amount of drinking water used in the refrigeration system, steam generation and particularly for the washing of machines, equipment and utensils, is a cause of concern in the cosmetics industry, since it results in a significant amount of effluent for waste disposal.

According to Banerjee and collaborators (2016), the effluents of the cosmetic industry are harmful from an environmental standpoint; they include high concentrations of organic compounds of slow biodegradability, as well as the presence of anionic surfactants, grease and oils. These impede the aeration of the medium and the transfer of soluble substrates to micro-organisms that might have been able to accomplish the degradation of the organic matter and which will hence reduce the rate of the substrate conversion.

Another important industry in Brazil is the dye sector; dyes are widely used in several industrial areas, such as textiles, leather, paper, plastics, foods and rubber for the coloration of their products. Annually more than 10.000 tons of dyes are used, and approximately 100 tons are released into watercourses (Ashrafi et al., 2017).

Water contamination caused by the release of textile dyes can have a considerable impact on the ecosystem. Effluents from the textile industry are considered to be the worst pollutants of water bodies and soils. The dyes cause an undesirable color in the water when they are disposed of without treatment and lead to a reduction of solar radiation, which affects the photosynthetic activity of aquatic organisms (Semeraro et al., 2015).

In addition, the textile dyes are characterized as mutagen agents that are carcinogenic, allergic and cytotoxic and represent a risk to all forms of life (Khandare and Govindwar 2015). Some of them are not biodegradable and undergo biomagnification along the food chain.

The integration of the coagulation/flocculation, advanced oxidation (Fenton reagent) and slow filtration processes was studied in an attempt to remediate the effluents discharged by the laboratory and address the problem outlined above. The effluent mainly contained compounds from the cosmetics and textile industry

2. MATERIALS AND METHODS

2.1. Site of experiments and effluent generation

The experiments were carried out at the Wastewater Treatment Laboratory, located at the Chemistry Institute, Goiás Federal University, where the coagulation/flocculation, Fenton

reagent and slow filtration were performed. The analysis to monitor the efficiency of the treatment was carried out in the Residual Wastewater Treatment Laboratory and the Laboratory of Analytical Instrumentation IQ-UFG.

The effluent used in this research was packaged in 50 liter bottles at room temperature and included the following: a liquid by-product of other treatments performed in the Laboratory of Wastewater Treatment, a liquid mixture with chemical compounds in sunscreen, hair conditioners, textile dyes, the analysis of liquid wastewater and environmental procedures carried out by the Wastewater Treatment Laboratory team.

2.2. Experimental methodology

In evaluating the rate of efficiency, we carried out analyses in duplicate such as: absorptivity performed at maximum absorption wavelength, turbidity, total residual soluble iron, residual hydrogen peroxide, chemical oxygen demand (COD), oil and greases, pH, total organic carbon (TOC), total solids (TS), total fixed solids (TFS), total volatile solids (TVS) and settleable solids (SS). All the analyses (except residual hydrogen peroxide) complied with the Standard Methods for Examination of Water and Wastewater (SMEWW) 20th Edition (APHA, 1998).

The determination of residual hydrogen peroxide follows procedures adopted by Oliveira et al. (2001), and depend on the reaction between hydrogen peroxide and vanadate ion (VO_3^-) in acid medium. The reaction leads to the formation of peroxovanadium ion (VO_2^{3+}), which is of a reddish color and absorbs strongly at 446 nm.

The coagulation/flocculation process and the Fenton reaction were performed with jar test equipment using a one-liter volume, and varying rotations between 30 and 300 rpm with sedimentation for 60 minutes. The parameters of rotation and treatment time, as well as the range of initial concentrations of chemical reagents were used for two treatment technologies (coagulation/flocculation and Fenton's reagent) and based on the values of the parameters used in the wastewater treatment station-Goiânia-GO and in Moraes and Brito (2015).

The studies of the treatment technology through coagulation/flocculation were initiated using a factorial experimental design (2^2) that was built for experiments 2 levels: a smaller level (-) and larger one (+) for the variables: concentration of anionic polymer (0.60 and 1.40 mg L⁻¹) and pH values (6.00 and 8.00); making a total of four experiments.

The treatment with Fenton's reagent to plan a factorial design for the studies was built at 2^3 where there were 2 levels: a smaller level (-) and larger one (+) for the variables: concentration of total soluble iron (2.33 and 6.50 mg L⁻¹), pH values (3.00 and 5.00) and concentration of peroxide of hydrogen (100.00 and 200.00 mg L⁻¹); making a total of eight experiments.

The dosage of chemical reagents used for the coagulation/flocculation assays ranged from 0.00 mg L⁻¹ to 2.00 mg L⁻¹ of anionic polymer and the pH values varied from 3.00 to 10.00. Fenton's reagent treatment concentration of chemical reagents ranged from 2.33 mg L⁻¹ to 90.00 mg L⁻¹ for a total soluble iron concentration, 100.00 mg L⁻¹ to 1200.00 mg L⁻¹ for hydrogen peroxide concentration and a variation in pH values from 3.00 to 5.50.

Approximately 200.00 mL of supernatant liquid were collected from the two treatment systems after sedimentation; values of pH, turbidity, chemical oxygen demand (COD), total residual soluble iron, absorbance and residual hydrogen peroxide (in the case of H₂O₂ residual, only in treatment via Fenton's reaction) were measured to determine the efficiency of the treatment through variations in the values of these parameters during the treatment operations.

The COD determination was performed after an increase of pH to 8.00 and the sample was then heated to 80°C for 40 minutes to remove residual H₂O₂. The sample remained at rest for 2 hours (ambient temperature) to sediment residual iron that interferes with the result (Cavalcante, 2005).

A cylindrical tank made of PVC (polyvinyl chloride) with a height of 55.00 cm and a diameter of 14.50 cm was used to construct a slow filter. The collection system, which also consisted of a PVC pipe, originated from the drainage system that follows the sand layer described above. Thus it was able to avoid the occurrence of negative pressure and, hence, the accumulation of air inside the filter medium, as well as ensuring the constant presence of effluent inside the filter. The adopted rate was $3 \text{ m}^3 \text{ m}^{-2} \text{ d}^{-1}$ and this was monitored periodically by the volumetric method of analysis.

The thickness of the filter material consisted of a 20 cm layer of fine sand used in civil engineering, followed by 5 cm of nº 2 crushed stone used as a support layer. The sand was washed with water and dried in the environment with solar radiation after being placed in a slow filter.

A non-woven synthetic blanket of brand GeotêxtilGeoFort GF Model 17 was placed between the crushed stone layer and the sand layer. Its features were as follows: 95.70% porosity, specific surface of $2.530 \text{ m}^2 \text{ m}^{-3}$, weight 400 g m^{-2} , thickness of 3 mm and 100% polypropylene composition to prevent sand from permeating the crushed stone. At the top of the sand layer, two non-woven synthetic blankets (Model 17) were used to prolong the effective time of the filter media and make the “*schmutzdecke*” development more efficient. The treatment began on June 13, 2016 at 1:00 p.m. and the treatment lasted for 264 minutes.

The slow filter was operated with a sample of effluent, batch processing system. The analyses were only employed for the monitoring of the treatment efficiency after some preliminary tests had been carried out. This was accomplished so that the filter could acquire a period of biological development and maturation.

The stages of the research treatment of the technologies tested are separated in the way shown in Figure 1 which represents the crude laboratory effluent involving a primary treatment through coagulation/flocculation, and subsequently a secondary treatment with Fenton’s reagent and concluding with the slow filtration to provide final polishing.

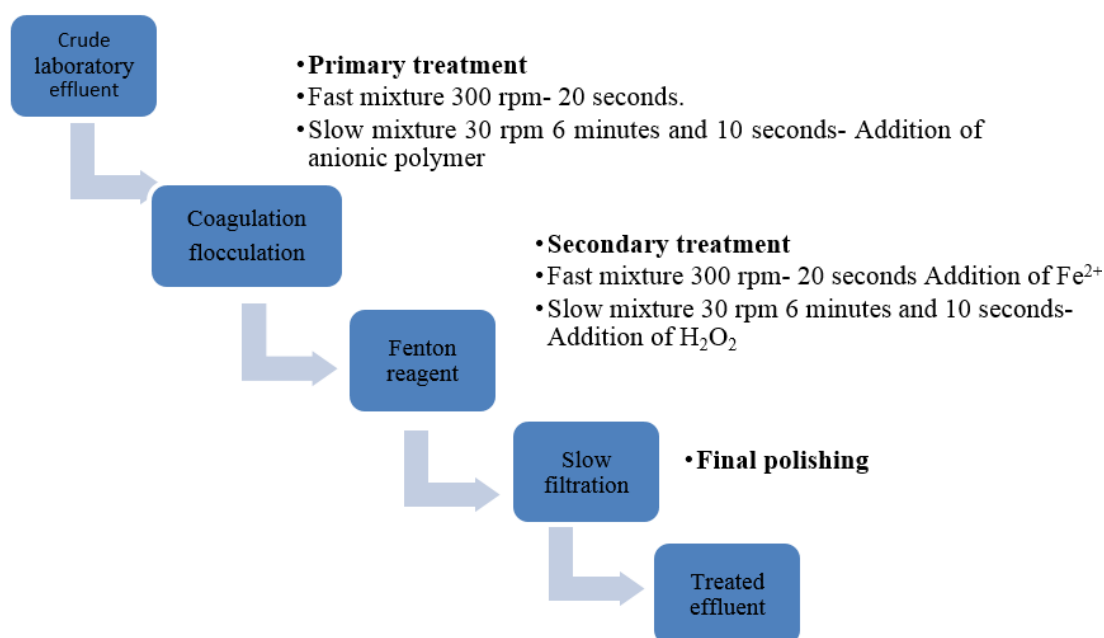


Figure 1. Treatment flowchart: Coagulation/flocculation, Fenton reagent and slow sand filtration.

As already explained, the studies began with the treatment technology through coagulation/flocculation so that a factorial experimental design (2^2) planned for the variables:

concentration of anionic polymer (0.60 and 1.40 mg L⁻¹) and pH values (6.00 and 8.00); subsequently the dosage of chemical reagents for the coagulation/flocculation assays ranged from 0.00 mg L⁻¹ to 2.00 mg L⁻¹ of anionic polymer and the pH values varied from 3.00 to 10.00.

Before starting the treatment with Fenton's reagent for the study, a factorial experimental design (2³) was planned for the variables: concentration of total soluble iron (2.33 and 6.50 mg L⁻¹), pH values (3.00 and 5.00) and concentration of peroxide of hydrogen (100.00 and 200.00 mg L⁻¹); subsequently the concentration of chemical reagents ranged from 2.33 mg L⁻¹ to 90.00 mg L⁻¹ for total soluble iron concentration, 100.00 mg L⁻¹ to 1200.00 mg L⁻¹ for hydrogen peroxide concentration and the pH values varied from 3.00 to 5.50.

The total treatment time was 400 minutes for the integration of the technologies. At the end of the combined treatment stages, the treatment efficiency was monitored through an environmental control system and further studies on classification of treated effluent for disposal and non-potable reuse purposes were carried out in compliance with the following resolutions: Conama (2005;2011), AGR (2009), ABNT (1997) and USEPA (2012).

2.3. Characterization of crude effluent

A characterization of effluent was carried out, which provided average and standard deviations from the calculations made with five different samples that are illustrated in Table 1.

Table 1. Average values and standard deviation of the analyzed physicochemical parameters for crude laboratory liquid residue.

Parameters	Values	Crude laboratory effluent	Resolution Conama (2005; 2011); AGR (2009)
pH	Average	1.73	5-9
	Standard deviation	0.06	
Turbidity (NTU)	Average	134.40	100 NTU
	Standard deviation	24.96	
Absorbance (320nm)	Average	1.60	450.00 mg O ₂ L ⁻¹
	Standard deviation	0.40	
COD (mg O₂ L⁻¹)	Average	1400.49	15.00 mg L ⁻¹
	Standard deviation	436.53	
Soluble Total Residual Iron (mg L⁻¹)	Average	82.49	
	Standard deviation	3.34	

The main parameter that needs to be monitored is the removal of the organic matter represented in this study by the COD. The Brazilian Resolution of 2011 (Conama, 2011) does not specify a COD value for effluent discharge, but only for BOD with a minimum removal of 60%. The Brazilian Resolution of 2009 (AGR, 2009) stipulates a COD maximum value of 450.00 mg O₂ L⁻¹ for wastewater discharge. In this article, crude laboratory effluent had a COD value of, on average, 1400.49 mg O₂ L⁻¹, which is also outside the standards for hydraulic bodies.

In light of the data, we believe that the integration of the proposed technologies gives a promising opportunity for the treatment of recalcitrant effluents.

3. RESULTS AND DISCUSSION

3.1. Coagulation/Flocculation treatment

The most interesting results obtained in the preliminary tests (2² factorial planning) were anionic polymer concentration 0.60 mg L⁻¹ and pH value 6.00 (31.16% COD removal), where the negative charges of the effective particles which led to chemical interactions were

neutralized, thus causing the formation of flake and incurring a lower cost, compared with the high cost of the chemical reagents involved in the process.

After making a choice of these data, the variation of anionic polymer concentration was studied, with pH value fixed at 6.00 and concentration of total soluble iron 82.49 mg L^{-1} (natural concentrations of effluent) to obtain maximum efficiency in treatment, as can be seen in Figure 2.

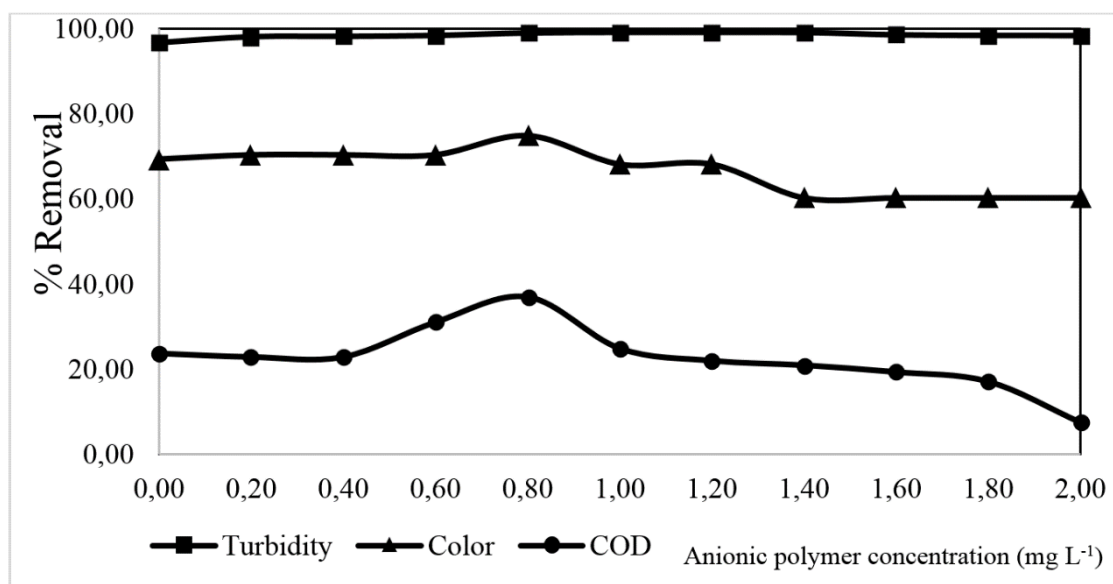


Figure 2. Effect of variation of anionic polymer concentration on COD, turbidity and absorbance removals by coagulation/flocculation treatment (pH 6.00 and residual total soluble iron = 82.49 mg L^{-1}).

It can be observed in Figure 2 that COD removal remained constant at 23.00% up to 0.40 mg L^{-1} , and it did not show any significant results in this range with the use of anionic polymer.

The removal of the organic matter increased around 11% when there was a concentration of polymeric anionic of 0.80 mg L^{-1} , which is a considerable value for treating recalcitrant compounds.

It is worth noting that this dosage of polymeric anionic also contributes significantly to the cohesion of the flake, since a better interaction of the iron ions prevents them from being released into the water bodies (Shaikh et al., 2017). An example of this was the removal of iron (in these conditions) of 97.17%. The initial iron concentration was 82.49 mg L^{-1} and the final concentration was 2.33 mg L^{-1} . This is because the polymer chains are able to adopt multiple conformations on the adsorbent surface, which distinguishes their adsorption behavior from that of simple molecules and ions (Wiśniewska et al., 2016). In this way, the soluble iron concentration also reaches parameters for disposal that comply with Brazil Resolution 2011 (Conama, 2011) which provides a maximum of 15 mg L^{-1} of total iron for disposal in hydric bodies.

It was found that when there was a concentration above 1.00 mg L^{-1} , it led to a reduction in the percentage of COD removal efficiency observed, which proves that an excess of anionic polymer led to an increase of solids, especially colloidal and particulate material. The excess of anionic polymers re-stabilizes the charge on colloidal particles by saturating its surface, damaging the formation of flocks in flocculation process and increasing the concentration of particulate matter in the effluent (Quartaroli et al., 2014).

During the removal of the color as a result of absorbance, measured at the maximum absorption wavelength (spectral sweep) 320 nm, it was noted that without the addition of an anionic polymer, a removal of 70.51% was obtained. With the addition of the anionic polymer, the maximum rate of efficiency for the removal of this parameter was 75.00%, with 0.80 mg L⁻¹ of anionic polymer. A decrease in the absorbance removal rate of efficiency of 1.00 mg L⁻¹ resulted in an excess of anionic polymer which led to an increase of solids interfering with absorption of the medium, and hence color parameter (Quartaroliet al., 2014).

With regard to the analysis of turbidity, it was observed that the turbidity remained almost constant with a mean rate of removal for all concentrations of anionic polymer. It should be noted that without the addition of the anionic polymer, there is a removal efficiency rate of around 96.80%, which is evidence that anionic polymer does not interfere in the removal of turbidity. It is only the presence of total soluble iron in residue and the suitable adjustment of pH that are sufficient for the removal of suspended and colloidal particles that cause turbidity in this effluent.

This fact can be explained by the formation of ferrous and ferric hydroxides of positive residual charge. These hydroxides, (also called gel), destabilize the colloidal particles of the effluent that possess a negative load, by reducing the potential Zeta and allow the approach and clustering of these particles, which characterizes the coagulation/flocculation process. It is also important to point out that there is a displacement of the chemical equilibrium when the pH increases to 6.00, which allows the formation of hydroxylated species, as is illustrated in Equation 1 (Morais and Brito 2016; Shaikh et al., 2017).



In view of the obtained data, it was confirmed that the optimum concentration of anionic polymer was 0.80 mg L⁻¹, and an attempt was made to evaluate the influence of pH values on the efficiency of the treatment. At this stage, the pH values were found to be in the range of 3.00 to 10.00 with total soluble iron concentration of 82.49 mg L⁻¹ and anionic polymer of 0.80 mg L⁻¹ (Figure 3).

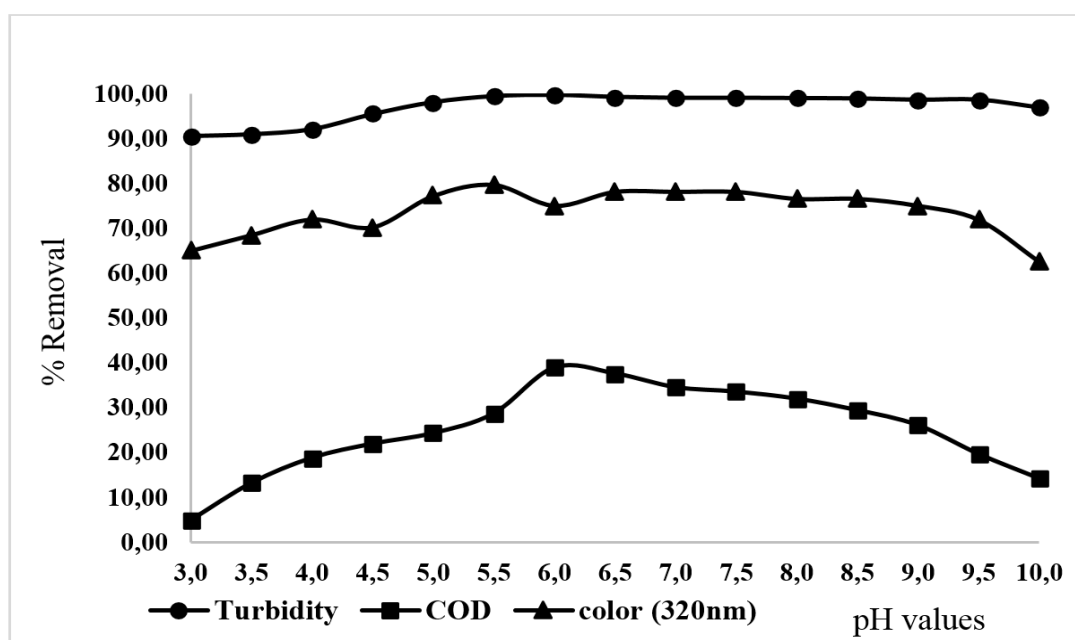


Figure 3. The effect of the variation of the pH values on COD, turbidity and absorbance removals for the treatment with coagulation/flocculation (total soluble iron =82.49 mg L⁻¹; anionic polymer =0.80 mg L⁻¹).

In the case of the analysis of COD, the best percentage for removal (38.86%) occurred at pH 6.00, because the polyacrylamide chains were completely ionized. As a result, there was an interaction of the anionic polymer with surface of flake mainly through the formation of hydrogen bonds between them (Wiśniewska et al., 2016).

The interactions between adsorbed species and surface hydroxyl groups, (whether protonated or not), altered the resulting charge of the particles and hence, the stability of suspension conditions. When pH of the solution is greater than the zero point charge of the solid in question (hydroxylated iron species), the surface of the solution becomes negative. This discourages electrostatic interactions with polyacrylamide polymer chains (since they contain completely dissociated carboxyl groups); this phenomenon may be occurring in this study when pH is above 6.00 (Wiśniewska et al., 2016).

With regard to the absorbance analysis at 320 nm, it can be seen that the removal efficiency at pH 6.00 was 75.00% on average, with removals up to 78.00% (pH values between 5.00-7.50) which can also be visualized, although when analyzing a set of three parameters, the choice of the best pH value was 6.00. At pH less than 4.50, there is a decrease in removal efficiency, with a removal efficiency of 64.91%.

When this occurs, low pH values (less than 5.00), and aquo-complexes $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ are predominant, while the hydroxylated $\text{Fe}(\text{OH})^{2+}$ species is generally at a pH greater than 4.00 (Gama, 2012). The $\text{Fe}(\text{OH})^{2+}$ species displays maximum absorbance with wavelengths of 300 nm, extending up to approximately 400 nm (Nogueira et al., 2007).

In the case of the turbidity analysis, it can be observed in Figure 3 that the percentages or removals were similar to the pH values studied, with results of reduction around 99.00%. A maximum removal rate at pH 6.00 was obtained of around 99.72%.

When there is pH of 3.00 to 5.50, turbidity has a smaller point of removal. It has been shown that at higher pH values (> 5.50), a displacement of the hydrolysis equilibrium of iron ions is achieved, leading to the formation of iron hydroxides that are essential for the interaction with colloids and formation of flakes (Guo et al., 2010).

A better operating condition was obtained from the results of the pH variation, where 82.49 mg L^{-1} of total soluble iron and 0.80 mg L^{-1} of anionic polymer were used in a pH value of 6.00; this made it possible to reach COD removals of 38.86%, turbidity of 99.72% and absorbance of 75.00%.

After the coagulation/flocculation treatment and the results were obtained for pH value, the turbidity and total iron were in accordance with the effluent release standards established by Brazil 2011 (Conama, 2011), with the exception of the COD parameter ($856.26 \text{ mg O}_2 \text{ L}^{-1}$) that had values above those allowed by the Brazil 2009 Resolution (AGR, 2009), at $450.00 \text{ mg O}_2 \text{ L}^{-1}$.

Thus, treatment must be provided via the Fenton reagent that can be combined with the coagulation/flocculation process, mainly to reduce the values related to chemical oxygen demand, that even after treatment by coagulation/flocculation still fail to meet the current standards for the discharge of effluents in water bodies, and hence would not be suitable for reuse for non-potable purposes.

3.2. Fenton reagent treatment

After the coagulation/flocculation treatment, the best test obtained from the Fenton reagent assays, was used to meet the standards and established values laid down by the studied resolutions. The residual total soluble iron 2.33 mg L^{-1} obtained from the previous treatment (coagulation / flocculation) was used for the factorial design, which avoided the unnecessary use of chemical reagents.

Preliminary tests were conducted using a 2^3 factorial planning design (two levels and three variables), since this had the highest removal efficiency for COD (29.82%), and is the most important parameter, given the difficulty of its removal and its estimated representation of organic matter. In the case of the other parameters, the results were turbidity (65.00%), absorbance (40.00%) and final pH of 2.83. In this experiment, 6.50 mg L^{-1} of total soluble iron and 200.00 mg L^{-1} of H_2O_2 were used at an initial pH of 3.00.

From the data produced by the factorial design, the total soluble iron concentration, H_2O_2 and pH values were varied to determine the best conditions for the Fenton reaction in the treatment of laboratory effluent after coagulation/flocculation, starting with a concentration of total soluble iron (2.33 to 90.00 mg L^{-1}) maintaining H_2O_2 at 200.00 mg L^{-1} and pH at 3.00.

In Figure 4, it was clear that the COD removal values had maximum values in 13.00 mg L^{-1} concentration of total soluble iron (COD 32.37%). When concentrations below 13.00 mg L^{-1} of total soluble iron were analyzed, there was a loss of efficiency, caused by the low concentration of iron species capable of reacting with H_2O_2 , and this reduced the generation of hydroxyl radicals in the medium.

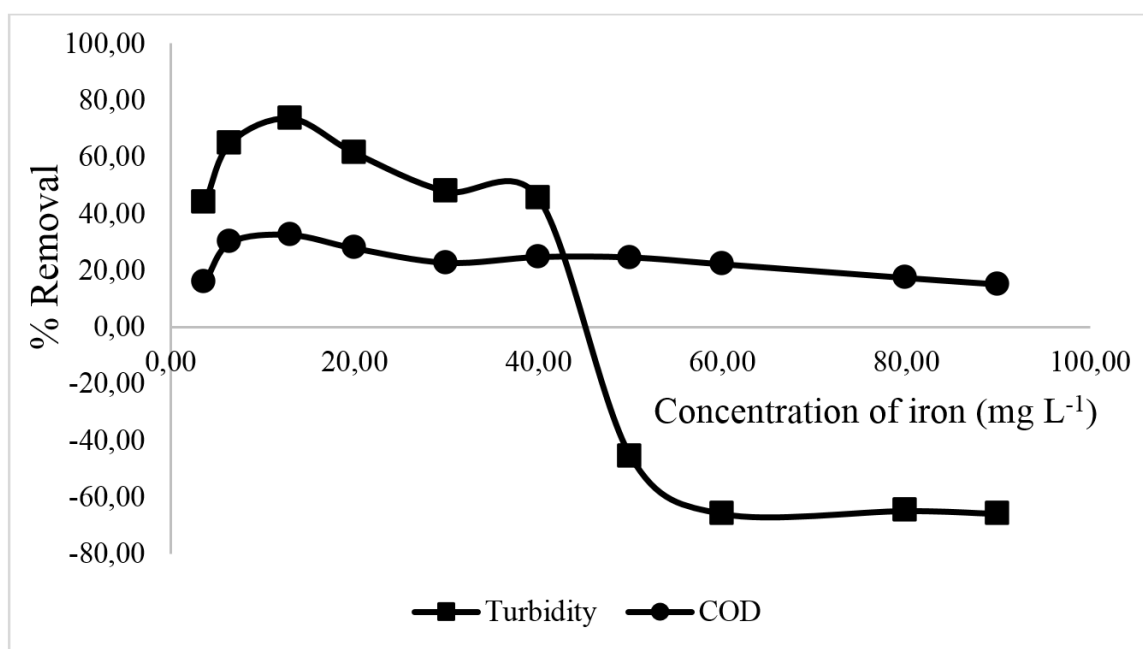
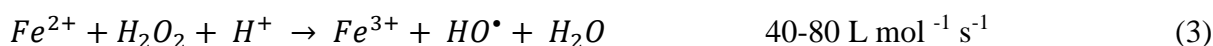


Figure 4. Effect of total soluble iron concentration on COD and turbidity removals through treatment with the Fenton reagent ($\text{H}_2\text{O}_2 = 200.00 \text{ mg L}^{-1}$, pH 3.00). Note: the negative (-) signals represent an increase in value of the parameter in question.

When going beyond a concentration of 13.00 mg L^{-1} , there begins to be a decline in efficiency in the removal of organic matter. This is due to an increase in concentration of Fe^{+2} that sequesters hydroxyl radicals, producing Fe^{3+} and OH^- ions, as expressed in Equation 2 (Ma and Xia, 2009).



Because the kinetic constant ($2.5 \cdot 10^8 \text{ L mol}^{-1} \text{ s}^{-1}$) is higher than the Fenton reaction Phase 1 reaction (Equation 2), the excess of Fe^{+2} in the system prefers the reaction represented by Equation 1 to the reaction of Equation 3, which confirms the negative result of the high increase in the concentration of Fe^{+2} in the treatment.



In an analysis of turbidity, it can be seen that an excessive increase of concentration of Fe^{+2} leads to a decrease in efficiency with regard to the removal of suspended/colloidal particles; this is shown in the graph of Figure 4, where there is a concentration of around 50.00 mg L^{-1} of total soluble iron. This behavior can be explained by the oxidation of Fe^{+2} to Fe^{+3} which, by hydrolysis in aqueous solution, leads to the formation of iron and ferric hydroxides, which cause increased turbidity in the treated effluent, as shown in Equations 4, 5 and 6 (Morais and Brito, 2016).

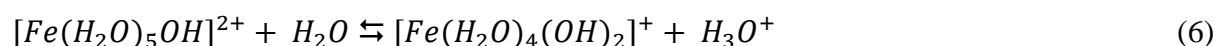
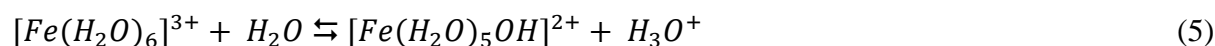


Figure 4 shows that keeping the concentration of H_2O_2 at 200.00 mg L^{-1} and pH 3.00 with a concentration of total soluble iron at 13.00 mg L^{-1} enabled the removal of COD to reach around 32.37%, with turbidity of 73.53% and final pH of 2.95. At the end of the treatment, the residual total soluble iron concentration was found to be 6.40 mg L^{-1} , which is within the release standards established by Brazil (Conama, 2011) (i.e., 15.00 mg L^{-1}).

In the next stage of the treatment, the influence of pH values was evaluated, and values of total soluble iron at 13.00 mg L^{-1} and H_2O_2 at 200.00 mg L^{-1} (Figure 5) were maintained.

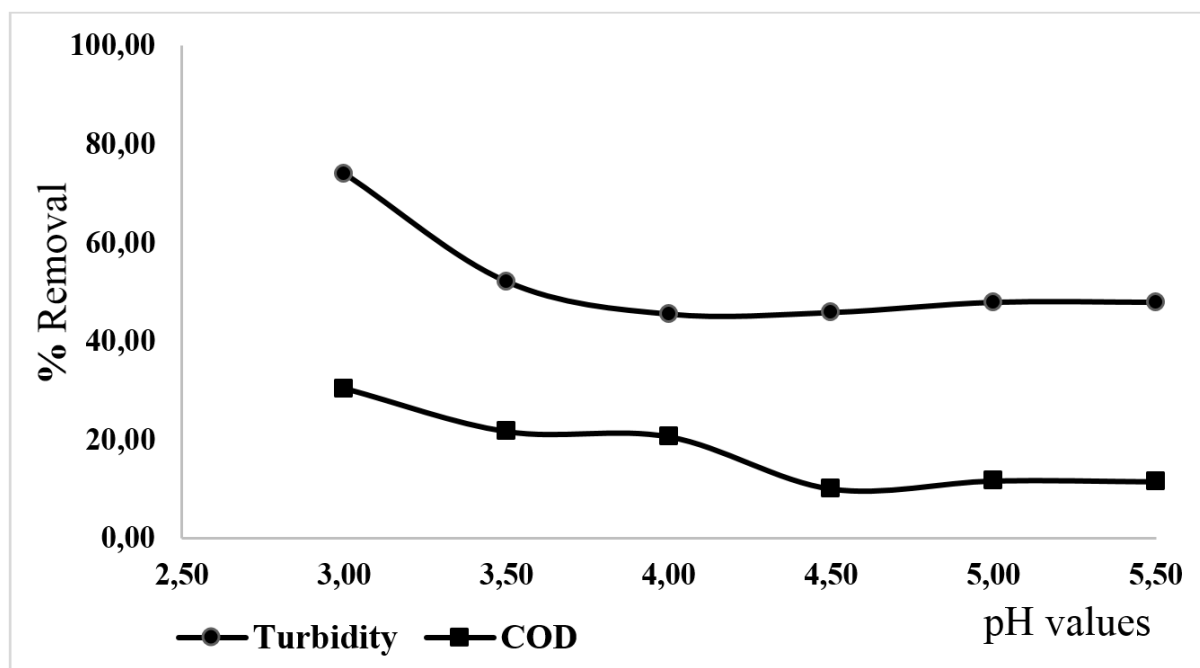


Figure 5. Effect of the variation of the pH values on COD and turbidity removals for the treatment with Fenton reagent (total soluble iron = 13.00 mg L^{-1} ; $\text{H}_2\text{O}_2 = 200.00 \text{ mg L}^{-1}$).

The efficiency removal rates of COD values (30.40%) and turbidity (74.00%), showed maximum values at pH 3.00. The treatment has a higher degree of efficiency in the acid range between pH 3-4, by directly increasing the concentration of Fe^{2+} in solution and the availability of the hydroxyl radicals in the medium (Lin et al., 2016).

The efficiency of Fenton reagent depends on the formation of the hydroxyl radicals shown in Equation 3 and the increase of H^+ ions induces the consumption of OH^- ions by undergoing a displacement in the equation for the formation of hydroxyl radicals (Mirzaei et al., 2017).

When there was a pH value higher than 4.50, a decrease in COD removal percentage was observed, owing to the decrease of free iron species, together with formation of complex species and iron precipitates, as well as the degradation of H_2O_2 to O_2 and H_2O or the formation of anion HO_2^- (Torradesa and Montañob, 2014). In conditions when there are neutral pH values, the reduction of H^+ ions is induced, which reduces the productive efficiency of the hydroxyl radical.

According to Nogueira et al. (2007), when the pH value is smaller than 3.00, the rate of degradation also decreases, although iron species remains soluble, since high concentrations of H^+ can sequester hydroxyl radicals, as shown in Equation 7:



In the last stage of Fenton's reagent, different concentrations of H_2O_2 were observed with constant values of total soluble iron of 13.00 mg L^{-1} and pH 3.00. The results for the variations of hydrogen peroxide are shown in Figure 6.

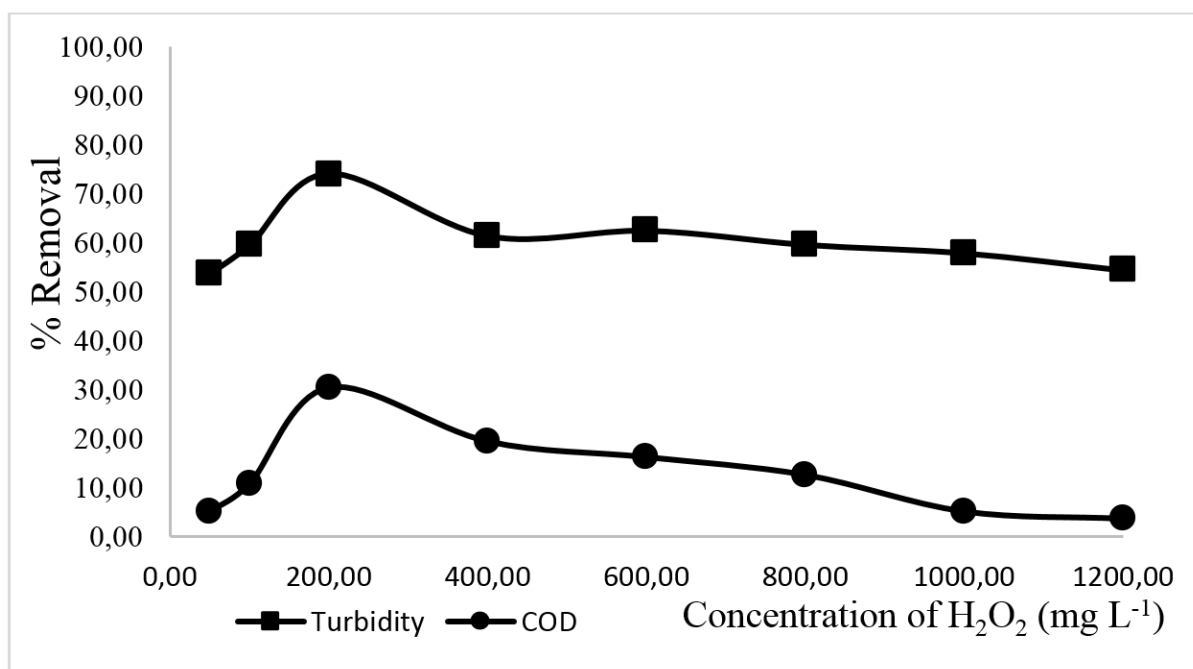


Figure 6. Effect of H_2O_2 concentration on COD and turbidity removals through treatment with Fenton reagent (Total soluble iron = 13.00 mg L^{-1} , pH 3.00).

In the analysis of Figure 6, it was confirmed that concentrations of 200.00 mg L^{-1} of H_2O_2 increase the removal of COD and turbidity; COD has a reduction efficiency of 30.40% and turbidity of 74.00%. Values greater than 200.00 mg L^{-1} reduce the removal efficiency, and have a negative effect on a high concentration of hydrogen peroxide, which reacts with hydroxyl radical and forms a hydroperoxyl radical (Equation 8) (Brito et al., 2010).



In addition to reacting to hydrogen peroxide, as is shown in Equation 9, hydroxyl radicals can react with each other, by regenerating hydrogen peroxide (eq 9), and reducing concentrations of the radical in the medium and reducing the efficiency of the Fenton reaction (Brito et al., 2010).



Ragasson (2013) specifies that the typical ratio of $[Fe^{+2}]$ and $[H_2O_2]$ is in the range of 1:5 to 1:10 for greater efficiency of Fenton reagent; however, this ratio depends on the nature of the effluent and its organic load.

After the reagent Fenton treatment for laboratory effluent, the relation between Fe^{2+} ions and hydrogen peroxide was 1:15.38.

These concentrations was varied to determine the ideal concentrations of iron and hydrogen peroxide, which are always around the ratio 1:15 of soluble iron and hydrogen peroxide (the best ratio that has been found in this article), as shown in Figure 7.

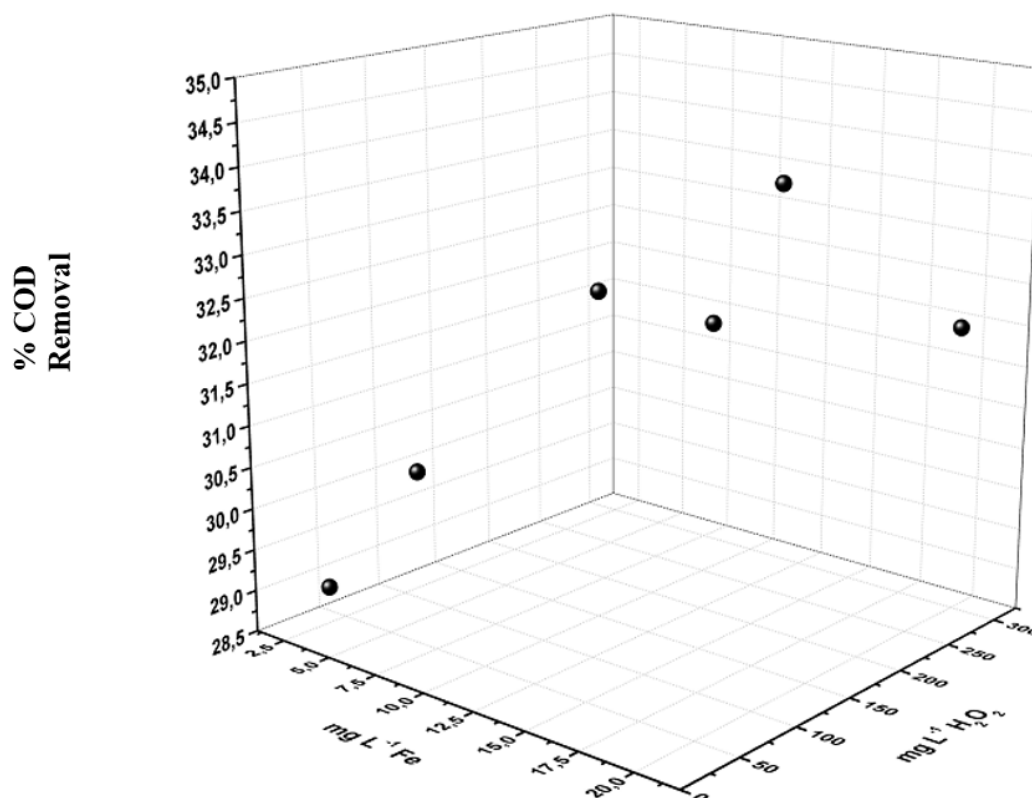


Figure 7. Study of Fe: H_2O_2 ratio in COD removal using the ratio obtained (1:15).

From the perspective of the Fe: H_2O_2 (ratio of 1:15) COD removals ranged from 28.94% to 33.70%. It was decided to keep total soluble iron concentrations of 13.00 mg L^{-1} and H_2O_2 of 200.00 mg L^{-1} with a removal percentage of around 32.05% COD. At concentrations of 15.00 mg L^{-1} of total soluble iron and H_2O_2 of 225.00 mg L^{-1} , a COD removal of 33.70% was achieved; however, the increase of only 2.00% does not compensate for the extra expense of the chemical reagents.

The selected concentration of 13.00 mg L^{-1} of total soluble iron may indicate whether there was removal of iron at the end of the slow filtration, by taking into account the assimilation of iron and manganese by the microbial consortium formed at the top of the layer of the filter bed called “schmutzdecke” and the adsorption of metals in microbial beds and from the formed biofilm in the sand layers (Arantes et al., 2014).

On the basis of the study of the variables that had a direct influence on Fenton's reagent treatment, it could be concluded what was the optimal concentration in removal of both COD (30.40%) and turbidity (74.00%) - using 200.00 mg L^{-1} of hydrogen peroxide and 13.00 mg L^{-1} of total soluble iron. With regard to pH, its greater efficiency was obtained at a

3.00 value. The final pH value after treatment was 2.95. The residual iron had a value of 6.67 mg L⁻¹ and residual H₂O₂ of 0.42 mg L⁻¹.

After the treatments, there continued to be a problem concerning the discharge of the organic matter, which reached a COD value of 595.95 mg O₂ L⁻¹, and thus failed to reach the minimum value (450.00 mgO₂ L⁻¹) stipulated by the Brazilian Resolution of 2009 (AGR, 2009). For this reason, the use of combined slow filtration with coagulation/flocculation processes was recommended together with the Fenton reagent discussed above. This was used for final polishing with the aim of meeting current standards of effluent release in water bodies and, perhaps later, the end of reuse for non-potable purposes.

3.3. Slow Filtration Treatment

After the treatment by Fenton reaction, the pH was raised to 7.00. The pH around 7.00 favors the growth of microorganisms that are essential for the development of the biological layer (“*schmutzdecke*”) at the top of sand layer and benefits from a good performance of the slow filter (Nascimento et al., 2012).

The main parameter, which is still failing to comply with federal legislation regarding COD, reached a parameter for disposal ranging from 595.95 mg O₂ L⁻¹ to 346.21 mg O₂ L⁻¹ in the best treatment batch.

After the fourth treatment batch, the slow filter starts the clogging process and it is necessary to clean the non-woven synthetic blankets on top of the sand filter layer before starting treatment again.

After the integration of technologies in these studies, the following analyses were performed before and after proposed treatments and properly discussed as follows.

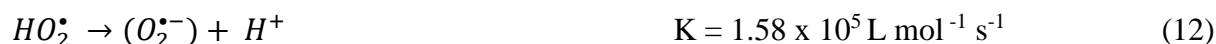
3.4. Analysis of the Parameters Investigated

In the study of pH values, it was observed that after the treatment by coagulation/flocculation, there is a reduction of the pH value which achieves a final value of 5.81. This is due to hydrolysis of iron which releases protons (H⁺) during the formation of hydroxylating species of iron in the reaction medium (Marcionilio et al., 2015).

In the treatment of the Fenton reagent, the initial pH value was 3.00, but at the end of the process, this value changed to 2.95. This reduction is due to the release of H⁺ ions during the regeneration process of the catalyst represented in Equations 10 and 11 (Sohrabi et al., 2017).



The decomposition of the hydroperoxyl radical also triggers the release of protons in medium (Equation 12) (Giannakis et al., 2016).



The reaction between ferric ion and the hydroperoxyl radical also leads to the regeneration of ferrous ions with a release of oxygen gas and H⁺ ions, as expressed in Equation 13 (Babuponnusami and Muthukumar, 2014).



At the end of the integration of the treatment technologies, the final pH value was 6.93, which is within the allowed range for disposal (between 5-9) according to Brazilian legislation (Pompei et al., 2017).

With regard to the absorbance analysis, it should be noted that at a wavelength of 320 nm researchers have found the presence of aromatic rings in solution. In addition, it is reported that at wavelengths of 200 to 300 nm, the main structures are carbonyl, carboxyl, nitroso and nitrite. At 292 nm, the most likely compounds are: aromatic amino acids and highly conjugated aliphatic chain volatile compounds, at 328 nm naphthalene and its derivatives (Hu and Wang, 2017).

Colored organic compounds have a chromophore group attached to a conjugated double bonds system. These chromophoric groups have unsaturated bonds such as $C=O$ and $N=N$ (Durigan et al., 2012).

Within this scenario for absorbance analysis, it was noted that during the post-reaction Fenton treatment, there was an increase of 65.71% of absorbance at 200 nm, possibly owing to the generation of intermediates that are strongly absorbent at this wavelength. According to Hu and Wang (2017), at 200-250 nm absorbance, the values may represent carboxyl functional groups that are strongly absorbent in this range.

The overall final removals of absorbance in this study were 83.50% at 200 nm and 86.25% at 320 nm for 400 minutes of treatment.

With regard to the turbidity analysis, the crude effluent showed a value of 134.40 NTU, and the final value obtained was 0.78 NTU in accordance with the limits allowed by Brazil 2005 (Conama, 2005), which establishes a value below 40.00 NTU in Class 1 waters. After crude laboratory effluent is treated by the coagulation/flocculation process, it already acquires a value of 4.1 NTU, which is within the limits imposed by federal legislation. After integrating three technologies, it was possible to obtain 99.41% of the removal value of this analytical parameter.

In the analysis of total soluble iron, the treated laboratory effluent had a total residual soluble iron concentration of 0.56 mg L^{-1} , after combined treatments. Compared with the crude laboratory effluent, this value obtained a reduction of 99.32% with regard to the initial concentration of total soluble iron (82.49 mg L^{-1}).

In the same way as in the first turbidity stage of coagulation/flocculation treatment, the greatest removal value of this parameter (iron) was 96.19%. This was due to the formation of iron hydroxylating species that interacts with colloids of medium to form cohesive flakes with good sedimentability; the action of the anionic polymer is of considerable importance for the interaction of the Fe^{2+}/Fe^{3+} ions and reducing their availability to aqueous medium (Ma and Xia, 2009).

The slow filtration process is also responsible for the removal of iron that enters in the treatment system with a concentration of 7.54 mg L^{-1} and has a final value of 0.56 mg L^{-1} . The processes of straining, sedimentation, inertial impaction, interception, adhesion, flocculation, diffusion, adsorption and biological activity (a microbial consortium called "*schmutzdecke*", which is described above and which consumes iron and manganese) have been suggested as mechanisms that can be employed for contaminant removal in slow filtration (Bagundol et al., 2013). In this way, soluble iron concentration also reaches parameters for disposal that comply with Resolution Brazil 2011 (Conama, 2011) which provides a maximum of 15 mg L^{-1} of total iron for disposal in hidric bodies.

The Residual hydrogen peroxide in crude laboratory effluent had a concentration of 1.13 mg L^{-1} . During the second stage, the Fenton reaction added 200.00 mg L^{-1} of H_2O_2 where almost all of the H_2O_2 was consumed during the homogeneous catalytic process (phase 1 of Fenton Equation 3) with a final concentration after the Fenton reaction of 0.41 mg L^{-1} and overall removal with combination treatments of 92.92%.

The removal of this value is extremely important because hydrogen peroxide is an oxidizing agent (1.42 V) that can continue the process of oxidation in aqueous medium if it

does not decompose rapidly in water and molecular oxygen (Brito et al., 2015). In addition, the high presence of H_2O_2 prevents a visualization of spectral changes in wavelengths between 200-300 nm, although in this article it was possible to determine the decrease of signals in these regions, which may indicate a degradation of aromatic intermediates (Salvador et al., 2012).

Through the analysis of the Chemical Oxygen Demand, it was possible to achieve an overall removal of 75.27% with a final COD of $346.21 \text{ mg O}_2 \text{ L}^{-1}$, thus obtaining a parameter for disposal that is in compliance with Brazilian Resolution 2009 (AGR, 2009), which recommends a final disposal of COD in the order of $450.00 \text{ mg O}_2 \text{ L}^{-1}$. The Fenton reaction, which is technology that involves the oxidation of organic matter, contributed to 30.40% of the removal. Fenton's reagent is an extremely promising remediation technology because it has a relatively low cost and high degree of efficiency compared with other advanced oxidations.

The dissolved oxygen (DO) analysis in crude laboratory effluent showed a value of 1.70 mg L^{-1} , a value that can be explained by the high organic load and oils and greases present in the studied effluent. After treatment with Fenton's reagent, the final dissolved oxygen concentration was 4.70 mg L^{-1} , which is an increase of 63.82% in the value of this parameter.

The intermediate reactions of the Fenton process explains the increase of O_2 concentration in the treated wastewater, hydroperoxyl radical in an aqueous solution is formed in reaction to Phase 02 during the regeneration of the catalyst. This leads to a reduction of ferric ions to ferrous ions and forming oxygen gas and acidifying medium (Equation 13) (Morais and Brito, 2016).



According to Frade (2013), the hydroperoxyl radical still undergoes decomposition and generates a superoxide radical ($O_2^{\bullet-}$), since it is involved in the reduction cycle between Fe^{+2} and Fe^{+3} . It also leads to an increase of concentration of O_2 in treated wastewater (Equations 12, 14 and 15).



In addition, residual hydrogen peroxide releases water and oxygen gas during its decomposition, and thus increases the DO concentration in treated effluent. It is worth noting that the decrease of the concentration of organic matter and the aeration in the jar test can also assist in increasing the dissolved oxygen.

After treatment by slow filtration, there was a reduction in the DO value and the final value was 3.22 mg L^{-1} at the end of the integration of the technologies, which included the following: physical absorption in the non-woven synthetic blanket, adsorption of chemicals in the sand layers and biological processes. For instance, dissolved organic material may be trapped in the filter that is adsorbed and metabolized by microorganisms. All these phenomena may affect the variation of the DO values (Arantes et al., 2014).

In the case of the phenolic compounds, crude laboratory effluent showed a high concentration of 84.83 mg L^{-1} , during combined treatments. It showed a 94.12% reduction with a final value of 4.98 mg L^{-1} , despite the fact that high reduction treated effluent is not suitable for the release standards established by Brazilian legislation of 2011. This has a maximum concentration of 0.5 mg L^{-1} of total phenols as a standard for the discharge of effluent, and requires other technologies to provide guidelines for the phenols.

The oils and greases parameter was determined for the crude laboratory effluent after the three treatment technologies had been combined, namely i): coagulation/flocculation; ii) the Fenton reaction; and iii) slow filtration resulting in a reduction of 99.10%, which is a significant value in percentage terms. The values fell from 134.60 mg L⁻¹ in crude laboratory effluent to 1.20 mg L⁻¹ in post-filtration effluent during 400 minutes of treatment. The final result of 1.20 mg L⁻¹ is a much lower value than the established threshold in Brazilian Resolution 2011 (Conama, 2011) for mineral oils up to 20 mg L⁻¹ and 50 mg L⁻¹ for vegetable oils and animal fats.

In the analysis of total organic carbon, the crude laboratory effluent showed a total organic carbon (TOC) value of 56.42 mg L⁻¹ after combined treatments, a reduction of 87.54% with a final value of 7.02 mg L⁻¹. The results showed that combined processes are efficient, not only in degrading organic compounds, but also perhaps in the degradation that is continued until a complete mineralization.

When the COD value (1400.49 mg O₂ L⁻¹) and TOC value (56.42 mg L⁻¹) were calculated for crude laboratory effluent, the value obtained in TOC / COD ratio was 0.04. After treatment, COD and TOC values reached 346.21 mg O₂ L⁻¹ and 7.02 mg L⁻¹ respectively. The TOC/COD ratio fell to 0.02, which is indicative of the destruction of heteroatomic groups and unsaturations (Aquino et al., 2006; Morais and Brito, 2015; 2016).

In an analysis of the solids series for this type of effluent, there are no settleable solids and the most significant results are a 54.73% increase over total fixed solids treatment. This increase over total fixed solids treatment represents an estimate of inorganic matter, an interesting fact that shows the mineralization of organic compounds in CO₂, H₂O and inorganic ions. In the analysis of total volatile solids that represent an estimate of organic matter, there was an overall reduction of 74.49%, a result that can be regarded as significant (Munoz et al., 2012).

3.5. Reuse of wastewater for non-potable purposes.

The criteria and guidelines used to determine the possibility of reusing treated liquid effluents, were based on publications from government agencies such as ABNT (1997) and USEPA (2012), as well as from Oenning Junior and Pawlowsky (2007); Tonetti et al., (2012); and Da Silva and Santana (2014).

In this article, a study was carried out of two guidelines mentioned previously and a summary was made of the limit values of the main analytical parameters, that included the feasibility of reuse for non-potable purposes.

The parameters studied for this purpose were as follows: pH, total iron, COD, dissolved oxygen, turbidity, Total Organic Carbon (TOC) and oils and greases.

The analytical parameters pH (6.73), total iron (0.56 mg L⁻¹), dissolved oxygen (DO 3.22 mg L⁻¹), turbidity (0.78 NTU) and total organic carbon (7.02 mg L⁻¹) are below the thresholds set by the guidelines with regard to reuse for non-potable purposes.

The values for reuse (non-potable purposes) are:

- pH value 6-9
- Total iron ≤ 5.00 mg L⁻¹
- DO ≥ 2.00 mg L⁻¹
- Turbidity ≤ 30 NTU
- TOC ≤ 10.00 mg L⁻¹

However, the COD analyses (346.21 mg O₂ L⁻¹) and oils and greases (1.20 mg L⁻¹) did not reach the thresholds set by the cited guidelines.

The values for reuse (non-potable purposes) are:

- COD $\leq 90.00 \text{ mg O}_2 \text{ L}^{-1}$
- Oils and greases $\leq 0.05 \text{ mg L}^{-1}$

When adopting an approach using only Brazilian legislation (ABNT, 1997), DO, Oils and greases, settleable solids and pH values were included in Class A for reuse for surface water recharge (the dam designed for public supply, or for rivers up to 10 km upstream of this dam, regardless of the distance from the catchment point and reservoir volume).

4. CONCLUSION

The optimum concentrations of anionic polymer, Fe^{+2} , H_2O_2 and pH values were obtained during a treatment that integrated coagulation/flocculation with the Fenton treatment followed by slow filtration. They had the following values: Coagulation/Flocculation (total iron = 82.49 mg L^{-1} , anionic polymer = 0.80 mg L^{-1} and pH = 6.00), Fenton reagent (total iron = 13.00 mg L^{-1} , H_2O_2 = 200.00 mg L^{-1} and pH = 3.00) and slow filtration (pH = 7.00 and filtration rate $3 \text{ m}^3 \text{ m}^{-2} \cdot \text{d}^{-1}$).

The coagulation/flocculation/Fenton reagent followed by slow filtration showed the main final removals: 75.27% COD removal with a final value of $346.21 \text{ mgO}_2 \text{ L}^{-1}$; 83.50% absorbance removal (200 nm) displaying a final value of 0.94; 86.25% absorbance removal (320nm) with a final value of 0.22 and 99.41% turbidity removal with final turbidity of 0.78 NTU. The total treatment time was 400 minutes.

The levels of total organic carbon (TOC) in treated effluent were evaluated by assessing the efficiency of the organic matter removal. The effluent treated at the end of the integration showed a reduction of 87.54% with a final value of 7.02 mg L^{-1} ; it can thus be concluded that combined treatments are efficient in the removal of organic compounds and can lead to degradation in the form of a complete mineralization.

The possibility of reuse for non-potable purposes of laboratory effluent when the obtained results are compared is established through values based on a set of guidelines. It can be concluded that treated effluent is not feasible for reuse, since it is necessary to increase the efficiency of the treatment by conducting more stages of treatment or improving the studied technologies, as well as monitoring other analytical parameters.

5. ACKNOWLEDGEMENTS

The authors would like to thank CAPES for the scholarship granted to Master's Degree student Chrystopher Allan Miranda Pereira.

6. BIBLIOGRAPHICAL REFERENCES

AGÊNCIA GOIANA DE REGULAÇÃO, CONTROLE E FISCALIZAÇÃO DE SERVIÇOS PÚBLICOS – AGR. Resolução nº 68/2009. Dispõe sobre o Regulamento dos Serviços de Abastecimento de Água e de Esgotamento Sanitário da empresa de Saneamento de Goiás S/A. **Diário Oficial [de] Goiás**, n. 20601, 24 abr. 2009.

AMERICAN PUBLIC HEALTH ASSOCIATION – APHA. **Standard Methods for the Examination of Water and Wastewater**. 20. ed. New York, 1998.

- AQUINO, S. F.; SILVA, S. Q.; CHERNICHARO, C. A. Considerações práticas sobre o teste de demanda química de oxigênio (DQO) aplicado à análise de efluentes anaeróbios. **Revista de Engenharia Sanitária Ambiental**, v. 11, n. 4, p. 295-304, 2006.
- ARANTES, C. C.; RIBEIRO, T. A. P.; PATERNIANI, J. E. S.; SLUSSAREK, M.; SILVA, G. K. E. Uso de coagulantes naturais à base de moringa oleífera e tanino como auxiliadores da filtração em mantas sintética não tecidas. **Engenharia Agrícola**, v. 34, p.780-788, 2014.
- ASHRAFI, M.; CHAMJANGALI, M. A.; BAGHERIAN, G.; GOUDARZI, N. Application of linear and non-linear methods for modeling removal efficiency of textile dyes from aqueous solutions using magnetic Fe₃O₄ impregnated onto walnut shell. **Spectrochimica Acta Part A: Molecular and Biomolecular Spectroscopy**, v. 171, p. 268-279, 2017. <https://doi.org/10.1016/j.saa.2016.07.049>
- ASSOCIAÇÃO BRASILEIRA DA INDÚSTRIA DE HIGIENE PESSOAL, PERFUMARIA E COSMÉTICOS – ABIHPEC. **Panorama do setor de HPPC**. 2016, disponível em: <https://goo.gl/csg5Nt>. Acesso em: 20 mar. 2017.
- ASSOCIAÇÃO BRASILEIRA DE NORMAS TÉCNICAS – ABNT. **Nbr 13969**: tanques sépticos - unidade de tratamento complementar e disposição dos efluentes líquidos – projeto, construção e operação. Rio de Janeiro, 1997.
- BABUPONNUSAMI, A.; MUTHUKUMAR, K. A review on Fenton and improvements to the Fenton process for wastewater treatment. **Journal of Environmental Chemical Engineering**, v. 2, p. 557-572, 2014. <https://doi.org/10.1016/j.jece.2013.10.011>
- BANERJEE, P.; DEY, S. S.; SWARNAKAR, S.; MUKHOPADHYAY, A.; GHOSH, S. Treatment of cosmetic effluent in different configurations of ceramic UF membrane based bioreactor: Toxicity evaluation of the untreated and treated wastewater using catfish (*Heteropneustes fossilis*). **Chemosphere**, v. 146, p. 133-144, 2016. <https://doi.org/10.1016/j.chemosphere.2015.12.004>
- BAGUNDOL, T. B.; AWA, A. L.; ENGUITO, M. R. Efficiency of slow sand filter in purifying well water. **Journal Multidisciplinary Studies**, v. 2, n. 1, 2013.
- BRITO, N. N.; PATERNIANI, J. E. S.; BROTA, G. A.; PELEGRINI, R. T. Ammonia removal from leachate by photochemical process using H₂O₂. **Revista Ambiente & Água**, v. 5, p. 51-60, 2010.
- BRITO, N. N.; LEITE, A. G. C.; CRUZ, T. M. G. S.; LIMA, T. C. R.; BORGES, F. A. Remediação de corante dispersivo utilizando tecnologia de tratamento: Fotoquímica H₂O₂/Vis. **Química Têxtil**, v. 118, p. 36-52, 2015.
- CAVALCANTE, A. M. **Técnicas oxidativas para a remoção de matéria orgânica de uma corrente de soda exausta de refinaria de petróleo**. 2005. 145p. Dissertação (Mestrado) - Universidade Federal do Rio de Janeiro, Rio de Janeiro, 2005.
- CONSELHO NACIONAL DO MEIO AMBIENTE - CONAMA. Resolução nº 357, de 17 de março de 2005. Dispõe sobre a classificação dos corpos de água e diretrizes ambientais para o seu enquadramento, bem como estabelece as condições e padrões de lançamento de efluentes, e dá outras providências. **Diário Oficial [da] União**, n. 53, p. 58-63, 18 mar. 2005.

- CONSELHO NACIONAL DO MEIO AMBIENTE - CONAMA. Resolução nº 430, de 13 de maio de 2011. Dispõe sobre as condições e padrões de lançamento de efluentes, complementa e altera a Resolução nº 357, de 17 de março de 2005, do Conselho Nacional do Meio Ambiente - CONAMA. **Diário Oficial [da] União**, n. 92, p. 89, 16 maio 2011.
- DA SILVA, M. A.; SANTANA, C. G. Reuso de água: Possibilidades de redução do desperdício nas atividades domésticas. **Periódico do centro de estudos em desenvolvimento sustentável da UNDB**, n. 1, p. 1-14, 2014.
- DURIGAN, M. A. B.; VAZ, S. R.; PERALTA-ZAMORA, P. Degradação de poluentes emergentes por processo Fenton e Foto Fenton. **Química Nova**, v. 35, n. 7, p. 1381-1387, 2012.
- FRADE, V. M. F. **Oxidação química de enrofloxacin pelo processo Fenton**. 2013. 193p. Dissertação (Mestrado em Tecnologia Químico-Farmacêutica) - Faculdade de Ciências Farmacêuticas, Universidade de São Paulo, São Paulo, 2013. <http://dx.doi.org/10.11606/D.9.2013.tde-27032014-162133>
- GAMA, M. R. Processos Fenton como Alternativa na Remoção de Interferentes Endócrinos e outros Micropoluentes Ambientais. **Revista Virtual de Química**, v. 4, n. 6, p. 777-787, 2012. <http://dx.doi.org/10.5935/1984-6835.20120056>
- GIANNAKIS, S.; LÓPEZ, S. M. I. P.; SPUHLER, D.; PÉREZ, J. A. S.; IBÁÑEZ, P. F.; PULGARIN, C. Solar disinfection is an augmentable, in situ-generated photo-Fenton reaction—Part 1: A review of the mechanisms and the fundamental aspects of the process. **Applied Catalysis B: Environmental**, v. 199, p. 199-223, 2016. <https://doi.org/10.1016/j.apcatb.2016.06.009>
- GUO, J. S.; ABBAS, A. A.; CHEN, Y. P.; LIU, Z. P.; FANG, F.; CHEN, P. Treatment of landfill leachate using a combined stripping, Fenton, SBR, and coagulation process. **Journal of Hazardous Materials**, v. 178, p. 699-705, 2010. <https://doi.org/10.1016/j.jhazmat.2010.01.144>
- HU, Y.; WANG, X. Application of surrogate parameters in characteristic UV-vis absorption bands for rapid analysis of water contaminants. **Sensors and Actuators B: Chemical**, v. 239, p. 718-726, 2017. <https://doi.org/10.1016/j.snb.2016.08.072>
- KHANDARE, R. V.; GOVINDWAR, S. P. Phytoremediation of textile dyes and effluents: Current scenario and future prospects. **Biotechnology Advances**, v. 33, p. 1697-1714, 2015. <https://doi.org/10.1016/j.biotechadv.2015.09.003>
- LIN, M.; NING, X.; AN, T.; ZHANG, J.; CHEN, C.; KE, Y. et al. Degradation of polycyclic aromatic hydrocarbons (PAHs) in textile dyeing sludge with ultrasound and Fenton processes: Effect of system parameters and synergistic effect study. **Journal of Hazardous Materials**, v. 307, p. 7-16, 2016. <https://doi.org/10.1016/j.jhazmat.2015.12.047>
- MA, X. J.; XIA, H. L. Treatment of water-based printing ink wastewater by Fenton process combined with coagulation. **Journal of Hazardous Materials**, v. 162, p. 386-390, 2009. <https://doi.org/10.1016/j.jhazmat.2008.05.068>
- MARCIONILIO, S. M. L. O.; BARROS, V. S.; LÉON, J. J. L. Avaliação da coagulação química e eletrocoagulação no tratamento de efluente de polo farmacêutico. **RQI- Revista Química Industrial**, n. 749, p. 22-31, 2015.

- MIRZAEI, A.; CHEN, Z.; HAGHIGHT, F.; YERUSHALMI, L. Removal of pharmaceuticals from water by homo/heterogeneous Fenton-type processes – A review. **Chemosphere**, v. 174, p. 665-688, 2017. <https://doi.org/10.1016/j.chemosphere.2017.02.019>
- MORAIS, W. O.; BRITO, N. N. Tratamento de efluente de indústria cosmética via reagente de Fenton. **Revista de Química Industrial**, v. 84, n. 753, p.52-68, 2016.
- MORAIS, W. O.; BRITO, N. N. Treatment of simulated wastewater via Fenton reaction at a production line of sunscreen. **Oxidation Communications**, v. 38, n. 4A, p. 2261-2272, 2015.
- MUNOZ, M.; PEDRO, Z. M.; PLIEGO, G.; CASAS, J. A.; RODRIGUEZ, J. J. Chlorinated byproducts from the Fenton-like oxidation of polychlorinated phenols. **Industrial e Engineering Chemistry Research**, v. 51, p. 13092-13099, 2012. <http://dx.doi.org/10.1021/ie3013105>
- NASCIMENTO, A. P.; PELEGRINI, R. T.; BRITO, N. N. Filtração Lenta para o tratamento de águas para pequenas comunidades rurais. **Revista Eletrônica de Engenharia Civil**, v. 2, p. 54-58, 2012. <https://doi.org/10.5216/reec.v4i2.19705>
- NOGUEIRA, R. F. P.; TROVÓ, A. G.; SILVA, M. R. A.; VILLA, R. D. Fundamentos e aplicações ambientais dos processos Fenton e foto-Fenton. **Química Nova**, v. 30, n. 2, p. 400-08, 2007. <http://dx.doi.org/10.1590/S0100-40422007000200030>
- OENNING JUNIOR, A.; PAWLOWSKY, U. Avaliação de tecnologias avançadas para o reuso de águas em indústria metal-mecânica. **Engenharia Sanitária e Ambiental**, v. 12, n. 3, p. 305-316, 2007. <https://doi.org/10.1590/S1413-41522007000300010>
- OLIVEIRA, M. C.; NOGUEIRA, R. F. P.; GOMES NETO, J. A.; JARDIM, W. F.; ROHWEDDER, J. J. R. Sistema de injeção em fluxo espectrofotométrico para monitorar peróxido de hidrogênio em processo de fotodegradação por reação foto-fenton. **Química Nova**, v. 24, n. 2, p. 188-190, 2001. <http://dx.doi.org/10.1590/S0100-40422001000200007>
- POMPEI, C. M. E.; CIRIC, L.; CANALES, M.; KARU, K.; VIEIRA, E. M.; CAMPOS, L. C. Influence of PPCPs on the performance of intermittently operated slow sand filters for household water purification. **Science of The Total Environment**, v. 581–582, p. 174-185, 2017. <https://doi.org/10.1016/j.scitotenv.2016.12.091>
- QUARTAROLI, L.; KURITZA, J. C.; CAVALLINI, G. S.; VIDAL, C. M. S.; SOUZA, J. B. Aplicação de polímeros aniônico e catiônico na flotação por ar dissolvido como forma de pós-tratamento de efluente de indústria de papel e celulose. **Scientia Forestalis**, v. 42, n. 101, p. 57-67, 2014.
- RAGASSON, M. K. Aplicação do reagente Fenton no tratamento de lixiviado de aterros sanitários. **Especialize**, n. 4, 2013.
- SALVADOR, T.; MARCOLINO, L. H.; PERALTA-ZAMORA, P. Degradação de corantes têxteis e remediação de resíduos de tingimento por processos Fenton, Foto Fenton e Eletro Fenton. **Química Nova**, v. 35, n. 5, p. 932-938, 2012.
- SHAIKH, S. M. R.; NASSER, M. S.; HUSSEIN, I. A.; BENAMOR, A. Investigation of the effect of polyelectrolyte structure and type on the electrokinetics and flocculation behavior of bentonite dispersions. **Chemical Engineering Journal**, v. 311, p. 265-276, 2017. <https://doi.org/10.1016/j.cej.2016.11.098>

- SEMERARO, P.; RIZZI, V.; FINI, P.; MATERA, S.; COSMA, P.; FRANCO, E. et al. Interaction between industrial textile dyes and cyclodextrins. **Dyes and Pigments**, v. 119, p. 84–94, 2015. <https://doi.org/10.1016/j.dyepig.2015.03.012>
- SOHRABI, M. R.; KHAVARAN, A.; SHARIATI, S.; SHARIATI, S. Removal of Carmoisine edible dye by Fenton and photo Fenton processes using Taguchi orthogonal array design. **Arabian Journal of Chemistry**, v. 10, p. S3523-S3531, 2017. <https://doi.org/10.1016/j.arabjc.2014.02.019>
- TONETTI, A. L.; CORAUCCI FILHO, B.; NICOLAU, C. E.; BARBOSA, M.; TONON, D. Tratamento de esgoto e produção de água de reuso com o emprego de filtro lento. **Engenharia Sanitária e Ambiental**, v. 17, n. 3, p. 287-294, 2012. <http://dx.doi.org/10.1590/S1413-41522012000300005>
- TORRADESA, F.; MONTAÑOB, J. G. Using central composite experimental design to optimize the degradation of real dye wastewater by Fenton and photo-Fenton reactions. **Dyes and Pigments**, v. 100, p. 184-189, 2014.
- UNITED STATES. Environmental Protection Agency - USEPA. **Guidelines for Water Reuse**. EPA/600/R-12/618. Washington, DC, 2012. 450 p.
- WIŚNIEWSKA, M.; CHIBOWSKI, S.; URBAN, T. Adsorption properties of the nanozirconia/anionic polyacrylamide system—Effects of surfactant presence, solution pH and polymer carboxyl groups content. **Applied Surface Science**, v. 370, p. 351-356, 2016. <https://doi.org/10.1016/j.apsusc.2016.02.188>