



Journal of the Mexican Chemical Society

ISSN: 1870-249X

editor.jmcs@gmail.com

Sociedad Química de México

México

Shah, Jasmin; Rasul Jan, M.; Inayatullah; Shah, Sultan  
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pharmaceuticals  
Journal of the Mexican Chemical Society, vol. 56, núm. 2, abril-junio, 2012, pp. 109-114  
Sociedad Química de México  
Distrito Federal, México

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## Sensitive spectrofluorimetric and spectrophotometric methods for determination of sparfloxacin in pharmaceuticals

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Received August 4, 2011; Accepted January 12, 2012

**Abstract.** Two simple, sensitive and accurate spectrofluorimetric and spectrophotometric methods have been developed for the determination of sparfloxacin in bulk and pharmaceutical preparations. The proposed methods were based on oxidation of sparfloxacin with Ce (IV) in acidic medium. The spectrophotometric method involved the measurement of unconsumed Ce (IV) concentration at 315 nm. The spectrofluorimetric method based on the measurement of reduced fluorescent Ce (III) at 352 nm after excitation at 250 nm. Different variables affecting the reaction such as concentration and volume of cerium (IV), type and concentration of acidic medium, heating temperature and time were carefully studied and optimized. Under the optimum conditions, linear relationship in the range of 0.02-0.2  $\mu\text{g mL}^{-1}$  and 0.02-0.1  $\mu\text{g mL}^{-1}$  were obtained using spectrophotometric and spectrofluorimetric methods, respectively. No interferences were observed from the common formulations excipients present in the dosage form of the drug. The proposed methods were successfully applied to the analysis of the investigated drug in pure and pharmaceutical formulations with good accuracy and precision. The recovery percentage ranged from 93-102  $\pm$  1.73-2.66%. The precision of the methods were good; RSD  $\leq$  2.55%.

**Key words:** Sparfloxacin, Cerium (IV), spectrophotometry, spectrofluorimetry, redox reaction.

**Resumen.** Se desarrollaron dos métodos espectrofluorimétricos simples, sensibles y exactos para la determinación de esparfloxacin en preparaciones a granel y farmacéuticas. Los métodos propuestos se basaron en la oxidación de esparfloxacin con Ce (IV) en medio ácido. El método espectrofotométrico involucra la concentración de Ce (IV) no consumido a 315 nm. El método espectrofluorimétrico se basa en la medida de Ce (III) fluorescente reducido a 352 nm después de excitación a 250 nm. Diferentes variables afectan la reacción como son la concentración y el volumen de cerio (IV), tipo y concentración del medio ácido, temperatura y tiempo de calentamiento fueron cuidadosamente estudiadas y optimizadas. Bajo condiciones óptimas, la relación lineal se encontró dentro de los rangos de 0.02-0.2  $\mu\text{g mL}^{-1}$  y 0.02-0.1  $\mu\text{g mL}^{-1}$  que fueron obtenidos utilizando los métodos espectrofotométricos y espectrofluorimétricos, respectivamente. No fueron observadas interferencias de los excipientes en las formulaciones comunes en la dosificación de la forma del fármaco. Los métodos propuestos fueron satisfactoriamente aplicados al análisis del fármaco investigado en formulaciones puras y farmacéuticas con buena exactitud y precisión. El porcentaje de recobro se encuentra en el rango de 93-102  $\pm$  1.73-2.66%. La precisión de los métodos fue buena; DRE  $\leq$  2.55%.

**Palabras clave:** Esparfloxacin, Cerio (IV), espectrofotometría, espectrofluorimetría, reacción redox.

### Introduction

Sparfloxacin [5-amino-1-cyclopropyl-7-(cis-3, 5-dimethyl-1-piperazinyl)-6, 8-difluoro-1, 4-dihydro-4-oxo-3-quinoline carboxylic acid] is a third generation fluoroquinolones which is the one of the most important and successful classes of manmade antibacterial used in the treatment of lung infection, urinary tract infection and cutaneous allergy. Sparfloxacin (spflx) is receiving attention due to its broad spectrum activity, potency and excellent pharmacokinetic profiles [1-3].

Different analytical methods have been reported for quantitative determination of sparfloxacin in pharmaceutical preparation like DC polarography [1], HPLC [4-6], and capillary zone electrophoresis [3, 7, 8]. Spectrophotometric methods have also been reported based on the indirect determination of sparfloxacin with N- bromosuccinimide [9], ion-pair complex with bromothymol blue [2], ternary complex with Pd(II) and eosin [10] and UV-Spectrophotometry [11, 12]. Few spectrofluorimetric methods have also been reported for the analysis of sparfloxacin. These methods are based on intrinsic fluorescence property or complexation reaction with aluminum chloride [13], derivatization of sparfloxacin with hydrobromic acid [14], hydrochloric acid [15] and potassium iodide [16].

Ce (IV) is a well known oxidizing agent. It is used for the indirect determination of a number of drugs [17-24].

The present work was carried out to develop simple and sensitive spectrophotometric and spectrofluorimetric methods for the determination of sparfloxacin in bulk and pharmaceutical formulations based on its reduction of Ce (IV) to Ce (III). The decrease in absorbance of Ce (IV) was measured at 315 nm in spectrophotometric method and the fluorescence intensity of produced Ce (III) was measured at 352 nm with excitation at 250 nm in spectrofluorimetric method.

### Experimental

#### Instruments and Chemicals

Fluorescence intensity was measured using RF 5301 Spectrofluorophotometer (Shimadzu, Japan) with Xenon lamp (UXL-155), Optima SP-3000 plus UV/Vis Spectrophotometer (Tokyo Japan), with quartz cells for absorbance measurements and thermostatically controlled water bath (Yu Jia China) was used for heating purpose.

All chemicals used were of high grade purity. Ammonium cerium (IV) sulfate dihydrate (reagent grade ACS, Scharlau Barcelona, Spain), sulphuric acid (95-98% Merck, Darmstadt Germany) and methanol (Merck, Darmstadt Germany) were used during this work. Standard reference sparfloxacin was

gifted by Libra Pharmaceutical Industry pvt. Ltd., Peshawar, Pakistan. Commercial formulations of sparfloxacin (Sparxin, Sparcin and Quspar) were purchased from local market. Sparfloxacin stock standard solution was prepared by dissolving 10 mg of authentic standard sparfloxacin in 10 mL of distilled methanol with vigorous shaking and diluted with distilled water to 100 mL volumetric flask.  $10\ \mu\text{g mL}^{-1}$  working standard solution was prepared by dilution with distilled water from the stock solution. Fresh solutions of Ce (IV) were prepared daily by dissolving appropriate amount of ammonium cerium (IV) sulfate dihydrate  $[(\text{NH}_4)_4\text{Ce}(\text{SO}_4)_4 \cdot 2\text{H}_2\text{O}]$  in 1.5 M and 1.0 M sulphuric acid, so as to obtain  $500\ \mu\text{g mL}^{-1}$  (Solution A) and  $100\ \mu\text{g mL}^{-1}$  (Solution B) solutions respectively. Stock solutions of sulphuric acid (1.0 M and 1.5 M) were prepared by adding 13.6 mL and 20.4 mL of 95-98% pure sulphuric acid to sufficient cold distilled water and diluted to 250 mL with cold distilled water.

### Sample preparation

Five tablets were weighed, grounded and mixed. After fine powdering of the tablets, an accurate weighed portion of the sample equivalent to contain 10 mg sparfloxacin was dissolved in 10 mL distilled methanol and vigorously shaken. The solution was filtered and diluted with distilled water in 100 mL flask.

### General Procedure

#### *Spectrofluorometric method*

Sparfloxacin solutions in the range of  $0.02\text{--}0.10\ \mu\text{g mL}^{-1}$  from  $10\ \mu\text{g mL}^{-1}$  working standard solution were transferred to 100 mL beakers. To each solution, 4 mL of Ce (IV) solution B was added. The mixture was heated on boiling water bath for 30 min. On cooling contents of the beakers were transferred quantitatively to 25 mL volumetric flask and diluted up to mark with distilled water. Fluorescence intensity of the produced Ce (III) was measured at 352 nm after excitation at 250 nm against a reagent blank prepared by the same method without addition of the drug.

#### *Spectrophotometric method*

Sparfloxacin solutions in the range of  $0.02\text{--}0.20\ \mu\text{g mL}^{-1}$  from  $10\ \mu\text{g mL}^{-1}$  working standard solution were transferred to 100 mL beakers. To each solution, 5 mL of Ce (IV) solution A was added. The mixture was heated on boiling water bath for 30 min. On cooling contents of the beakers were transferred to 25 mL volumetric flask and diluted up to mark with distilled water. The blank was prepared by the same method without addition of the drug and absorbance was measured at 315 nm by UV/Vis Spectrophotometer (Model SP-3000 plus, Optima Tokyo Japan) and the analyte absorbances were subtracted from the blank absorbance.

## Result and discussion

Cerium (IV) is a good oxidizing agent. Its solution in acidic medium is yellow in colour with  $\lambda_{\text{max}}$  at 315 nm while its reduced form Ce (III) is colorless and possesses an intrinsic fluorescence property having emission at 352 nm after excitation at 250 nm. Sparfloxacin reacts with Ce (IV) in acidic medium, reduces Ce (IV) to Ce (III). Quantification of the drug can be carried out by monitoring either the decrease in absorbance of Ce (IV) with the help of spectrophotometer or fluorescence intensity of the produced Ce (III) with the help of spectrofluorometer. The absorbance measurement at 315 nm for reagent solution [Ce (IV)] against a reaction solution [Ce (IV) with sparfloxacin] will give the amount of reacted Ce (IV) with sparfloxacin. While the fluorescence measurement of the reaction solution at 352 nm after excitation at 250 nm gives the amount of Ce (III) produced equal to the amount of sparfloxacin.

Factors affecting the redox reaction of sparfloxacin with Ce (IV) were investigated by changing one variable while keeping the others constant through a series of experiments.

### Effect of temperature and heating time

The reaction between sparfloxacin and Ce (IV) was relatively slow at room temperature. Therefore, the effect of temperature and heating time on the redox reaction were studied in the range of  $50\text{--}100\ ^\circ\text{C}$  and  $10\text{--}60$  min, respectively. Maximum absorbance of the reagent blank against the analyte and maximum fluorescence intensity of the produced Ce (III) was found at  $100\ ^\circ\text{C}$  with 30 min heating time.

### Effect of acidity

The effect of acidity on the reaction was studied in the range of  $0.0\text{--}2.5$  M sulphuric acid solution. Maximum analytical signal was obtained with 1.0 M  $\text{H}_2\text{SO}_4$ . Above and below the optimum concentration of acid, decrease in signal was observed. To obtain stable and high fluorescence intensity 1.0 M  $\text{H}_2\text{SO}_4$  solution was used throughout the work for preparation of Ce (IV) solution.

### Effect of concentration of Ce (IV) solution

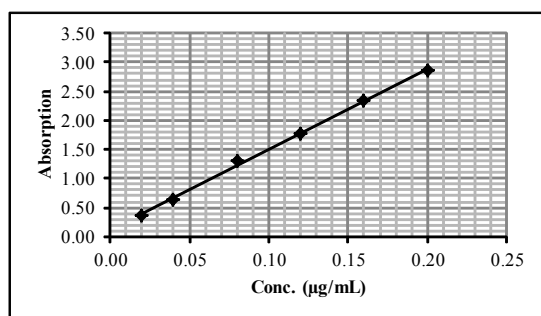
Concentration of Ce (IV) solution affects the reaction very markedly. The influence of Ce (IV) was studied in the range of  $50\text{--}500\ \mu\text{g mL}^{-1}$  for both methods. It was found that with increase in concentration of Ce (IV) from  $50\ \mu\text{g mL}^{-1}$  to  $500\ \mu\text{g mL}^{-1}$ , absorbance of the reagent blank against analyte continuously increases. Due to the maximum signal on the spectrophotometer,  $500\ \mu\text{g mL}^{-1}$  Ce (IV) solution was selected for spectrophotometric method and  $100\ \mu\text{g mL}^{-1}$  Ce (IV) solution was selected for spectrofluorometric method due to stable and reproducible results. The effect of volume of Ce (IV) solution on the absorbance behavior of Ce (IV) and fluorescence intensity of the produced Ce (III) was also investigated in the

range 1-5 mL for spectrophotometric method and 1-8 mL for spectrofluorometric method. 5 mL of Ce (IV) solution A for the spectrophotometric method was found optimum volume and 4 mL of Ce (IV) solution B for the spectrofluorometric method was found optimum volume. Therefore, 5 and 4 mL of Ce (IV) solutions were used throughout the analysis.

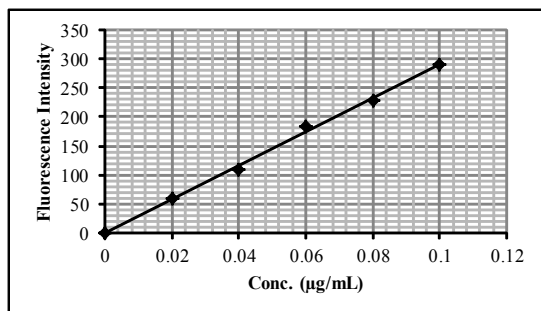
### Validation of method

Under the optimum experimental conditions of the proposed methods, a linear relationship between the absorbance and concentration of sparfloxacin was observed in the concentration range of 0.02-0.2  $\mu\text{g mL}^{-1}$ . A linear plot of fluorescence intensity with concentration of sparfloxacin was also obtained in the range of 0.02-0.1  $\mu\text{g mL}^{-1}$  for spectrofluorometric method. A calibration curve for spectrophotometric and spectrofluorometric method for determination of sparfloxacin is given in figure 1 and 2. The linear regression equations, slopes, intercepts, correlation coefficients, standard deviation and relative standard deviation of the response factors are given in table 1 and 2. The molar absorptivity was found to be  $3.24 \times 10^6 \text{ L mol}^{-1} \text{ cm}^{-1}$ . The LOD (3 s/b) and LOQ (10 s/b) values were found to be  $1.6 \times 10^{-3} \mu\text{g mL}^{-1}$  and  $5.2 \times 10^{-3} \mu\text{g mL}^{-1}$  for spectrophotometric method and  $6.0 \times 10^{-3} \mu\text{g mL}^{-1}$  and  $20.0 \times 10^{-3} \mu\text{g mL}^{-1}$  for spectrofluorometric method respectively.

The proposed method was validated according to ICH (International Conference on Harmonization) guidelines on the validation of analytical methods [25].



**Fig. 1.** Effect of concentration of the drug on the absorption behavior of sparfloxacin.



**Fig. 2.** Effect of concentration of the drug on the fluorometric behavior of Ce (III).

**Table 1.** Analytical parameters for the spectrofluorimetric determination of sparfloxacin.

Parameter	Value
$\lambda_{\text{ex}}$ (nm)	250
$\lambda_{\text{em}}$ (nm)	352
Concentration range ( $\mu\text{g mL}^{-1}$ )	0.02-0.1
Limit of detection ( $\mu\text{g mL}^{-1}$ )	$6.0 \times 10^{-3}$
Limit of quantification ( $\mu\text{g mL}^{-1}$ )	$2.0 \times 10^{-2}$
Regression equation (y)	$Y = 2907.3X + 0.327$
Slope (b)	$2.91 \times 10^3$
Intercept (a)	0.327
Correlation coefficient (r)	0.9975
Relative standard deviation (%)	1.12

**Table 2.** Analytical parameters for the spectrophotometric determination of sparfloxacin.

Parameter	Value
$\lambda_{\text{max}}$ (nm)	315
Concentration range ( $\mu\text{g mL}^{-1}$ )	0.02-0.2
Molar absorptivity ( $\text{L mol}^{-1} \text{ cm}^{-1}$ )	$3.24 \times 10^6$
Limit of detection ( $\mu\text{g mL}^{-1}$ )	$1.6 \times 10^{-3}$
Limit of quantification ( $\mu\text{g mL}^{-1}$ )	$5.2 \times 10^{-3}$
Regression equation (y)	$Y = 8.98X + 0.018$
Slope (b)	8.98
Intercept (a)	0.018
Correlation coefficient (r)	0.999
Relative standard deviation (%)	2.55

### Specificity of the method

To evaluate the specificity (selectivity) of the developed methods for the analysis of pharmaceutical preparations containing sparfloxacin, the interferences effect of various excipients used in pharmaceutical preparations such as lactose, glucose, sucrose, sorbitol, magnesium stearate, talc and starch were studied. Solutions containing sparfloxacin and one of the excipients taken separately in concentrations four, eight and twelve times greater than that of the sparfloxacin were analyzed by the proposed spectrophotometric method and two, four and eight times greater than that of the sparfloxacin were analyzed by the proposed spectrofluorometric method. A level of interference was considered to be acceptable if the error was not higher than 3% relative to the expected sparfloxacin value. No interferences were observed in the determination of sparfloxacin in the presence of the common excipients studied (Table 3 and 4). The recovery of sparfloxacin was 95-105% and 94.5-103% for spectrofluorometric and spectrophotometric methods respectively.

The degradation products of sparfloxacin were also studied to check the specificity of the present method. The method was found free of interferences from the degraded products of sparfloxacin. The degraded products of sparfloxacin are

**Table 3.** Spectrofluorimetric determination of sparfloxacin (0.04  $\mu\text{g mL}^{-1}$ ) in the presence of excipients ( $n = 3$ ).

Excipient	Excipient added ( $\mu\text{g mL}^{-1}$ )	% Recovery $\pm$ RSD
Glucose	0.08	96.6 $\pm$ 2.28
	0.16	95.0 $\pm$ 2.72
	0.32	102.0 $\pm$ 0.78
Sucrose	0.08	100.0 $\pm$ 2.02
	0.16	100.0 $\pm$ 2.29
	0.32	97.0 $\pm$ 1.29
Sorbitol	0.08	100.0 $\pm$ 0.25
	0.16	100.0 $\pm$ 1.54
	0.32	102.0 $\pm$ 1.94
Lactose	0.08	100.0 $\pm$ 1.70
	0.16	96.0 $\pm$ 0.93
	0.32	99.0 $\pm$ 0.60
Starch	0.08	101.0 $\pm$ 0.27
	0.16	96.0 $\pm$ 2.90
	0.32	100.0 $\pm$ 1.75
Talc	0.08	101.0 $\pm$ 1.50
	0.16	101.0 $\pm$ 1.75
	0.32	97.0 $\pm$ 1.00
Mg. Stearate	0.08	103.0 $\pm$ 0.35
	0.16	103.0 $\pm$ 0.30
	0.32	101.0 $\pm$ 0.30

**Table 4.** Spectrophotometric determination of sparfloxacin (0.04  $\mu\text{g mL}^{-1}$ ) in the presence of excipients ( $n = 3$ ).

Excipient	Excipient added ( $\mu\text{g mL}^{-1}$ )	% Recovery $\pm$ RSD
Glucose	0.16	100.0 $\pm$ 2.28
	0.32	100.0 $\pm$ 1.8
	0.48	95.0 $\pm$ 1.8
Sucrose	0.16	100.0 $\pm$ 1.8
	0.32	100.0 $\pm$ 2.5
	0.48	105.0 $\pm$ 1.8
Sorbitol	0.16	100.0 $\pm$ 1.8
	0.32	100.0 $\pm$ 1.8
	0.48	105.0 $\pm$ 1.8
Lactose	0.16	100.0 $\pm$ 1.8
	0.32	105.0 $\pm$ 1.8
	0.48	100.0 $\pm$ 0.6
Starch	0.16	100.0 $\pm$ 1.8
	0.32	105.0 $\pm$ 1.8
	0.48	95.0 $\pm$ 1.8
Talc	0.16	100.0 $\pm$ 1.8
	0.32	100.0 $\pm$ 1.8
	0.48	100.0 $\pm$ 1.8
Mg. Stearate	0.16	105.0 $\pm$ 1.8
	0.32	95.0 $\pm$ 1.8
	0.48	100.0 $\pm$ 1.8

mainly oxidized products with changes at the piperazine ring (figure 3).

### Repeatability of the method

Repeatability of the proposed methods was checked by determining three different concentrations of sparfloxacin in replicate for each concentration (0.02-0.06  $\mu\text{g mL}^{-1}$  for spectrofluorometric and 0.04-1.3  $\mu\text{g mL}^{-1}$  for spectrophotometric method). The results are summarized in table 5 and 6. The relative standard deviation (RSD) considered being very satisfactory (0.86-3.18%) and the recovery test was found to be in the range of 91.0-98.83% for spectrofluorometric method and recovery from 97.5-100.83% with RSD from 0.89-1.82 for spectrophotometric method, indicates a good accuracy. The proposed methods were found useful for routine analytical

and quality control analysis of sparfloxacin in pharmaceutical preparations as well as in pure form.

### Accuracy of the method

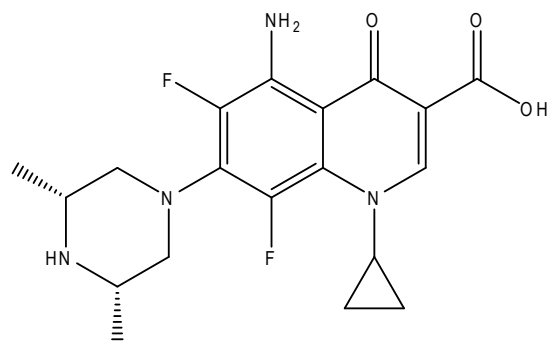
To check accuracy of the proposed method, the standard addition technique was applied using three different brands formulations. The percent recoveries were in the range of 92.75  $\pm$  0.92 % to 102.5  $\pm$  1.73% (tables 7 and 8). This indicates high accuracy of the developed method for sparfloxacin determination in pharmaceutical preparations.

### Reaction stoichiometry

The stoichiometry between sparfloxacin and Ce (IV) was studied by molar ratio method. The results revealed that each mole of sparfloxacin consumes 50 mole of Ce (IV), giving a mixture of oxidized products. The oxidation products were also checked by TLC and many spots were observed on TLC. It was attributed due to the presence of many centers in the sparfloxacin molecule liable to oxidation.

### Analysis of pharmaceutical preparations

The proposed method has been successfully applied to the determination of sparfloxacin in commercial formulation. The results obtained for pharmaceutical dosage forms were compared statistically with respect to accuracy by student t-test at 95% confidence level and precision by the variance ratio F-test with those of the literature HPLC method. There was also no signifi-

**Fig. 3.** Structural formula of sparfloxacin.

**Table 5.** Evaluation of accuracy and precision of the proposed spectrofluorimetric method for sparfloxacin determination (n = 3).

Sample	Amount taken ( $\mu\text{g mL}^{-1}$ )	Amount found ( $\mu\text{g mL}^{-1}$ )	%Recovery $\pm$ RSD	Confidence limit
Sparcin tablets	0.020	0.0193	96.5 $\pm$ 3.0	0.0193 $\pm$ 2.98 $\times 10^{-3}$
	0.040	0.0380	95.0 $\pm$ 1.86	0.0380 $\pm$ 2.98 $\times 10^{-3}$
	0.060	0.0593	98.8 $\pm$ 1.52	0.0593 $\pm$ 2.98 $\times 10^{-3}$
Quspar tablets	0.020	0.0193	95.0 $\pm$ 3.12	0.0186 $\pm$ 1.44 $\times 10^{-3}$
	0.040	0.0383	95.7 $\pm$ 3.13	0.0383 $\pm$ 2.98 $\times 10^{-3}$
	0.060	0.0580	96.7 $\pm$ 1.72	0.0580 $\pm$ 2.48 $\times 10^{-3}$
Sparxin tablets	0.020	0.0192	91.0 $\pm$ 3.18	0.0182 $\pm$ 1.89 $\times 10^{-3}$
	0.040	0.0380	95.0 $\pm$ 1.32	0.0380 $\pm$ 1.24 $\times 10^{-3}$
	0.060	0.058	96.7 $\pm$ 0.86	0.0580 $\pm$ 1.24 $\times 10^{-3}$

**Table 6.** Evaluation of accuracy and precision of the proposed spectrophotometric method for sparfloxacin determination (n = 3).

Sample	Amount taken ( $\mu\text{g mL}^{-1}$ )	Amount found ( $\mu\text{g mL}^{-1}$ )	%Recovery $\pm$ RSD	Confidence limit
Sparcin tablets	0.040	0.039	97.5 $\pm$ 1.8	0.039 $\pm$ 1.76 $\times 10^{-3}$
	0.080	0.080	100.0 $\pm$ 0.9	0.080 $\pm$ 1.76 $\times 10^{-3}$
	0.120	0.119	99.2 $\pm$ 0.8	0.119 $\pm$ 2.48 $\times 10^{-3}$
Quspar tablets	0.040	0.040	100.0 $\pm$ 1.8	0.040 $\pm$ 1.76 $\times 10^{-3}$
	0.080	0.080	100.0 $\pm$ 0.9	0.080 $\pm$ 1.76 $\times 10^{-3}$
	0.120	0.119	99.2 $\pm$ 0.6	0.119 $\pm$ 1.76 $\times 10^{-3}$
Sparxin tablets	0.040	0.040	100.0 $\pm$ 1.8	0.040 $\pm$ 1.76 $\times 10^{-3}$
	0.080	0.080	100.0 $\pm$ 0.9	0.080 $\pm$ 1.76 $\times 10^{-3}$
	0.120	0.121	100.8 $\pm$ 0.6	0.121 $\pm$ 1.76 $\times 10^{-3}$

**Table 7.** Recovery test of sparfloxacin in tablets by the proposed spectrofluorimetric method.

Pharmaceutical preparation	Amount added ( $\mu\text{g mL}^{-1}$ )	Amount found ( $\mu\text{g mL}^{-1}$ )	RE %	% Recovery $\pm$ RSD
Sparcin tablets (100 mg/tablet)	0.02	0.0200	0.0	100.0 $\pm$ 2.54
	0.04	0.0374	6.5	93.5 $\pm$ 1.76
	0.06	0.0558	7.0	93.0 $\pm$ 2.66
Quspar tablets (100mg/tablet)	0.02	0.0202	1.0	101.0 $\pm$ 1.61
	0.04	0.0371	7.25	92.8 $\pm$ 0.92
	0.06	0.0588	2.0	98.0 $\pm$ 0.26
Sparxin tablets (100 mg/ tablet)	0.02	0.0202	1.0	101.0 $\pm$ 1.04
	0.04	0.0409	2.25	102.3 $\pm$ 2.33
	0.06	0.0610	1.67	101.7 $\pm$ 0.99

(Each result is average of the separate triplicate analysis).

**Table 8.** Recovery test of sparfloxacin in tablets by the proposed spectrophotometric method.

Pharmaceutical preparation	Amount added ( $\mu\text{g mL}^{-1}$ )	Amount found ( $\mu\text{g mL}^{-1}$ )	RE %	% Recovery $\pm$ RSD
Sparcin tablets (100 mg/tablet)	0.02	0.019	5.0	95.0 $\pm$ 3.74
	0.04	0.041	2.5	102.5 $\pm$ 1.73
	0.08	0.08	0.0	100.0 $\pm$ 0.88
Quspar tablets (100mg/tablet)	0.02	0.02	0.0	100.0 $\pm$ 3.55
	0.04	0.04	0.0	100.0 $\pm$ 1.78
	0.08	0.079	1.25	98.8 $\pm$ 0.59
Sparxin tablets (100 mg/ tablet)	0.02	0.02	0.0	100.0 $\pm$ 3.55
	0.04	0.039	2.5	97.5 $\pm$ 1.82
	0.08	0.079	1.25	98.8 $\pm$ 1.26

(Each result is average of the separate triplicate analysis).

**Table 9** Determination of sparfloxacin in commercial formulation and validation by reference method.

Name of commercial formulation		Mass (mg) determined	
		Proposed method	Literature method
Tab. Sparcin (100 mg/tab)	mean $\pm$ % RSD	99.91 $\pm$ 0.22	101.85 $\pm$ 1.78
	F (9.28)	0.015	
	t (2.776)	1.73	
Tab. Sparaxin (100 mg/tab)	mean $\pm$ % RSD	100 $\pm$ 0.33	99.81 $\pm$ 0.32
	F (9.28)	0.083	
	t (2.776)	0.91	
Tab. Quspar (100 mg/tab)	mean $\pm$ % RSD	99.5 $\pm$ 0.17	99.97 $\pm$ 1.04
	F (9.28)	0.028	
	t (2.776)	0.81	

cant difference in precision between the proposed and literature methods. The results show similar precision and accuracy in the analysis of pharmaceutical dosage forms (table 9).

## Conclusion

The present study described the use of Ce (IV) reagent for the development of simple, sensitive, accurate and precise spectrophotometric and spectrofluorometric methods for the determination of sparfloxacin in bulk and pharmaceutical formulations. As compared to HPLC methods, the reagents used in the proposed spectrophotometric and spectrofluorimetric methods are readily available and the procedure does not involve any tedious sample preparation. The developed methods are simple and readily adaptable in the quality control laboratories.

## References

- Jain, S.; Jain, N.K.; Pitre, K.S. *J. Pharm. Biomed Anal.* **2002**, 29, 795-801.
- Marona, H.R.N.; Schapoval, E.E.S. *J. Pharm. Biomed Anal.* **2001**, 26, 501-504.
- Faria, A.F.; de Souza, M.V.N.; de Almeida, M.V.; de Oliveira, M.A.L. *Anal. Chim. Acta* **2006**, 579, 185-192.
- Marona, H.R.N.; Schapoval, E.E.S. *J. Pharm. Biomed Anal.* **1999**, 20, 413-417.
- Kowalczyk, D.; Hopkata, H.; Pietras, R. *Chemia Analityczna* **2004**, 49, 201-211.
- CaO, S.X.; Zhang, J.Y.; Liu, H.M. *Chinese J. Chromatogr.* **2001**, 19, 454-456.
- Sun, S.W.; Wu, A.C. *J. Liq. Chromatogr. Relat. Technol.* **1999**, 22, 281-296.
- Fierrens, C.; Hillaert, S.; Bossche, W.V. *J. Pharm. Biomed. Anal.* **2000**, 22, 763-772.
- Askal, H.; Refaat, I.; Darwish, I.; Marzouq, M. *Chem. Pharm. Bull.* **2007**, 55, 1551-1556.
- El-Didamony, A.M. *Anal. Lett.* **2007**, 40, 2708-2720.
- Chowdary, K.P.R.; Rao, G.D. *Indian drugs*, **1997**, 34, 107-110.
- Fratini, L.; Schapoval, E.E.S. *Int. J. Pharm.* **1996**, 127, 279-282.
- Rizk, M.; Belal, F.; Ibrahim, F.; Ahmed, S.; El-Enany, N. *Pharmazie. Acta Helveticae* **2000**, 74, 371-377.
- Du, L.M.; Zhou, J.; Xu, Q.Q.; Fan, Z.F.; Yuan, R. *Chin. Chem. Lett.* **2000**, 11, 255-258.
- Du, L.M.; Xu, Q.Q.; Wang, J.P.; Zuo, H.Q. *Guang Pu XueYu Guang Pu Fen Xi* **2003**, 23, 131-133.
- Liming, D.; Weijun, J.; Chuan, D.; Changsong, L.; Qingqin, X. *Chinese J. Internet* **2000**, 2, 16-19.
- Sultan, S.M.; Alzamil, I.Z.; Alrahman, A.M.A.; Altamrah, S.A.; Asha, Y. *Analyst* **1986**, 111, 919-921.
- Koukli, I.I.; Calokerinos, A.C.; Hadjiioannou, T.P. *Analyst* **1989**, 114, 711-716.
- Calatyud, J.M. *Anal. Chim. Acta* **1992**, 264, 283-289.
- Perez-Rinz, T.; Martinez-Lozano, C.; Tomas, V.; Sidrach de Cardona, C. *Talanta* **1993**, 40, 1361-1370.
- Alwarthan, A.A. *Anal. Chim. Acta* **1995**, 317, 233-237.
- Aly, F.A.; Alarfaj, N.A.; Alwarthan, A.A. *Anal. Chim. Acta*, **1998**, 358, 255-261.
- Adegoke, O.A.; Balogun, B.B. *Int. J. Pharm. Sci. Rev. Res.* **2010**, 4, 1-10.
- Shah, J.; Jan, M.R.; Naeem, M.; Inayatullah, J. *Appl. Spect.* **2012**, in the press.
- International Conference on Harmonization Topic Q2 (R1): Validation of Analytical Procedures: text and methodology, 2005; <http://www.ich.org/LOB/media/MEDIA417.pdf>.