Abstract

The experimental adsorption data of salicylate onto activated carbon were interpreted satisfactorily by the Prausnitz-Radke isotherm. The adsorption capacity increases diminishing the pH, but decreases increasing the temperature and when the carbon is oxidized with HNO3. At pH 5 and 11, the capacity increases lightly with the NaCl addition, but no at pH 3. It is concluded that the adsorption mechanism of salicylate is dependent upon the pH and the interaction between the salicylate molecule and the carbon surface.

Keywords

Adsorption, salicylate, Prausnitz-Radke, carbon.