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# Hydrate Equilibrium Modelling with the Cubic two-state Equation of State

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**Abstract:** A combination of the cubic two-state equation of state and the van der Waals–Platteeuw model was employed for the description of the hydrate formation thermodynamics, with and without inhibitors. The model parameters were determined from the properties of pure compounds and from phase equilibrium data of binary gas–water and water–inhibitor mixtures. With these parameters, predictions were performed for multicomponent gas–water–inhibitor systems. For single-gas hydrates, the average absolute deviations found for equilibrium pressure did not exceed 8%, whereas for multi-gas hydrates, these deviations were 23% maximum. The proposed model produced good predictions in the presence of inhibitors (methanol, ethanol, ethylene glycol and triethylene glycol), with increasing deviations at higher concentrations of inhibitors.

**Keywords:** hydrate; equation of state; phase equilibria; two-state model; hydrate inhibitor.

**Resumen:** Se empleó la combinación de la ecuación de estado “cubic two-state” con el modelo de van der Waals–Platteeuw para la descripción de la termodinámica de la formación de hidratos, en la presencia y ausencia de inhibidores. Los parámetros del modelo se determinaron a partir de las propiedades de los compuestos puros y de datos de equilibrio de fases de mezclas binarias gas-agua y agua-inhibidores. Para la presión de formación de hidratos de gases sencillos, se encontró un error promedio máximo de 8%, mientras que para mezclas multicomponente la desviación máxima fue de 23%. El modelo propuesto produjo buenas predicciones en la presencia de inhibidores (metanol, etanol, etilenglicol y trietilenglicol), con desviaciones crecientes a medida que se aumenta la concentración del inhibidor.

**Palabras clave:** hidratos; ecuación de estado; equilibrio de fases; modelo de dos estados; inhibidor de hidratos.

## 1. Introduction

Clathrates are polyhedral lattice structures in which a group of molecules traps other molecules. They are also known as “host–guest” complexes, inclusion compounds and, in the natural gas industry, gas hydrates. Gas hydrates are crystalline solids that are similar to ice. The temperature at which hydrates form is above the freezing point of water. These structures are characterized by a hydrogen-bonded water cage around a hydrophobic solute. Guest molecules are generally non-polar small molecules (e.g. light hydrocarbon gases, carbon dioxide and nitrogen) or polar molecules with large hydrophobic segments [1].

In natural gas processing facilities, hydrate formation can cause several problems, such as the plugging of pipelines and fittings, which prevents the free flow of gas. The understanding of this phenomenon from a theoretical point of view may help to plan and optimize prevention measures, such as the injection of inhibitors, which can displace the gas–solid equilibrium to a higher temperature or pressure. Another important field where hydrate thermodynamics has a fundamental role is the recovery of methane from the seabed. In recent years, large amounts of gas hydrate have been discovered in the seabed and in permafrost zones. Gas hydrates, in the near future, may be an important energy source.

Depending on the size and composition of the guest molecules, gas hydrates may have three different structures, namely, structures I, II and H. Normally, light gases such as methane,

ethane and nitrogen induce sI-type structures, whereas larger molecules like propane, *i*-butane, and *n*-butane form sII structures. Examples of sH hydrates are those formed by isopentane and cycloheptane. Hydrate formation is favoured at high pressures and low temperatures.

To predict hydrate formation, there are several empirical correlations and theoretical models. Wilcox et al. [2] introduced the first predictive method. They assumed that hydrates are solid solutions, treating them as an ideal liquid solution, and the equilibrium conditions were obtained from the gas–solid distribution coefficients. Then, after the determination of the crystalline structures in the 1950s, thermo-statistics models were established in order to describe the distribution of gas molecules within the hydrate cavities, such as that proposed by van der Waals and Platteeuw [3] (vdWP). On the other hand, for fluid phases, many equations of state (EoS) have been employed, from cubic EoSs [4] to more sophisticated ones like SAFT [5-7] or CPA [8, 9]. At the present time, the coupling of vdWP with EoS is the most used approach for hydrate–gas–liquid equilibrium modelling.

Recently, Medeiros and Téllez-Arredondo have presented an EoS that can deal with the association of substances, namely, the cubic two-state (CTS) EoS [10]. Its basis is the combination of the Soave–Redlich–Kwong (SRK) EoS, representing the physical part, and the two-state association model for the hydrogen bonding contribution. The CTS EoS has an equivalent polynomial form, which is easily solvable by robust numerical

methods. This model successfully describes the phase equilibria of many non-ideal systems, such as alkane–alcohol [10], water–alcohol, alcohol–alcohol [11] and alkanolamine–water–acid gas mixtures [12]. In this last paper, the CTS equation combined with an electrolyte model produced very good predictions of multicomponent systems from binary data only. Its simplicity encouraged us to employ it in modelling fluid phases in hydrate equilibria, coupled to the vdWP model. Accordingly, we organized this paper as follows. First, we present briefly the CTS EoS, the vdWP model and the hydrate–fluid-phase thermodynamics. Then, we describe the correlations between the measured data for pure compounds (water, nitrogen, hydrocarbons and inhibitors), for water–inhibitor vapour–liquid equilibrium (VLE) experimental data and for single-gas hydrate equilibrium data. The third section shows the prediction of multicomponent gas hydrate equilibrium conditions, with or without inhibitors. Finally, conclusions and additional remarks are presented in the last section.

## 2. The model

When a system reaches thermodynamic equilibrium, the fugacity of each component in all phases is equal. For hydrate(H)–Liquid(L)–Vapour(V) equilibrium this condition is expressed by the following system of equations:

$$f_{H_2O}^H(T, p, \mathbf{w}) = f_{H_2O}^L(T, p, \mathbf{x}) = f_{H_2O}^V(T, p, \mathbf{y}) \quad (1)$$

where  $f$  is the fugacity,  $T$  the temperature,  $p$  the pressure, the vectors  $\mathbf{w}$ ,  $\mathbf{x}$  and  $\mathbf{y}$  are the molar composition of the phases. As mentioned in the introduction, EoSs produce expressions to calculate the fugacities of the fluid phases, whereas the vdWP model is the common model for the hydrate phase. In this work, we employed the CTS EoS–vdWP combination, which is briefly described below.

### 2.1 The CTS equation of state

In 2008, Medeiros and Tellez-Arredondo presented the CTS EoS [10]. In this equation, the SRK EoS describes the physical contribution to the Helmholtz energy and the modified two-state association model of Cerdeiriña et al. [13] gives the association part. The complete development of the model can be seen elsewhere [10–12]. For this EoS, the equation for the pressure is given by:

$$p(\mathbf{x}, v, T) = \frac{RT}{v - b(\mathbf{x})} - \frac{a(\mathbf{x}, T)}{v[v + b(\mathbf{x})]} - \frac{RT \sum_i x_i \frac{\sum_j x_j v_{ij} f_{ij}(T)}{v[v + \sum_j x_j v_{ij} f_{ij}(T)]}}{v[v + \sum_j x_j v_{ij} f_{ij}(T)]} \quad (2)$$

where  $R$  is the gas constant,  $T$  the temperature,  $v$  the molar volume,  $\mathbf{x}$  is the vector of molar fractions  $x_i$ ,  $a$  is the mixture's

dispersive parameter and  $b$  is the van der Waals mixture co-volume. In the association term,  $v_{ij}$  is the characteristic volume of association of a pair of molecules  $i$ – $j$ . The parameter  $f_{ij}$  is given by:

$$f_{ij}(T) = e^{\frac{-E_{ij}}{RT}} - 1 \quad (3)$$

where  $E_{ij}$  is the energy of association between molecules  $i$  and  $j$ . For pure compounds, the dispersive function  $a(T)$  is given by the conventional Soave function:

$$a_i(T) = a_{0i} \left[ 1 + c_{1i} (1 - \sqrt{T_{ri}}) \right]^2 \quad (4)$$

where  $T_{ri}$  is the reduced temperature  $T_{ri} = T / T_{ci}$ . This EoS has five parameters for pure compounds:  $a_{0i}$ ,  $b_i$  and  $c_{1i}$  for the physical part and  $v_{ii}$  and  $E_{ii}$  for the association contribution. These parameters are regressed by fitting the equation to the measured saturation pressures and densities of saturated liquids.

For non-associating compounds, like hydrocarbons, CTS reduces to the SRK EoS. The combining and mixing rules for the physical part are the conventional van der Waals rules:

$$a(\mathbf{x}, T) = \sum_i \sum_j x_i x_j a_{ij}(T) \quad (5)$$

$$a_{ij}(T) = \sqrt{a_i(T) a_j(T)} (1 - k_{ij}) \quad (6)$$

$$b(\mathbf{x}) = \sum_i x_i b_i \quad (7)$$

where  $k_{ij}$  is the binary interaction parameter, which corrects the geometric mean rule for  $a_{ij}$ . For the cross-association parameters, the combining rules employed here were those suggested by Reynoso-Lopez et al. [11]:

$$v_{ij} = \min(v_{ii}, v_{jj}) \quad (8)$$

$$E_{ij} = (1 - l_{ij}) \frac{E_{ii} + E_{jj}}{2} \quad (9)$$

where  $l_{ij}$  is the binary interaction parameter for correcting the arithmetic mean rule for the cross-association energy. These binary parameters among associative substances are determined by fitting the EoS to measured VLE experimental data. From direct thermodynamics, the fugacity coefficient of species  $k$  ( $\Phi_k = f_k / x_k p$ ) from the CTS EoS is given by:

$$\begin{aligned} \ln \Phi_k = & \ln \left( \frac{v}{v - b} \right) + \frac{b_k}{b} \cdot \left( \frac{b}{v - b} - \frac{a/RT}{v + b} \right) \\ & + \frac{a}{bRT} \cdot \left( 1 + \frac{a'_k}{a} - \frac{b_k}{b} \right) \cdot \ln \left( \frac{v}{v + b} \right) + \\ & + \ln \left( \frac{v}{v + \sum_j x_j v_{kj} f_{kj}} \right) - \sum_i \frac{x_i v_{ik} f_{ik}}{v + \sum_j x_j v_{ij} f_{ij}} - \ln \left( \frac{pv}{RT} \right) \end{aligned} \quad (10)$$

where:

$$a'_k = \left( \frac{\partial}{\partial N_k} Na \right)_{T, N_{j \neq k}} \quad (11)$$

In this work, we used Eq.10 for calculating the fugacities in the fluid phases.

## 2.2 The vdWP model

In the vdWP model, the thermodynamic properties of hydrates are the result of a three-dimensional generalization of the ideal adsorption. In this model, the hydrates are solid solutions. The basic hypotheses of this model are as follows: (i) the contribution of the guest molecules to the free energy is independent of the way they occupy the cavities, i.e. the caged molecules do not distort the host lattice. (ii) Each cavity can cage only one guest molecule, which cannot diffuse out of the cavity. (iii) There are no interactions between solute molecules. (iv) Quantum effects are neglected, so that classical statistics is valid.

The result is an expression for calculating the water fugacity in the hydrate phase:

$$f_w^H = f_w^{EH} \exp \left( - \frac{\Delta \mu_w}{RT} \right) \quad (12)$$

The vdWP model allows the estimation of the water chemical-potential change in the process of enclathration of the guest molecules:

$$\Delta \mu_w = \mu_w^{EH} - \mu_w^H = -RT \sum_m v_m \ln \left( 1 - \sum_j \theta_{jm} \right) \quad (13)$$

where  $\mu$  is the chemical potential, subscript  $w$  denotes water, superscripts  $H$  and  $EH$  refer to hydrate and empty hydrate, respectively, and  $n_m$  is the number of type  $m$  cavities per water molecule. For sI hydrates  $n_1 = 1/23$  and  $n_2 = 3/23$ , and for sII structures  $n_1 = 2/17$  and  $n_2 = 1/17$ .  $q_{jm}$  is the occupancy of molecules  $j$  in type  $m$  cages. The Langmuir adsorption relation gives the occupancy, assuming one guest per cavity:

$$\theta_{mj} = \frac{C_{mj} f_j}{1 + \sum_i C_{mi} f_i} \quad (14)$$

where  $f$  is the fugacity calculated by an EoS, and  $C_{mj}$  are the Langmuir constants for the cage type  $m$  and guest molecule  $j$ . These constants can be determined by evaluating the configurational integral of guest–host and guest–guest intermolecular potentials or, empirically, by the method of Parrish and Prausnitz [14]:

$$C_{mj}(T) = \frac{A_{mj}}{T} \exp \left( \frac{B_{mj}}{T} \right) \quad (15)$$

Parrish and Prausnitz adjusted  $A_{mj}$  and  $B_{mj}$  to fit Eq.15 to the values calculated with the Kihara intermolecular potential [14]. On the other hand, Folas *et al.* [9] made them adjustable in order to correlate experimental hydrate data with the CPA EoS. They followed the suggestion of Munck *et al.* [15], which states

that adjusting these parameters allows better predictions of dissociation pressures for multicomponent hydrates. We employed the same approach.

Concerning the first hypothesis of the vdWP model, it is important to mention that there are evidences which show that guest affects host-host interactions [16]. Based on these evidences, Lee and Holder [17] devised a model which allows the calculation of the reference chemical potential depending on the guest, which is used to determine the Langmuir constants  $C_{mj}$  via the configurational integral. Here, as previously mentioned, we decided to use Eq.15 to estimate the Langmuir constants with adjustable  $A_{mj}$  and  $B_{mj}$ , following Folas *et al.* [9] and Munck *et al.* [15]. Indeed, the effect of cage distortion is implicit in the regressed parameters.

The following relation estimates the fugacity of an empty hydrate:

$$f_w^{EH} = p_w^{EH} \exp \left[ \frac{v_{EH} (p - p_w^{EH})}{RT} \right] \quad (16)$$

where  $p_w^{EH}$  is the vapour pressure of a hypothetical empty hydrate, which is assumed to be very low.  $v_{EH}$  is the molar volume of the empty hydrate, which depends on temperature and on hydrate structure but not on the guest type. For the empty hydrate vapour pressure, we employed the correlation of Sloan [1]:

$$\ln \frac{p_w^{EH}}{\text{atm}} = a - \frac{b}{T} \quad (17)$$

We evaluated the empty hydrates' molar volumes with the Avlonitis correlation [18]:

$$v_w^{EH} = v_0 \left[ 1 + k_1 (T - T_0) + k_2 (T - T_0)^2 + k_3 (T - T_0)^3 \right] \quad (18)$$

where  $T_0$  is 273.15 K. The parameters of Eqs. 17 and 18 are listed in Table 1 for sI and sII structures.

**Table 1.** Parameters for Eqs. 17 and 18, from references 1 and 18.

	sI	sII
$a$	17.44	17.332
$b / \text{K}$	6003.9	6017.6
$v_0 \times 10^6 / \text{m}^3 \cdot \text{mol}^{-1}$	22.35	22.57
$k_1 \times 10^4 / \text{K}^{-1}$	3.1075	1.9335
$k_2 \times 10^7 / \text{K}^{-2}$	5.9537	2.1768
$k_3 \times 10^{10} / \text{K}^{-3}$	1.3707	-1.4786

In the next section, we present and discuss the regressions of the model's parameters.

## 3. Parameter estimation

According to the previous section, the CTS–vdWP model has the following undetermined parameters: (i) for pure compounds, five CTS parameters ( $a_0$ ,  $b$ ,  $c_1$ ,  $v_{ii}$ ,  $E_{ii}$ ); (ii) for each

pair, two CTS binary parameters ( $k_{ij}$  and  $l_{ij}$ ); (iii) for each hydrate guest molecule, two parameters for the Parrish–Prausnitz equation (Eq. ) ( $A_{mj}$  and  $B_{mj}$ ). We applied the following strategy to calculate them:

Determination of the five parameters ( $a_0$ ,  $b$ ,  $c_1$ ,  $v_{ii}$ ,  $E_{ii}$ ) for pure species (water, nitrogen, methane, ethane, propane, *n*-butane, *i*-butane, ethylene glycol and triethylene glycol) from their measured saturation pressures and liquid densities at saturation.

Estimation of binary parameters for the pairs of associating species H<sub>2</sub>O–MEG, H<sub>2</sub>O–TEG, H<sub>2</sub>O–methanol and H<sub>2</sub>O–ethanol. The required experimental data were binary VLE data.

Regression of Parrish–Prausnitz parameters for solutes in both structures sI and sII, using gas hydrate measured data from the NIST database [19].

Regarding the water-hydrocarbon binary parameters, it is well known that predicting/correlating hydrocarbon aqueous solubility is a challenge for almost all EoSs. Economou and Tsionopoulos [20] have shown that the solubility is normally underestimated in many cases orders of magnitude with quadratic mixing rules. This happens even for more sophisticated EoSs, such as SAFT, APACT and CTS. Despite their poor solubility description, many successful coupling of EoS–vdWP for hydrate equilibrium correlation was obtained. Unconventional mixing rules are required in order to simultaneously correlate hydrate and vapor–liquid equilibrium. In the case of CTS, we are currently working on a new non-quadratic mixing rule, similar to those suggested by: (a) Prausnitz and co-workers for SRK EoS [21] and by (b) ourselves for CPA EoS [22]. In these mixing rules, water–water dispersive parameter is corrected by the presence of a hydrophobic solute. In this paper, we only focused on correlating hydrate dissociation pressures. Then, for the sake of simplicity, no binary interaction parameters were used for hydrocarbons and nitrogen, i.e.  $k_{ij}$  for all pairs including these compounds were set to 0.

### 3.1 Pure compounds

The objective function for parameter regression of the pure compounds was

$$F(\Omega) = \sum_{i=1}^{\# \text{exp}} \left( \frac{p_i^{s, \text{exp}} - p_i^{s, \text{CTS}}(\Omega)}{p_i^{s, \text{exp}}} \right)^2 + \sum_{i=1}^{\# \text{exp}} \left( \frac{\rho_i^{s, \text{exp}} - \rho_i^{s, \text{CTS}}(\Omega)}{\rho_i^{s, \text{exp}}} \right)^2 \quad (19)$$

where  $\Omega$  is the vector of the parameters [ $\Omega = (a_0, b, c_1, v_{ii}, E_{ii})$ ],  $p^s$  is the saturation pressure, and  $\rho^s$  is the liquid density at saturation conditions. Superscripts *exp* and *CTS* mean experimental and calculated with the CTS EoS. The summations run over the available measured data. Table 2 shows the parameters and deviations from experimental data. As expected, the CTS data correlated closely with the pure compounds' properties. For water, the parameters were regressed in the temperature range of hydrate formation.

### 3.2 Binary parameters for water–inhibitors

Thermodynamic inhibitors have the ability to shift the pressure–temperature equilibrium conditions. Common inhibitors in industry are alcohols, glycols and salts. These compounds have the capacity to interact with water and to inhibit cage formation. Methanol is the cheapest one, and the most used. Here, we evaluated how methanol, ethanol, ethylene glycol and triethylene glycol affected gas hydrate equilibria. Then, the VLE data of water–inhibitor mixtures were correlated using the CTS EoS. The binary parameters were determined by minimizing the following objective function:

**Table 2.** CTS pure parameters. Experimental correlated data from CDATA [23]

	$a_0$ Pa.m <sup>6</sup> .mol <sup>-2</sup>	$b \times 10^6$ m <sup>3</sup> .mol <sup>-1</sup>	$c_1$	$-E_{ii} / R$ K	$v_{ii} \times 10^6$ m <sup>3</sup> .mol <sup>-1</sup>	$T_r$ range	AAD % $p$	AAD% $\rho$
water	0.2720	15.22	0.8344	1966	2.074	0.42-0.50	0.1	0.7
methanol	0.5105	31.78	0.5137	2405	0.6958	0.55-0.90	0.29	0.11
ethanol	0.8409	47.37	0.6332	2493	0.5030	0.55-0.90	0.22	0.11
MEG	1.411	50.82	0.8857	2307	1.291	0.40-0.92	0.46	0.64
TEG	4.742	13.07	1.091	2722	0.2043	0.49-0.87	0.40	0.93
nitrogen	0.1384	26.24	0.4850	-	-	0.51-0.90	0.6	1.8
methane	0.2324	29.15	0.4486	-	-	0.48-0.89	0.3	1.8
ethane	0.5515	42.94	0.5840	-	-	0.49-0.90	0.2	1.4
propane	0.9262	58.99	0.6334	-	-	0.45-0.85	0.6	0.6
i-butane	1.2919	76.18	0.8118	-	-	0.48-0.83	0.3	0.7
n-butane	1.3594	75.12	0.7021	-	-	0.50-0.80	0.3	0.3



$$F(k_{ij}, l_{ij}) = \sum_i \left( \frac{\varphi_i^{\text{exp}} - \varphi_i^{\text{CTS}}(k_{ij}, l_{ij})}{\varphi_i^{\text{exp}}} \right)^2 \quad (20)$$

where  $\varphi$  is the bubble pressure or temperature, depending on the experimental data set. Table 3 shows the regressed parameters and the deviations produced by the EoS modelling. Figs. 1 and 2 show the phase envelopes for the glycol mixtures studied here. Alcohol–water pairs could be described with only one parameter (see reference 11), whereas for water–glycol pairs both  $k_{ij}$  and  $l_{ij}$  have nonzero values. The CTS EoS successfully correlated the VLE data.

### 3.3 Parrish–Prausnitz parameters for calculation of Langmuir constants

The parameters  $A_{mj}$  and  $B_{mj}$  in Eq. 15, required for the evaluation of the Langmuir constants, were adjusted in order to correlate the single and binary hydrate dissociation pressures. In these calculations, as previously mentioned, binary parameters for liquid water–gas and gas–gas were set to 0. The minimized objective function was:

$$F(A_{mj}, B_{mj}) = \sum_i \left[ \frac{p_i^{\text{exp}} - p_i^{\text{calc}}(A_{mj}, B_{mj})}{p_i^{\text{exp}}} \right]^2 \quad (21)$$

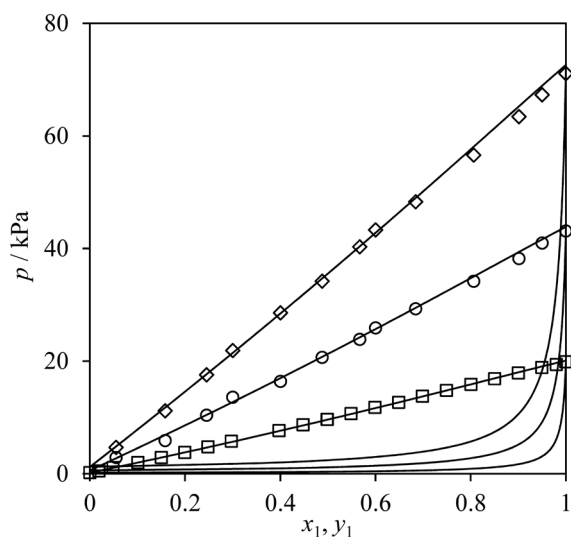
where  $p$  is the hydrate dissociation pressure. Table 4 shows the regressed constants for single and binary gas hydrates. Methane and ethane constants in the sII structures were regressed from the dissociation pressures for the mixtures methane–propane and ethane–propane. For *n*-butane constants in both structures sI and sII, the regressions ran over the data of the binaries methane–butane and propane–butane. Parameters for nitrogen in sI structures were determined from the methane–nitrogen equilibrium pressures.

Table 5 presents the correlation deviations. For the sake of comparison, this table also shows the deviations from the Bandyopadhyay–Klauda (BK) model [4]. For the BK model, the program kindly provided by Dr. Klauda generated the predictions. Both models produce quite similar results.

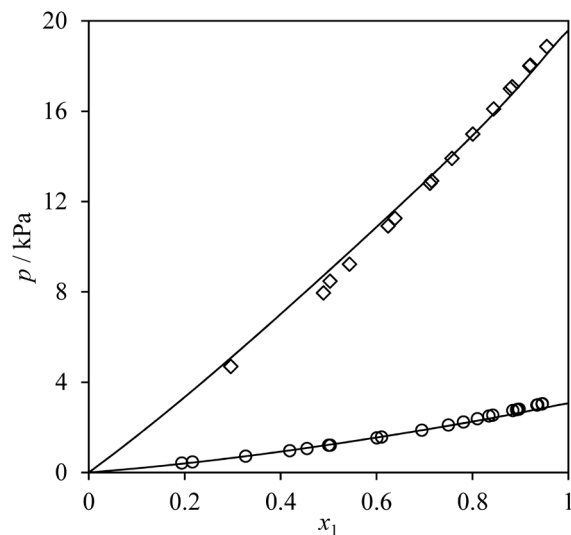
Fig. 3 shows graphically the CTS–vdWP correlation for the binary data. Very good agreement between the measured data and the correlation was achieved. The CTS–vdWP model produced average absolute deviations (AADs) from the experimental data not exceeding 8%, except for *n*-butane with 21%.

**Table 3.** Binary interaction parameters for water–inhibitor mixtures and references for experimental data

Mixture	References	$k_{ij}$	$l_{ij}$	AAD% $T$	AAD% $p$	AAD% $y$
methanol–water	[24–34]	-0.173	0	0.2	4.0	1.3
ethanol–water	[24–34]	-0.209	0	0.2	3.6	1.8
MEG–water	[24–34]	-0.0655	-0.0485	0.4	2.3	-
TEG–water	[24–34]	-0.215	-0.194	-	3.6	-



**Fig. 1.** Water(1)–MEG(2) phase envelope at 333.15 K (squares), 350.85 K (circles) and 363.45 K (diamonds). Symbols: experimental data from references [24–34]. Lines: CTS EoS.



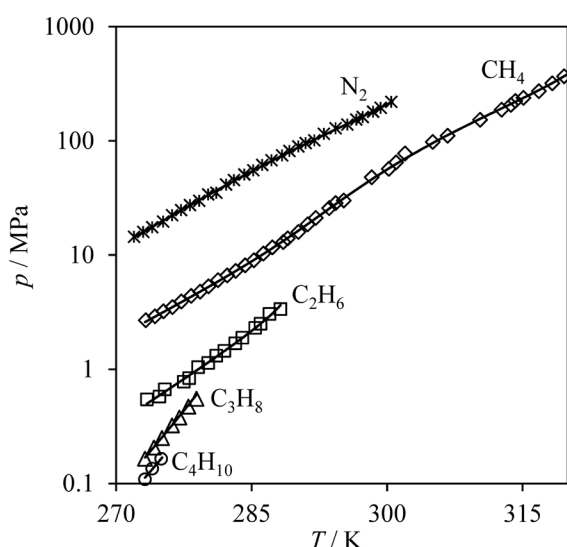
**Fig. 2.** Water(1)–TEG(2) phase envelope at 297.6 K (circles) and 332.6 K (diamonds). Experimental data from references [24–34]. Lines: CTS EoS.

**Table 4.** Parrish–Prausnitz equation constants

		Small Cavity		Large Cavity	
		$A_{mj} \times 10^8 / \text{K} \cdot \text{Pa}^{-1}$	$B_{mj} / \text{K}$	$A_{mj} \times 10^8 / \text{K} \cdot \text{Pa}^{-1}$	$B_{mj} / \text{K}$
Methane	sI	0.06511	3682	71.88	2428
	sII	4.0289	2619	22.23	174.6
Ethane	sI			13.10	3487
	sII			89.16	3454
Propane	sII			10.24	4560
i-Butane	sII			42.30	4283
n-Butane	sI			61.02	3152
	sII			62.82	3032
Nitrogen	sI	25.412	1954	40.47	2003
	sII	486.12	1301	84.68	718.4

**Table 5.** Equilibrium pressure deviations (H-L-V) and references for experimental data

	References	$T / \text{K}$	$p / \text{MPa}$	Structure	AAD% $p$	
					CTS	BK
Methane	[35,36,42-45,46-51,53-55,60,64,66]	273.2–312.7	2.6–187	sI	2.1	4.5
	[35,44,51,57,65,66]	274.5–303.7	0.3–62	sII	5.4	-
Ethane	[41,44-47,59,60]	273.4–288.2	0.5–3.4	sI	2.6	4.9
	[37-40,47,57]	273.2–279.4	0.5–1.4	sII	7.5	3.8
Propane	[35,37,44,52,59,61,64,67]	273.2–278.9	0.17–0.55	sII	3.0	3.6
i-Butane	[62,63,66]	273.2–275.1	0.11–0.17	sII	1.7	-
n-Butane	[44,51,56]	273.8–297.9	1.0–48	sI	21	42
	[58]	273.2–275.2	0.17–0.34	sII	5.0	6.6
Nitrogen	[57,68]	273.7–292.4	3.0–24.4	sI	1.3	1.9
	[46,50,69-71]	272.0–300.6	14.5–220	sII	1.4	9.6

**Fig. 3.** H-L-V equilibrium pressure for one-component hydrates. Lines: CTS–vdWP model. Symbols: experimental data for methane sI (diamonds), ethane sI (squares), propane sI (triangles), *i*-butane sII (circles) and nitrogen sII (crosses). Experimental data from references [35–71].

The agreement with the measured data was good, considering the large interval of pressures covered (from 0.1 to 220 MPa).

#### 4. Predictions: multicomponent mixtures and inhibitors

The final step in this work was to check whether the proposed model is able to predict uncorrelated dissociation pressures for multicomponent hydrates and the effect of inhibitors in shifting equilibrium conditions. The employed parameters were those determined in the previous steps and with no further regressions. In the following calculations, the binary parameters for gas–gas, gas–liquid water and gas–inhibitors were set to 0. Kontogeorgis et al. [72] employed this same approximation using the CPA EoS.

##### 4.1 Mixed-gas hydrates

Table 6 presents the studied mixtures and the sources of experimental data for comparison. Table 7 shows the predictions’

**Table 6.** Studied multicomponent gas hydrates and references for experimental data

Gas	Structure	References
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub>	sI & sII	[44,46,51,57,73]
CH <sub>4</sub> + iC <sub>4</sub> H <sub>10</sub>	sII	[44,51,66,74]
C <sub>3</sub> H <sub>8</sub> + N <sub>2</sub>	sII	[75]
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub> + C <sub>3</sub> H <sub>8</sub>	sII	[76,57]
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub> + C <sub>3</sub> H <sub>8</sub> + iC <sub>4</sub> H <sub>10</sub> (mixture 1)	sII	[77]
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub> + C <sub>3</sub> H <sub>8</sub> + C <sub>4</sub> H <sub>10</sub> + N <sub>2</sub> (mixture 2)	sII	[44]
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub> + C <sub>3</sub> H <sub>8</sub> + C <sub>4</sub> H <sub>10</sub> + iC <sub>4</sub> H <sub>10</sub> (mixture 3)	sII	[51]

deviations, and they are compared with the BK model. These mixed-gas hydrates form both sI and sII structures.

The average deviation of the predictions from the measured data was 18%, with a regular behaviour among the studied mixtures. We believe that these deviations could have been lowered if the binary parameters fixed to 0 had been adjusted, especially for water–hydrocarbon pairs. The BK model, for instance, estimates hydrocarbons solubilities via excess Gibbs energy mixing rule which better estimates the interaction between water and the non-polar compounds. Compared with other EoS-vdWP combinations, the deviations observed here are similar to those obtained with Peng-Robinson (11–28%) and PC-SAFT EoSs (15–28%), as reported by Economou and co-workers [78]. In this study, they used the quadratic rules with no binary parameters too. In the near future, we are

planning to use complex mixing rules, such as those suggested by ourselves [22], in order better describe water–hydrocarbon mutual solubilities. Fig. 4 shows the reasonable agreement with the measured pressures for CH<sub>4</sub> + *i*-C<sub>4</sub>H<sub>10</sub> and for CH<sub>4</sub> + C<sub>2</sub>H<sub>6</sub> + C<sub>3</sub>H<sub>8</sub> mixtures.

Depending on the conditions, multiple-guest hydrates can present different structures. Table 8 shows that for the mixture methane-ethane both sI and sII structures can be formed. The most stable structure is the one with the lowest equilibrium pressure. Table 8 also shows that sI is favoured at lower concentrations of methane, as mentioned by Bandyopadhyay *et al.* [4].

In summary, the correlations/predictions of hydrate equilibrium pressures have a maximum deviation of 23% in a very broad range of pressures that are common in natural gas facilities.

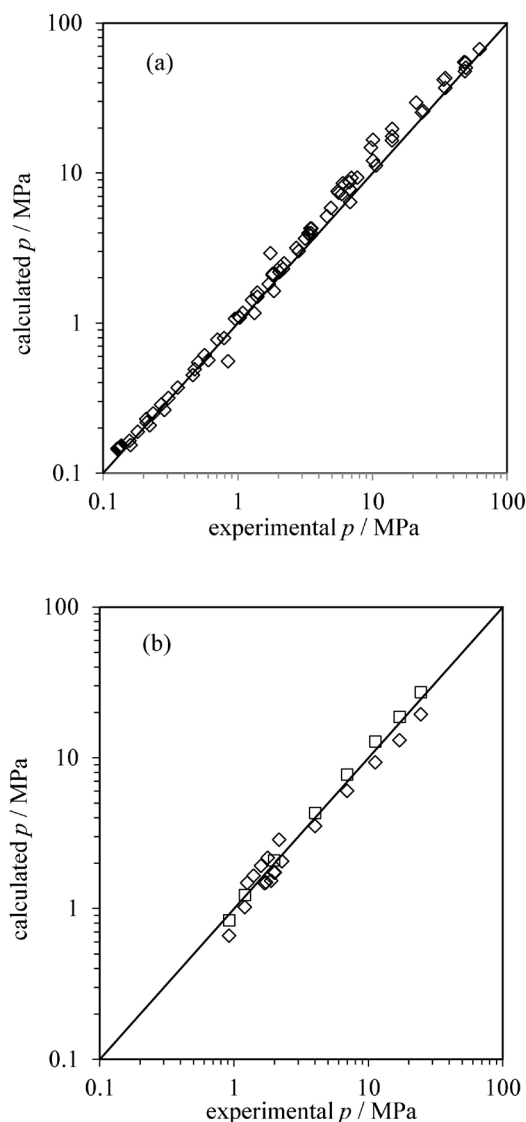
**Table 7.** Predictions' deviations from measured data for multi-gas hydrates

Mixtures	<i>T</i> / K	<i>p</i> / MPa	AAD % <i>p</i>	
			CTS	BK
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub>	274.2–301.3	0.883–48.64	12	9
CH <sub>4</sub> + iC <sub>4</sub> H <sub>10</sub>	273.8–303.1	0.159–49.06	18	-
C <sub>3</sub> H <sub>8</sub> + N <sub>2</sub>	274.2–289.2	0.332–13.71	19	67
CH <sub>4</sub> + C <sub>2</sub> H <sub>6</sub> + C <sub>3</sub> H <sub>8</sub>	275.8–298.1	0.917–24.47	16	8.9
mixture 1	273.5–281.9	0.92–2.67	18	-
mixture 2	273.7–280.4	0.793–1.813	16	24
mixture 3	293.6–301.7	13.55–52.16	21	-

**Table 8.** Comparison of different structure properties for CH<sub>4</sub> + C<sub>2</sub>H<sub>6</sub> mixtures.

<i>T</i> / K	<i>p<sub>exp</sub></i> / MPa	<i>x<sub>C<sub>2</sub>H<sub>6</sub></sub></i>	<i>p<sub>calc</sub></i> / MPa		stable structure
			sI	sII	
284.8	2.14	0.823	2.26	4.12	sI
286.2	2.66	0.823	2.72	4.73	sI
287.0	3.0	0.823	3.05	6.31	sI
278.21	2.254	0.0953	2.92	1.84	sII
279.6	2.628	0.0953	3.40	2.15	sII
283.69	4.191	0.0953	5.42	3.43	sII

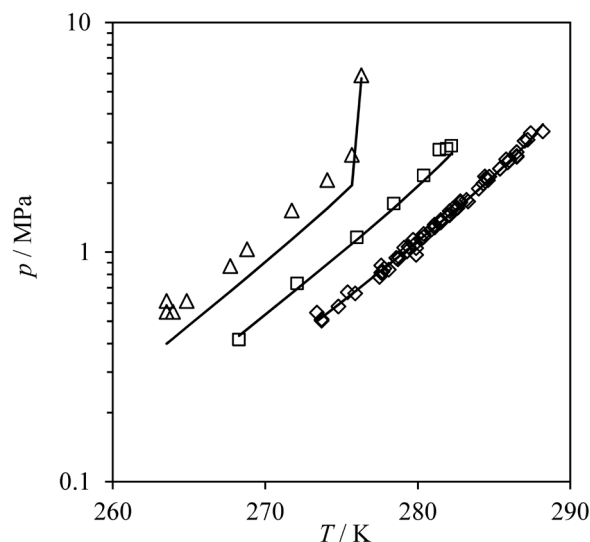




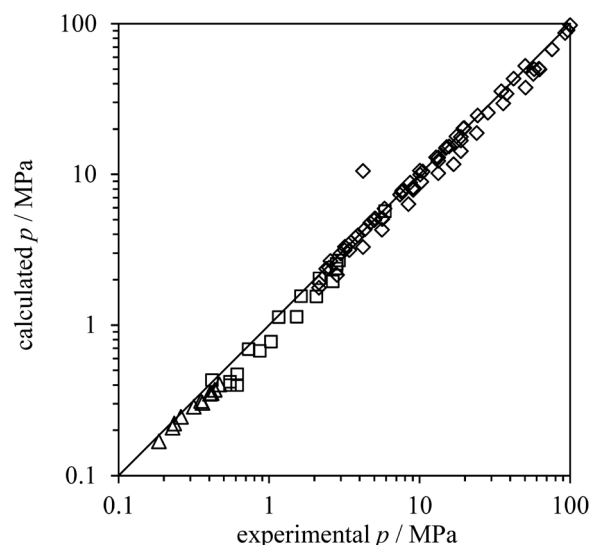
**Fig. 4.** CTS–vdWP predictions versus measured hydrate equilibrium pressures for (a)  $\text{CH}_4 + i\text{-C}_4\text{H}_{10}$  sII and (b)  $\text{CH}_4 + \text{C}_2\text{H}_6 + \text{C}_3\text{H}_8$  sII mixtures. Experimental data from references [44, 51, 57, 66, 74, 76].

#### 4.2 Gas hydrates in presence of inhibitors

Similarly to the multi-gas hydrate predictions, no regressions were performed when inhibitors were present in the gas mixtures. Furthermore, the inhibitor–gas binary parameters were also set to 0. Table 9 presents the deviations of the predictions for the mixtures studied here. Yet again, the deviations from the measured equilibrium pressures are very acceptable for all inhibitors except for 2-propanol. The overall average AAD was 13%, with a maximum of 33% (2-propanol excluded). As expected, an increase of the inhibitor fraction decreases the predicted equilibrium temperature at a given pressure. Furthermore, at constant temperature, the equilibrium pressure rises. Fig. 5 illustrates this fact. Also from this figure, the model correctly predicts the transition from HLV to HLL equilibrium when



**Fig. 5.** Hydrate equilibrium pressures as a function of temperature and methanol concentration for methane–water–methanol mixtures. Lines: CTS–vdWP predictions. Symbols: experimental data [41, 80]. Diamonds: no methanol; squares: methanol 10% wt; triangles: methanol 20% wt.



**Fig. 6.** Predictions versus experimental pressures for water–methanol (methanol 0.57%, 1.97%, 5.0%, 5.44%, 10.0%, 10.34%, 20.0%) mixtures with: methane sI (diamonds), ethane sI (squares), propane sII (triangles). Experimental data from references [41, 79–81].

20% wt methanol is present. Both the model prediction and the measured data present a subtle change in the slope at approximately 276 K. Fig. 6 presents the accuracy of the predictions of hydrate equilibrium pressures for methane sI, ethane sI and propane sII in the presence of methanol. Although the deviations presented in Table 9 reach 24% for some cases, the overall predictions are in very good agreement with the experimental data. The model estimations show that the higher the inhibitor concentrations are, the more substantial the predictions' deviations.

**Table 9.** Deviations of the predictions from measured data for gas–inhibitor hydrates and references for experimental data

Inhibitor	Gas / Structure	Inhibitor wt%	<i>T</i> / K	<i>p</i> / MPa	AAD %	References
Metanol	Methane (sI)	0.57	274.5–289.6	2.98–15.7	1.3	[79]
		1.97	271.6–291.0	2.39–19.8	2.2	[79]
		5.44	270.4–296.8	2.49–50.4	2.0	[79]
		10.	266.2–286.4	2.14–18.8	8.9	[41,80]
		20.	261.4–295.2	2.14–100	23.	[41,80,81]
	Ethane (sI)	10.	268.3–282.2	0.417–2.91	6.8	[41,79]
		20.	263.5–276.3	0.614–5.89	24.	[42,77]
	Propane (sII)	5.	272.1–274.8	0.234–0.468	9.4	[41,80]
		10.39	268.3–271.8	0.185–0.434	12.	[41,80]
Etanol	Methane (sI)	15.	273.3–284.7	3.378–13.67	20.	[82]
2-Propanol	Methane (sI)	10.	280.6–298.2	4.45–45.1	38.	[83]
		16.4	281.1–297.8	4.69–43.5	56.	[83]
		25.	281.1–299.0	5.39–49.5	54.	[83]
MEG	Methane (sI)	10	270.2–287.1	2.42–15.6	3.2	[84]
		30	267.6–279.9	3.77–16.4	10.	[84]
		50	263.4–266.5	9.89–15.2	26.	[84]
	Ethane (sI)	14.99	271.7–282.7	0.51–2.6	11.	[85]
		25	262.0–278.6	0.24–2.69	7.6	[86]
		30	269.0–276.9	0.71–2.38	7.1	[85]
		50	247.9–263.6	0.24–1.87	6.2	[86]
	Propane (sII)	9.96	272.4–274.5	0.215–0.413	10.	[87]
		14.99	271.4–273.8	0.210–0.427	28.	[87]
		29.96	264.7–266.4	0.192–0.304	25.	[87]
		10	274.6–293	3.17–25.6	4.5	[88]
		20.2	275.0–293.0	4.37–39.9	15.	[88]
		40	274.5–283.0	7.27–35.2	33.	[88]
TEG	Methane (sI)	10	274.6–293	3.17–25.6	4.5	[88]
		20.2	275.0–293.0	4.37–39.9	15.	[88]
		40	274.5–283.0	7.27–35.2	33.	[88]
	Ethane (sI)	10	273.9–283.0	0.6–1.74	27.	[88]
	Propane (sII)	10	272.3–276.8	0.18–0.51	5.3	[89]
		20	271.7–275.2	0.25–0.5	3.5	[89]
		30	270.2–272.4	0.29–0.425	19.	[89]

The assumption of null binary parameters for inhibitor–gas pairs is probably the cause of the greater deviations under those conditions. For 2-propanol, it is possible that interactions with the gas are more important and the null binary parameter leads to great deviations.

## 5. Concluding remarks

We employed the combination of van der Waals–Platteeuw with the CTS EoS model for correlating and predicting hydrate formation conditions. The modelling of hydrate dissociation pressure required the adjustment of the following model parameters: (i) CTS constants of the pure components from experimental data within a temperature range relevant for hydrate

formation; (ii) water–inhibitors (methanol, ethanol, 2-propanol, ethylene glycol and triethylene glycol) EoS binary parameters from VLE experimental data, and (iii) Parrish–Prausnitz parameters for the calculation of Langmuir constants from single-hydrate measured data. The correlations are in very good agreement with the measured data. The above mentioned parameters were then employed to predict the behaviour of multi-component hydrates, with and without thermodynamic inhibitors. The resulting predictions are very close to the measured data, except when 2-propanol is the inhibitor. It was found that when increasing the inhibitor concentration, the inhibitor–gas interactions played a greater role; therefore, binary parameters for these interactions should be taken into account. This holds especially for inhibitors capable of interacting with gases, such as 2-propanol.

Future developments will include an electrolytic contribution to the EoS. This is particularly important for studying seafloor hydrates and their dissociation conditions. The CTS-electrolyte approach successfully correlated/predicted acid gas solubilities in alkanolamines [12], and we expect that it will also work for the study of hydrates.

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## Symbols

$a_i$	Energetic parameter for pure components in the EoS
$a_0$	Energetic parameter in the EoS CTS
$a_{ij}$	Energetic parameter for $i$ - $j$ pair
$a_M$	Energetic parameter for the mixture in the EoSs
$A$	Parrish and Prausnitz adjusted parameter
$b_i$	Co-volume parameter for pure components in the EoS and CTS equation
$b_M$	Co-volume parameter for the mixture in the EoSs
$B$	Parrish and Prausnitz adjusted parameter
$c_1$	Energetic parameter in the CTS equation
$C$	Langmuir's constant
$f$	Fugacity or Mayer's function
$F_{obj}$	Objective function
$k_{ij}$	Binary interaction parameter
$l_{ij}$	Binary interaction parameter
$p, p_c$	Pressure, critical pressure
$R$	Gas constant or spherical cavity's free radius
$T, T_c, T_r$	Temperature, critical temperature, reduced temperature
$v$	Molar volume
$v_c$	Critical volume
$v^{as}$	Volume of association
$v_i$	Number of cavities type $i$ per host molecule
$v_{ij}$	Association volume for $i$ - $j$ pair
$\mathbf{x}$	Liquid molar fraction's vector
$\mathbf{y}$	Vapor molar fraction's vector

**Greek letters**

$\Phi$	Fugacity coefficient
$\mu$	Chemical potential
$\rho$	Density
$\theta$	Occupancy fraction of molecules
$\Omega$	Vector of CTS parameters

**Abbreviations**

AAD	Absolute average deviation
BL	Bandyopadhyay-Klauda

CTS	Cubic Two-State
CPA	Cubic plus association
EoS	Equation of state
EH	Empty hydrate
HLV	Hydrate-liquid-vapor
HLL	Hydrate-liquid-liquid
SRK	Soave-Redlich-Kwong
vdWP	van der Waals-Platteeuw
VLE	Vapor-liquid equilibria
SAFT	Statistical association fluid theory